Experimental delineation of the $\mathrm{CT} \rightleftharpoons \mathrm{IT}$ transformation in intermediate plagioclase feldspars

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Abstract

The $\mathrm{CT} \rightleftharpoons \mathrm{IT}$ Al/Si order/disorder transformation in intermediate plagioclase feldspars has been experimentally bracketed by dry and hydrothermal annealing of natural samples in the composition range $\mathrm{An}_{50} - \mathrm{An}_{80}$. It can be represented on the albite–anorthite binary phase diagram by a steep line passing through the points $\sim \mathrm{An}_{50}$, 1000°C and $\sim \mathrm{An}_{77}$, 1440°C (where it meets the one atmosphere solidus). The transformation appears to be characterized by a change from long range order (sharp $b$ reflections in electron diffraction patterns) to short range order (diffuse $b$ reflections) over a fairly narrow temperature interval. No sign of any two phase region separating the $\mathrm{CT}$ and $\mathrm{IT}$ fields was found on an electron optical scale and the observations are thus consistent with (but do not prove) non-first order properties for the transformation.

Assuming the transformation to be of higher than first order leads to some reappraisal of existing thermodynamic data for the high plagioclase solid solution. At the simplest level such data are consistent with approximately ideal mixing of albite and anorthite in both the $\mathrm{CT}$ and $\mathrm{IT}$ structures but with a continuous, non-ideal transition region between them. The apparent deviations from ideality can be ascribed to $\mathrm{Al}/\mathrm{Si}$ ordering in anorthite rich compositions.

The order/disorder transformation line extrapolates at lower temperatures into the exact composition range of Bøggild exsolution, and must, therefore, be implicated in that miscibility gap.

Introduction

Although a number of experimental studies of $\mathrm{Al}/\mathrm{Si}$ ordering in plagioclase feldspars have been reported in the literature, no really systematic attempts to define the stability fields of the different ordered phases appear to have been made. Much of the data on which present versions of the plagioclase subsolidus phase diagram are based (e.g., Smith, 1972, 1974, 1975) are due to Gay (1954), Gay and Bown (1956) and McConnell (1974). Their experiments provide a sound basis for estimating the approximate positions of the phase boundaries but suffer from being unreversed and, in many cases, of too short duration for complete equilibration. New data are slowly becoming available (Slimmings, 1976a; Wenk, 1978; Kroll and Müller, 1980; Tagai and Korekawa, 1981) which further constrain the stability limits and kinetics of the cation ordering but still do not unambiguously define the location of the equilibrium order–disorder boundaries. It is against this background that the present experimental investigation of the ordering behavior of plagioclases has been undertaken. In this paper we report the experimental bracketing of the $\mathrm{CT} \rightleftharpoons \mathrm{IT}$ transformation for the composition range $\sim \mathrm{An}_{50} - \mathrm{An}_{80}$. Careful reversals at two compositions and extended heat treatments at three others yield results which are consistent with the view that the transformation is non-first order in character, has a larger temperature variation with composition than proposed by Smith (1974, 1975) and occurs at more albite-rich compositions than shown by Grove et al. (1983).

In contrast with intermediate compositions the stability limits of the ordered end members, albite and anorthite, are quite well defined. Anorthite has an ordered ($\mathrm{IT}$) $\mathrm{Al}/\mathrm{Si}$ distribution up to its melting point (Czank, 1973, in Grove et al., 1983; Smith, 1974; McLaren and Marshall, 1974; Bruno et al., 1976) while at high temperatures albite is essentially $\mathrm{Al}/\mathrm{Si}$ disordered (Smith and Ribbe, 1969; Smith, 1974). The change in symmetry from $\mathrm{IT}$ to $\mathrm{CT}$ at some intermediate composition in the high temperature solid solution must surely be reflected in its thermodynamic mixing properties. Moreover, the nature of that transformation and its precise position must be determined before its influence on the complex phase relations at lower temperatures can be fully evaluated, since exsolution in the so-called peristerite, Bøggild and Huttenlocher miscibility gaps is inextricably associated with $\mathrm{Al}/\mathrm{Si}$ ordering (Smith and Ribbe, 1969; Smith, 1974, 1975;
Table 1. Details of the starting materials used for the transformation experiments. Note that compositions are for probe analyses.  

<table>
<thead>
<tr>
<th>Specimen number</th>
<th>Locality</th>
<th>Composition range</th>
<th>Most abundant composition</th>
<th>Structural state</th>
</tr>
</thead>
<tbody>
<tr>
<td>1691</td>
<td>Skaergaard, E.Greeland</td>
<td>Ab_{40-49}Or_{4}An_{30-53}</td>
<td>Ab_{51}</td>
<td>&quot;a&quot;</td>
</tr>
<tr>
<td>101249</td>
<td>Duluth, Minnesota</td>
<td>Ab_{42-50}Or_{1}An_{36-59}</td>
<td>Ab_{28}</td>
<td>&quot;a&quot;</td>
</tr>
<tr>
<td>11044</td>
<td>Duluth, Minnesota</td>
<td>Ab_{35-50}Or_{1}An_{39-53}</td>
<td>Ab_{41}</td>
<td>&quot;a&quot;</td>
</tr>
<tr>
<td>CB</td>
<td>Crystal Bay, Minnesota</td>
<td>Ab_{31-50}Or_{1}An_{39-71}</td>
<td>Ab_{70}</td>
<td>IT</td>
</tr>
<tr>
<td>54091</td>
<td>Bushveld, S.Africa</td>
<td>Ab_{25-28}Or_{1}An_{37-77}</td>
<td>Ab_{76}</td>
<td>IT + &quot;a&quot; (exsolved on a fine scale)</td>
</tr>
</tbody>
</table>

McConnell, 1974; Grove et al., 1983). It is not the purpose here, however, to go beyond a simple analysis of the high temperature solid solution; the high/low albite transformation, the displacive $\text{IT} \rightleftharpoons \text{PT}$ anorthite transformation and the exsolution reactions will not be considered further.

**Specimen description and experimental methods**

Five natural plagioclase samples in the composition range An$_{50-56}$-An$_{59}$ were selected for the experiments on the basis of their homogeneity and comparative lack of alteration or inclusions. A summary of their compositions and structural states is given in Table 1. Sample no. 1691 ($\text{An}_{52}$), from the middle gabbro zone of the Skaergaard igneous intrusion, has been described by Gay and Muir (1962). Numbers $^1$ 101249 ($\text{An}_{58}$) and 11044 ($\text{An}_{60}$) were gabbros from Duluth, Minnesota; plagioclase from the latter has already been used for experimental work by McConnell (1974). The fourth specimen ($\text{An}_{50}$), an anorthosite from Crystal Bay, Minnesota, has been described by Gay (1953, 1954). The final specimen, 54091 ($\text{An}_{59}$), an anorthosite from the Bushveld intrusion, was previously studied by Slimming (1976a). Small chips of plagioclase from the Crystal Bay anorthosite were heat treated without being crushed up, while feldspars from the other four samples were ground to a size range of 50–200 μm and separated into narrow density fractions using heavy liquids.

Grain mounts of all the samples were prepared for electron microprobe analysis. The microprobe used had a Harwell Si(Li) detector and pulse processor (Statham, 1976) and followed the correction procedures of Sweatman and Long (1969). Approximately ten grains per sample were analyzed giving compositions of Ab$_{49-56}$Or$_{1}$An$_{50-53}$ (1691); Ab$_{42-50}$Or$_{1}$An$_{56-59}$ (101249); Ab$_{42-50}$Or$_{1}$An$_{72-74}$ (Crystal Bay); and Ab$_{22-19}$Or$_{1}$An$_{77-80}$ (54091). Each analysis indicated the presence of $\sim$0.18–0.37 wt.% iron oxide (given as FeO), and the proportions of albite (Ab), orthoclase (Or) and anorthite (An) components were calculated from the alkali ratios. Wet chemical analyses giving Ab$_{50}$An$_{70}$ for the Crystal Bay feldspars (Gay, 1953) and Ab$_{21}$Or$_{1}$An$_{59}$ for 54091 plagioclase (Slimming, 1976a) suggested that the probe data have systematically low Na contents. A series of further tests on the probe using broad electron beam techniques and longer counting times showed that this systematic error was due to incomplete stripping of the energy spectra at albite poor compositions, rather than to alkali loss. The same problem did not arise at the albite richer compositions; probe data for 1691 and 11044 feldspars correspond quite closely to independent analyses. The correct compositions are therefore considered to be an An$_{50-53}$, An$_{56-59}$, An$_{59-63}$, An$_{60-71}$, An$_{74-77}$.

An AEI EM6G transmission electron microscope, operating at 100 kV, was used to characterize the structural states of starting materials and experimental products. Overexposed selected area electron diffraction patterns of 10–12 crushed feldspar fragments (deposited from alcohol onto carbon film) were examined for each run. Of principal interest was the nature of diffraction maxima at $h + k = odd, l = odd$ positions (systematically absent in the C\text{T} structure with a $c = 14\text{Å}$ cell, present in IT, or paired in the intermediate plagioclase structure). These were noted as being either absent, weak and diffuse, weak and sharp, strong and sharp, or paired. Single reflections are usually referred to as type $b$ reflections and paired reflections as type $e$. Of the starting materials, Crystal Bay plagioclase had sharp $b$ reflections, 1691, 11044 and 101249 had $e$'s and 54091 had fine scale composition modulations (with a wavelength of a few hundred \text{Å}) and both $b$ and $e$ reflections (plus diffuse, streaked $c$'s). Dark field images of ion thinned crystals from the initial powders showed the presence of type $b$ APB's up to a few μm apart in the Crystal Bay sample, and typical $e$ APB's in 1691, 11044 and 101249. In one ion thinned crystal from 11044 inconclusive evidence of a faint composition modulation was also observed, but even if really present and pervasive, it could not have had sufficient amplitude to affect the experimental results discussed below.

Dry annealing runs were performed by wrapping a few milligrams of powder, or, in the case of the Crystal Bay sample, small $\sim$0.5 mm chips, in Pt foil and suspending them in air in vertical Pt wound furnaces. Externally heated cold seal vessels were used for the hydrothermal runs for which a few mg of plagioclase powder were sealed in gold tubes with $\sim$4 mg distilled water. Run times and temperatures for the dry runs were 7–34 days, 1200–1400°C and 7–104 days, 850–1000°C for the hydrothermal treatments (P$_{H_2O} : 600$ bars). Disordered starting materials used for reversing the $\text{CT} \rightleftharpoons \text{IT}$ transition were prepared by annealing $\sim$50 mg batches of the separated powders or chips in air at high temperatures.

It was of some concern as to whether composition changes occurred during the heat treatments. Small portions of all the products were therefore examined optically in an immersion oil of refractive index close to that of...
the feldspar grains. The grains usually emerged clean and unaltered but in some cases, where the gold tube had leaked during hydrothermal annealing, were clearly partly dissolved and had noticeably higher refractive indices. Such run products were discarded. In addition, polished mounts of grains from a number of heat treated samples were prepared and these then analysed with the electron probe. No significant deviations from the composition ranges of the starting materials were found.

Results

The following changes of structure were induced experimentally: \( e \) plagioclase \( \rightarrow I1 \) with sharp \( b \) reflections; \( e \rightarrow CT \), \( b \) reflections absent or diffuse; \( CT \rightarrow I1 \) with sharp \( b \) reflections; and \( I1 \rightarrow CT \) with diffuse or absent \( b \) reflections. Both ordering and disordering reactions were thus observed and the position of the \( CT = I1 \) transformation defined. The complete results are given in Table 2.

Table 2. Details of the transformation experiments. *The cold seal bombs could only just withstand a pressure of 600 bars at 1000°C and in some cases lost some pressure during a run.

<table>
<thead>
<tr>
<th>Run no.</th>
<th>Starting material</th>
<th>Annealing temperature (°C)</th>
<th>Annealing time (days)</th>
<th>( P_{H_2O} ) (bars)</th>
<th>Product</th>
</tr>
</thead>
<tbody>
<tr>
<td>1691 (An50-53)</td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>1691/1I</td>
<td>( e ) (natural, untreated)</td>
<td>845-848</td>
<td>104</td>
<td>600</td>
<td>( v, ) diffuse intensity at ( b ) positions</td>
</tr>
<tr>
<td>1691/1I</td>
<td>( e ) (natural, untreated)</td>
<td>944-960</td>
<td>27</td>
<td>&quot;</td>
<td>( b )'s absent, or only ( v, ) diffuse intensity around ( b ) positions</td>
</tr>
<tr>
<td>1691/1I/1</td>
<td>( CT ) (annealed 13 days in air, 1255°C; diffuse/absent ( b )'s)</td>
<td>845-848</td>
<td>104</td>
<td>&quot;</td>
<td>( v, ) diffuse intensity around ( b ) positions</td>
</tr>
<tr>
<td>1691/1I/1</td>
<td>( CI ) (natural, untreated)</td>
<td>944-960</td>
<td>27</td>
<td>&quot;</td>
<td>( b )'s absent, or only ( v, ) diffuse intensity around ( b ) positions</td>
</tr>
<tr>
<td>101249 (An50-53)</td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>101249/1I</td>
<td>( e ) (natural, untreated)</td>
<td>998-1006</td>
<td>14</td>
<td>600</td>
<td>( b )'s diffuse or absent</td>
</tr>
<tr>
<td>101249/1I</td>
<td>( e ) (natural, untreated)</td>
<td>984-1002</td>
<td>24</td>
<td>500-600*</td>
<td>diffuse ( b )'s</td>
</tr>
<tr>
<td>101249/1I</td>
<td>( CT ) (annealed 14 days in air, 1275°C; diffuse/absent ( b )'s)</td>
<td>944-951</td>
<td>49</td>
<td>600</td>
<td>some crystals with sharp ( b )'s and some with diffuse ( b )'s</td>
</tr>
<tr>
<td>101249/1I</td>
<td>( CI ) (natural, untreated)</td>
<td>998-1000</td>
<td>24</td>
<td>500-600*</td>
<td>( b )'s diffuse or absent</td>
</tr>
<tr>
<td>101249/1I</td>
<td>( CI ) (natural, untreated)</td>
<td>949-955</td>
<td>28</td>
<td>600</td>
<td>most crystals with sharp ( b )'s and some with diffuse ( b )'s</td>
</tr>
</tbody>
</table>

Crystal Bay (An59-63)

<table>
<thead>
<tr>
<th>Run no.</th>
<th>Starting material</th>
<th>Annealing temperature (°C)</th>
<th>Annealing time (days)</th>
<th>( P_{H_2O} ) (bars)</th>
<th>Product</th>
</tr>
</thead>
<tbody>
<tr>
<td>11044 (An59-63)</td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>11044/1I</td>
<td>( e ) (natural, untreated)</td>
<td>999-1013</td>
<td>7</td>
<td>600</td>
<td>most crystals with diffuse ( b )'s and some with sharp ( b )'s</td>
</tr>
<tr>
<td>11044/1I</td>
<td>( CT ) (annealed 10 days in air, 1275°C; diffuse/absent ( b )'s)</td>
<td>995-1010</td>
<td>7</td>
<td>&quot;</td>
<td>sharp ( b )'s</td>
</tr>
<tr>
<td>11044/1I</td>
<td>( CI ) (natural, untreated)</td>
<td>975-992</td>
<td>24</td>
<td>500-600*</td>
<td>most crystals with sharp ( b )'s and some with diffuse ( b )'s</td>
</tr>
</tbody>
</table>

<table>
<thead>
<tr>
<th>Run no.</th>
<th>Starting material</th>
<th>Annealing temperature (°C)</th>
<th>Annealing time (days)</th>
<th>( P_{H_2O} ) (bars)</th>
<th>Product</th>
</tr>
</thead>
<tbody>
<tr>
<td>CB/5</td>
<td>( I1 ) (natural, untreated)</td>
<td>1334-1337</td>
<td>11</td>
<td>&quot;</td>
<td>dry, 1 atm.</td>
</tr>
<tr>
<td>CB/5</td>
<td>( I1 ) (natural, untreated)</td>
<td>1301-1302</td>
<td>7</td>
<td>&quot;</td>
<td>diffuse intensity at ( b ) positions</td>
</tr>
<tr>
<td>CB/10</td>
<td>( I1 ) (natural, untreated)</td>
<td>1239-1234</td>
<td>1 day 1286 + 16 days</td>
<td>&quot;</td>
<td>diffuse intensity at ( b ) positions</td>
</tr>
<tr>
<td>CB/6</td>
<td>( I1 ) (natural, untreated)</td>
<td>1301-1305</td>
<td>7</td>
<td>&quot;</td>
<td>most crystals with diffuse ( b )'s and some with sharp ( b )'s</td>
</tr>
<tr>
<td>CB/6</td>
<td>( I1 ) (natural, untreated)</td>
<td>1240-1259</td>
<td>28</td>
<td>&quot;</td>
<td>diffuse intensity at ( b ) positions</td>
</tr>
<tr>
<td>CB/6</td>
<td>( I1 ) (natural, untreated)</td>
<td>1194-1201</td>
<td>7</td>
<td>&quot;</td>
<td>diffuse intensity at ( b ) positions</td>
</tr>
<tr>
<td>CB/6</td>
<td>( I1 ) (natural, untreated)</td>
<td>1201-1215</td>
<td>34</td>
<td>&quot;</td>
<td>diffuse intensity at ( b ) positions</td>
</tr>
<tr>
<td>CB/6</td>
<td>( I1 ) (natural, untreated)</td>
<td>1201-1215</td>
<td>1333°C; diffuse/absent ( b )'s)</td>
<td>&quot;</td>
<td>diffuse intensity at ( b ) positions</td>
</tr>
<tr>
<td>CB/5</td>
<td>( I1 ) (natural, untreated)</td>
<td>1336-1390</td>
<td>10</td>
<td>&quot;</td>
<td>homogeneous crystals, some with diffuse ( b )'s and some with sharp ( b )'s</td>
</tr>
<tr>
<td>CB/5</td>
<td>( I1 ) (natural, untreated)</td>
<td>1249-1259</td>
<td>28</td>
<td>&quot;</td>
<td>homogeneous crystals with sharp, weak ( b )'s</td>
</tr>
</tbody>
</table>
and are summarized in Figure 1. A complete kinetic analysis of the Al/Si ordering and disordering was not attempted and only the critical experiments are described in this section. Evidence of exsolution was not observed in bright field TEM images of any of the run products. Attempts were made to obtain dark field images of antiphase textures but the $b$ reflections were usually too weak. In one case (11044/6/H1), however, small, normal looking $b$ domains were successfully imaged.

1691, An$_{50-53}$

Hydrothermal treatment of 1691 plagioclase crystals in both $e$ and $CT$ starting states for 104 days at $\sim$850°C and for 27 days at $\sim$950°C yielded diffuse intensity around $b$ positions in electron diffraction patterns. Even diffuse intensity could not be detected at the $b$ positions for some crystals annealed at 950°C. The $CT$ material was prepared by annealing the natural sample at $\sim$1255°C for 14 days. Both annealing temperatures are considered to have been in the $CT$ stability field.

101249, An$_{56-59}$

Annealing this sample hydrothermally at 1000°C for 24 days changed sharp $e$ reflections into diffuse or absent $b$'s. Taking grains previously disorderd at 1250°C for 14 days as starting material, a 24 day hydrothermal anneal at 1000°C also produced diffuse or absent $b$'s. Hydrothermally treating disordered ($CT$) starting material for 28 days and ordered ($e$) starting material for 49 days at $\sim$950°C produced mainly sharp $b$'s though in some crystals the $b$ reflections were diffuse. The $CT = IIT$ transformation has been placed between 950 and 1000°C for this composition.

11044, An$_{59-63}$

The critical runs at An$_{59-63}$ also involved ordered ($e$) and disordered ($CT$, heated 10 days, $\sim$1275°C in air) starting materials. These were annealed hydrothermally at $\sim$1000°C for 7 days yielding sharp $b$'s, though some crystals from the run with initially ordered material had diffuse $b$'s. The $IIT$ stability field extends to above $\sim$1000°C.

Crystal Bay, An$_{69-71}$

All runs on the Crystal Bay plagioclase were performed dry, in air. A 17 day heat treatment at $\sim$1300°C of both ordered ($IIT$) and disordered ($CT$, 11 days, $\sim$1300°C) material produced crystals with diffuse or absent $b$'s. Sharp, though weak, $b$ reflections were produced in $CT$ and $IIT$ crystals annealed for 34 days at $\sim$1200°C. At $\sim$1250°C, 28 day runs yielded only slightly diffuse $b$ reflections. The $CT = IIT$ transformation has been placed between 1200 and 1300°C at this composition, and is probably very close to 1250°C.

54091, An$_{74-77}$

A batch of powder was annealed in air for 10 days at 1360-1390°C to homogenize the fine scale composition modulations. The product had some crystals with sharp $b$ reflections and some with diffuse $b$ reflections. Annealing this for a further 7 days at $\sim$1390°C gave sharp, weak $b$ reflections, constraining the position of the $CT = IIT$ transformation to a higher temperature.

Uncertainties in the experiments arose principally from temperature drift of the furnaces, particularly during the longer runs, a possible error of $\pm 2^\circ$C in temperature determinations, and from the range of compositions within each feldspar powder. During hydrothermal runs at $\sim$1000°C, the cold seal vessels tended to expand giving some pressure variations down to 50 bars. Because the volume changes associated with the cation ordering are small, however, the differences in pressure between the dry and hydrothermal runs are not likely to displace the $CT = IIT$ boundary to a detectable degree. In addition it has been assumed that the presence of water merely influences the kinetics of the reactions without changing the relative stabilities of the structures. Further factors such as the small but variable orthoclase and iron oxide contents of the plagioclases are not likely to represent significant sources of error in placing the $CT = IIT$ boundary, given the other uncertainties in the experimental data.
Discussion

The experiments described above have been used to define a line marking the change from sharp $b$ reflections to diffuse $b$ reflections in intermediate plagioclase compositions, as a function of temperature and composition. It has been assumed that this line may be equated with the $CT \rightleftharpoons IT$, Al/Si order/disorder transformation. Such an assumption is reasonable so long as the established boundary describes equilibrium behavior. Diffuse superlattice reflections can arise under two quite different circumstances; either they may represent equilibrium short range ordering in anticipation of a transformation, at some lower temperature, to a structure with long range ordering (e.g., see Chen and Cohen, 1979), or they may be due to a non-equilibrium, intermediate state formed during the isothermal transition from a disordered to a long range ordered structure. If sharp superlattice reflections are produced from a disordered starting material, however, there are no such ambiguities and the experimental annealing conditions must have been within the equilibrium stability field of the long range ordered structure.

An important consideration in studying ordering in feldspars is the sluggishness of Al/Si diffusion. This means, at least, that the whole range of possible ordered states is quenchable, but also that even for very long heat treatments there may remain doubts over whether equilibrium really has been achieved. It is essential, therefore, either to perform a complete time-temperature-transformation analysis at each composition, so that the kinetic constraints are fully defined, or to reverse the transformations. In the present study we have followed both ordering and disordering reactions, producing sharp $b$ reflections indicative of long range ordering at four of the five compositions (Table 2). Thus, even though a thorough kinetic analysis was not attempted, it is clear that substantial intracrystalline equilibration can be achieved in a matter of days by dry annealing at temperatures greater than $\sim$1200°C and in a matter of weeks by hydrothermal treatment at 950–1000°C. The establishment of equilibrium at $\sim$850°C (run nos. 1691/H2, 1691/1/H2), however, has not been unambiguously demonstrated. A further problem exists close to the equilibrium phase boundary due to the fact that the free energy driving force for the reaction may become very small, causing a slowing down of the reaction rates. It is to be expected that the heat treatments close to the boundary may not yield an equilibrium product as readily as those on either side of it.

At $An_{59-71}$ (Crystal Bay) the transformation was bracketed between 1200 and 1300°C by following both $CT \rightarrow IT$ and $IT \rightarrow CT$. Gay and Bown (1956) reported that $b$ reflections were absent after heating the same sample at $\sim$1200°C for three days but, in the present study, 7–34 day annealing gave sharp but weak $b$'s at the same temperature. Electron diffraction is generally a more sensitive technique for detecting weak reflections than X-ray diffraction and it is possible that Gay and Bown were unable to detect these reflections in their single crystal X-ray photographs. At $\sim$1250°C, long anneals (28 days) gave diffraction intensity at the $b$ positions which was only slightly diffuse and these runs are presumed to have been very close to the $CT \rightleftharpoons IT$ boundary. At $An_{56-59}$ (101249) and $An_{59-63}$ (11044) sharp $b$ reflections were produced from crystals which had been initially ordered ($e$) or disordered ($CT$) giving unequivocal temperatures for $IT$ ordering of greater than or approximately equal to 950°C and 1000°C respectively for the two compositions. Also present in these runs, however, were some crystals with diffuse $b$ reflections, indicating that the composition ranges either straddle or come very close to the $CT/IT$ boundary at such temperatures. The time scale of 950°C treatments at $An_{58-53}$ (1691) was similar to that for $An_{56-59}$ runs which yielded crystals with long range ordering. It is concluded that the diffuse $b$ reflections which resulted represent an equilibrium (or near equilibrium) state of short range ordering, and are not due to kinetic problems of ordering.

As shown in Figure 1 the suggested position for the $CT \rightleftharpoons IT$ boundary is consistent with experimental data from other sources. Kroll and Müller (1980) annealed a range of synthetic plagioclases at high temperatures and found a change from sharp to diffuse $b$ reflections (also using electron diffraction) between $An_{67.5}$, 1350°C and $An_{77}$, 1400°C. Gay (1954) effectively bracketed the transformation between $An_{70}$ and $An_{80}$ at 1350°C. Tagai and Korukawa (1981) produced disordered crystals of composition $An_{68}$ at 1180°C from an $IT$ starting state but at 1150°C sharp $b$ reflections remained. Their experiments were of relatively short duration (160 hours), however, and they reported the presence of diffuse $e$ reflections in diffraction patterns both before and after the heat treatments. McConnell (1974) transformed $e$ reflections into $b$'s at 1010°C, $An_{65}$.

With increasing temperature, at a fixed composition, there appears to be a gradation in the nature of the $b$ reflections in samples annealed on either side of the $CT \rightleftharpoons IT$ boundary. At relatively low temperatures the $b$ reflections are sharp, but weaken as the annealing temperature is raised and then become diffuse. The diffuseness increases until at high temperatures the reflections cannot easily be detected. This is consistent with a thermodynamically non-first order transformation in which long range ordering (sharp reflections) gives way to short range ordering (diffuse reflections) over a narrow temperature interval but in a continuous way. Kroll and Müller (1980) noted that the transformation as a function of composition also appears to have no obvious discontinuity. Evidence for a high order transformation is necessarily negative, i.e., first order properties are not detected (see Carpenter and Smith, 1981, for a discussion of this problem). $IT$ is a normal subgroup of $CT$ and fulfills the Landau-Lifshitz (1958) criteria for a second order transformation, and the apparent absence of exsolution in any
of the annealed samples is also consistent with, but not proof of, other than first order transformation behavior. It can safely be stated, therefore, that the CT = IT transformation shows no signs of being first order but has the appearance, at least, of being non-first order and truly continuous in the thermodynamic sense.

**Plagioclase mixing properties**

As stated earlier, the CT = IT transformation occurring at some composition (depending on the temperature) between albite and anorthite must surely affect the thermodynamic mixing properties of the high temperature solid solution. With the added insight into the ordering behavior provided by the experiments it is instructive to reconsider data relating to plagioclase mixing from this point of view. It is necessary first, however, to discuss what form the principal parameters might have. The evidence outlined above favors a non-first order transformation with short range ordering above the transition temperature. Free energy (\(G\)), enthalpy (\(H\)), entropy (\(S\)) and an order parameter (\(\eta\)), as functions of temperature, should therefore have the forms given in Figure 2, which Thompson and Perkins (1981) define as being due to a “mixed or \(m'\)” transition. Crossing the transformation line by varying composition rather than temperature should give an analogous series of relations (Fig. 3). Because the end members of the solid solutions have different symmetries the conventional free energy (\(\Delta G_n\)), enthalpy (\(\Delta H_n\)) and entropy (\(\Delta S_n\)) of mixing, which formally should go to zero at pure albite and pure anorthite, could be mislead-

A clearer picture of the order/disorder contribution to the mixing at high temperatures is most easily illustrated by referring to CT (i.e., disordered) end members as the standard states. The new mixing parameters (\(\Delta G_m\), \(\Delta H_m\), \(\Delta S_m\)), as shown in Figure 3, do not then go to zero at \(An_{100}\). In the simplest possible case the CT solid solution could be represented by ideal mixing of albite and anorthite, with \(G\), \(H\) and \(S\) displacements due solely to a non-first order Al/Si ordering transformation (Fig. 3). CT albite is defined as having \(\eta = 0\) and IT anorthite as having \(\eta = 1\). The CT = IT transformation in pure
anorthite has a free energy, enthalpy and entropy of ordering indicated by $\Delta G_{\text{ord}}$, $\Delta H_{\text{ord}}$ and $\Delta S_{\text{ord}}$ in Figure 3.

As an aside before considering the activity data of Orville (1972) and the enthalpies of solution measured by Newton et al. (1980), it is worth highlighting a difference in their sample characterization methods compared with those of the present study. X-ray powder diffraction, and in particular the $\Delta 2\theta(131-131)$ parameter, is useful for distinguishing between "high" and "low" plagioclases (Ribbe, 1975; Kroll and Ribbe, 1980) but, unfortunately, is rather insensitive to changes of ordering in An rich compositions, and does not lead to a distinction between $\bar{I}T$ and $CT$ structures. To be sure of the state of order of synthetic crystals with compositions of An$_{50}$An$_{100}$ it is necessary to use single crystal X-ray or electron diffraction. Both Orville (1972) and Newton et al. (1980) used only powder diffraction to characterize their experimental products and there remains some uncertainty as to the state of Al/Si order in them. This represents an important limitation on the degree of confidence with which their data may be discussed in the present context.

**Enthalpy of mixing**

Enthalpies of mixing for a high plagioclase series were measured by Newton et al. (1980) using oxide melt solution calorimetry. Their crystalline samples were prepared by annealing glasses at 1200°C, 20 kbar, for 3 hours. In spite of their careful characterization by X-ray powder diffraction it is not clear whether their An rich samples achieved short range or long range ordering on an $\bar{I}T$ basis or retained a metastable $CT$ structure. The enthalpy of solution ($\Delta H_{\text{soln}}$) values they quote for An$_{100}$ are the same, within experimental error limits, for natural crystals and for samples annealed at 1200°C, 20 kbars, or 1300–1400°C, 1 bar, for 3 hours. Since Kroll and Müller (1980) showed that synthetic anorthite orders in less than 3 hours at 1400°C, 1 bar, and no natural disordered anorthites have yet been reported, it may be concluded that the $\Delta H_{\text{soln}}$ is for $\bar{I}T$ anorthite. Similar predictions of the state of order for intermediate compositions are rather less certain because the rate of ordering varies with composition, typically requiring ~5 days to reach a steady value of $\Delta 2\theta(131-131)$ at $T = 1060$–1380°C, 1 bar (Kroll and Müller, 1980). For the sake of continuing this discussion it will be assumed that some consistent degree of order occurred under the synthesis conditions in compositions where it was energetically favored. Newton et al. (1980) interpreted their data (reproduced in Fig. 4) in terms of a small, positive excess $\Delta H_m$ and used a two parameter Margules expression to describe the effect. It is not strictly valid, however, to treat a solid solution in this way if the end members have different symmetries and there is a distinct phase transformation at some intermediate composition.

According to the present analysis, the $CT \leftarrow \bar{I}T$ transformation at 1200°C (1 bar) would occur at $\sim$An$_{67}$. Pressure is unlikely to change this significantly because the volume change associated with ordering is small. An alternative explanation of the observed enthalpy effects ascribes the non-ideality to ordering at An rich compositions leaving the $CT$ structure with a linear $\Delta H_{\text{soln}}$/composition relationship, that is, with ideal mixing (Fig. 4). The form of $\Delta H_m$ should then, in the simplest case, be that shown in Figure 3b. Extrapolation of the $\Delta H_{\text{soln}}$ line for the $CT$ structure to An$_{100}$ gives an estimate for the $CT \rightarrow \bar{I}T$ ordering enthalpy in pure anorthite ($\Delta H_{\text{ord}}$) of 2.3±0.6 kcal/mole (Fig. 4).

**Activities**

Orville (1972) determined activities of the end member components, Ab and An, in the plagioclase solid solution (Fig. 5) by considering ion-exchange equilibria between synthetic crystals and a fluid. His data do not necessarily reflect equilibrium Al/Si order because thefeldspars were synthesized hydrothermally from glasses at only 800°C, 1 kbar, and then annealed in the ion exchange experiments at 700°C, 2 kbar, for up to 12 days. Significant Al/Si ordering clearly did occur, however, as evidenced by variations of $\Delta 2\theta(131-131)$ with annealing time. For the composition range An$_{85-90}$–An$_{100}$ the activities were also consistent with ideal mixing, but in the $\bar{I}T$ structure. The continuous, non-ideal range between these two states was then taken to represent the transition from $CT$ to $\bar{I}T$ symmetry. Orville also pointed out that the solid solution for An$_{90-95}$ may be
CARPENTER AND McCONNELL: CT = IT TRANSFORMATION IN PLAGIOCLASE

Fig. 5. Smoothed activity-composition relations from Orville (1972), obtained using ion exchange between crystals and a fluid at 700°C (solid lines). The straight segments correspond to ideal mixing in the CT and IT structures with a non-ideal transitional region between them (Orville, 1972). The dashed activity line for anorthite has been recalculated for a "fictive" CT anorthite end member assuming ideal mixing (\( a_{An} = \frac{X_{An}}{X_{Ca}} \)) for the composition range An\(_{65}\)-An\(_{55}\); on this scheme pure ordered IT anorthite has an activity of 0.78.

represented by an ideal CT structure having a "fictive" high albite structure, An\(_{100}\) end member. Recalculating the activities in terms of an ideal CT solid solution, with disordered albite and anorthite as standard end member states, reduces the activity of IT ordered anorthite to 0.78 (Fig. 5). This corresponds to a \( \Delta G_{ord} \) of \(-480\) cal/mole (at 700°C).

Saxena and Ribbe (1972) calculated very similar activities from Orville's data but chose to apply a continuous mathematical function to describe them. They justified their choice on the basis that there is "no reason to suppose that the structural changes are anything but continuous." A non-first order transformation allows for such changes to be continuous and Orville's interpretation is consistent with the simple analysis, illustrated in Figure 3, in which deviations from ideality are concentrated into the composition range of an order-disorder transformation. Similarly, in their reinterpretation of Orville's data Blencoe et al. (1982) also did not take into account the presence of a distinct order/disorder transformation.

**Free energy and entropy of mixing**

Orville's activity data, with CT end member standard states, may be used to calculate \( \Delta G_m' \) (Fig. 6). In form, \( \Delta G_m' \) is wholly consistent with the free energy of mixing shown in Figure 3d, i.e. ideal mixing in the composition range An\(_{65}\)-An\(_{55}\) with a continuous (non-first order) transformation to IT at calcic compositions. Assuming \( \Delta H_m' = 0 \) for the CT solid solution leads to a form of \( \Delta S_m' \) which corresponds to Figure 3c, with ideal mixing of NaAlSi\(_3\)O\(_8\) and CaAl\(_2\)Si\(_2\)O\(_8\) molecules from albite to An\(_{55}\).

Given \( \Delta G_{ord} \approx 480\) cal/mole at 700°C, and \( \Delta H_{ord} = 2.3 \pm 0.6\) kcal/mole the entropy of ordering for pure anorthite (\( \Delta S_{ord} \)) is \(-1.3\) to \(-2.5\) cal/deg/mole (see Fig. 3c). This may be compared with an extreme value for complete disorder to complete order of \( 4R(\frac{1}{2}\ln 3 + \frac{1}{2}\ln 3) \approx 5.5\) cal/deg/mole.

**Implications and conclusions**

It has not been the purpose here to present a new model for plagioclase mixing. Rather, the intention has been simply to demonstrate that deviations from ideality in the high temperature solid solution may be ascribed to cation ordering. Possible values for the free energy, enthalpy and entropy of ordering for pure anorthite have been extracted and, in principle, data can be obtained in a similar manner for CT = IT ordering in more albitic compositions. Too much credence should not necessarily be attached to these numbers, however, because of the uncertainties relating to the state of order of the experimental material from which the data were derived. As pointed out by Newton et al. (1980), for thermodynamic calculations it is not actually too critical what kind of...
model is used for a solid solution so long as it predicts correct enthalpies and entropies under changing P,T conditions. Unfortunately, although straightforward non-ideal enthalpies of mixing are expected to be fairly constant with temperature, the effects of cation ordering are generally strongly temperature dependent. Thus the heat of mixing data of Newton et al. may apply strictly only to 1200°C (assuming no large ΔC_p effects between 1200°C and the calorimeter temperature, ~700°C).

The existence of a non-first order transformation at an intermediate composition between albite and anorthite also illustrates a limitation of the Al avoidance model for the entropy of mixing terms. The expression given by Kerrick and Darken (1975) for Al avoidance effectively averages out the contributions of a true order-disorder transformation by producing a completely smooth transition between high albite and fully ordered anorthite. Probably for this reason, it fails to predict the correct composition ranges for miscibility gaps at low temperatures (Henry et al., 1982).

The position of the \( C1 = \bar{I} \) transformation line is interesting from two further points of view. First, it intersects the solidus determined by Murphy (1977, given by Henry et al., 1982) for the binary system Ab-An at \( \sim \text{An}_{77}, 1440^\circ \text{C} \). At this point the experimental solidus appears to vary quite smoothly. A marked effect would not be expected for the case of a non-first order transition, however, because the free energy of mixing for the crystalline phase would be a continuous function. Only a change in slope of the solidus might be anticipated. Second, extrapolating the \( C1 = \bar{I} \) transformation line to lower temperatures takes it exactly into the region of the Böggild miscibility gap. It would pass through \( \text{An}_{55} \) at \( \sim 750^\circ \text{C} \) and, given the steep slope, would probably not reach compositions which are much richer in albite. This coincidence with the Böggild gap and also, incidentally, with the break in structural parameters observed in natural low plagioclases (Smith and Gay, 1958; Doman et al., 1965; Bambauer et al., 1967a,b; Slimming, 1976b), must imply some involvement of the transformation in the observed exsolution behavior. It is interesting to note that in sodic pyroxenes (Carpenter, 1980) and in the peristerite range of plagioclases (Carpenter, 1981) it may be possible to relate exsolution to non-first order cation ordering reactions. Speculation along these lines raises a number of further issues, however, as outlined, for example, by Grove et al. (1983), and is unlikely to be convincing until a clearer understanding of the relationship between the \( e \) and \( \bar{I} \) ordering schemes has been established. It is notable that the Huttenlocher miscibility gap is located entirely below the \( C1 = \bar{I} \) transformation.

In conclusion, it seems likely that the key to the plagioclase solid solution lies in the order–disorder phenomena. Further experiments and thermochemical data, which define the stability limits of the different ordered phases (and of \( C2/m \), monoclinic albite) as functions of temperature, pressure and composition, are required.

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