

MAGNESIOPASCOITE, A NEW MEMBER OF THE PASCOITE GROUP: DESCRIPTION AND CRYSTAL STRUCTURE

ANTHONY R. KAMPF[§]

*Mineral Sciences Department, Natural History Museum of Los Angeles County, 900 Exposition Boulevard,
 Los Angeles, California 90007, U.S.A.*

IAN M. STEELE

Department of the Geophysical Sciences, The University of Chicago, 5734 South Ellis Avenue, Chicago, Illinois 60637, U.S.A.

ABSTRACT

Magnesiopascoite, $\text{Ca}_2\text{Mg}(\text{V}_{10}\text{O}_{28}) \cdot 16\text{H}_2\text{O}$, a new mineral species from the Blue Cap mine, near La Sal, San Juan County, Utah, U.S.A., is named as the Mg analogue of pascoite, Mg occupying one of the two nonequivalent Ca sites in the structure of pascoite. Magnesiopascoite formed from post-mining leaching and oxidation of V–U oxide replacements of organic matter. Magnesiopascoite occurs as parallel stacks of lustrous, transparent, bright-orange crystals that vary in aspect from tabular to equant to prismatic. Numerous small faces in the [110] and $[\bar{1}\bar{1}0]$ zones (at approximately 120°) yield shapes and prominent striations that mimic hexagonal symmetry. Crystal intergrowths reach several mm in the largest dimension. The mineral has a yellow streak, an adamantine luster, a Mohs hardness of about $2\frac{1}{2}$, and one cleavage, {001} perfect. The measured density is $2.43(2) \text{ g/cm}^3$, and the calculated density is 2.442 g/cm^3 . It is optically biaxial (–), α 1.769(3), β 1.802(3), γ 1.807(3), $2V$ (meas.) = $45(5)^\circ$; the dispersion, $r < v$, is crossed. The orientation is $X = b$, $Z \wedge a = 20^\circ$ (in β obtuse), and the pleochroic scheme is Y (orange) > X (yellow), Z (yellow). Electron-microprobe analyses provided: CaO 7.78, MgO 2.67, ZnO 0.23, CoO 0.05, V_2O_5 71.32, H_2O 21.94 (calc.), total 103.99% which, based on $V = 10 \text{ apfu}$ with $O = 44 \text{ apfu}$ and sufficient H for charge balance, yields $\text{Ca}_{1.77}(\text{Mg}_{0.85}\text{Zn}_{0.04}\text{Co}_{0.01})(\text{H}_2\text{O})_{15.34}(\text{H}_3\text{O})_{0.66}(\text{V}_{10}\text{O}_{28})$. Magnesiopascoite is monoclinic, $C2/m$, a 19.8442(15), b 9.9353(8), c 10.7149(8) Å, β 120.305(1) $^\circ$, V 1823.9(2) Å³, $Z = 2$. Its atomic structure [$R_1 = 2.27\%$ for $F_o > 4\sigma(F_o)$] is similar to that of other pascoite-group minerals. It consists of an anionic decavanadate $(\text{V}_{10}\text{O}_{28})^{6-}$ structural unit and an interstitial hydrated cationic complex of composition $\{\text{Ca}_2\text{Mg}(\text{H}_2\text{O})_{16}\}^{6+}$, which is formed of $\text{Mg}(\text{H}_2\text{O})_6$ octahedra and seven-coordinated $\text{CaO}_2(\text{H}_2\text{O})_5$ groups.

Keywords: magnesiopascoite, new mineral species, crystal structure, decavanadate, pascoite group, Blue Cap mine, Utah.

SOMMAIRE

Nous décrivons la magnésiopascoite, $\text{Ca}_2\text{Mg}(\text{V}_{10}\text{O}_{28}) \cdot 16\text{H}_2\text{O}$, nouvelle espèce minérale provenant de la mine Blue Cap, près de La Sal, comté de San Juan, au Utah; il s'agit de l'analogue Mg de la pascoite, le Mg occupant un des deux sites non-équivalents dans la structure de cette dernière. La magnésiopascoite est apparue suite aux exploitations minières par lessivage et oxydation des oxydes de V et de U formés par remplacement de la matière organique. Elle se présente en empilements parallèles de cristaux transparents brillants de couleur orange et d'aspect tabulaire à équidimensionnel à prismatique. De nombreuses faces dans les zones [110] et $[\bar{1}\bar{1}0]$ (à environ 120°) donnent une morphologie et des stries importantes qui rappellent une symétrie hexagonale. Des intercroissances de cristaux atteignent plusieurs mm de diamètre. Le minéral possède une rayure jaune, un éclat adamantin, une dureté de Mohs d'environ $2\frac{1}{2}$, et un clivage, {001}, parfait. La densité mesurée est $2.43(2) \text{ g/cm}^3$, et la densité calculée est 2.442 g/cm^3 . La magnésiopascoite est biaxe (–), α 1.769(3), β 1.802(3), γ 1.807(3), $2V$ (mes.) = $45(5)^\circ$; la dispersion, $r < v$, est croisée. L'orientation est $X = b$, $Z \wedge a = 20^\circ$ (β obtus), et le schéma pléochroïque est Y (orange) > X (jaune), Z (jaune). Les analyses effectuées avec une microsonde électronique ont donné: CaO 7.78, MgO 2.67, ZnO 0.23, CoO 0.05, V_2O_5 71.32, H_2O 21.94 (calc.), pour un total de 103.99%, ce qui a mené à $\text{Ca}_{1.77}(\text{Mg}_{0.85}\text{Zn}_{0.04}\text{Co}_{0.01})(\text{H}_2\text{O})_{15.34}(\text{H}_3\text{O})_{0.66}(\text{V}_{10}\text{O}_{28})$ en supposant 10 atomes de V et 44 d'oxygène par formule unitaire (apfu), et une quantité suffisante de H pour assurer l'électronéutralité. La magnésiopascoite est monoclinique, $C2/m$, a 19.8442(15), b 9.9353(8), c 10.7149(8) Å, β 120.305(1) $^\circ$, V 1823.9(2) Å³, $Z = 2$. Sa structure atomique [$R_1 = 2.27\%$ pour $F_o > 4\sigma(F_o)$] ressemble à celle des autres membres du groupe de la pascoite. Elle est faite d'une unité structurale anionique décavanadate $(\text{V}_{10}\text{O}_{28})^{6-}$ et d'un complexe interstitiel hydraté cationique de composition $\{\text{Ca}_2\text{Mg}(\text{H}_2\text{O})_{16}\}^{6+}$, formé d'octaèdres $\text{Mg}(\text{H}_2\text{O})_6$ et de groupes $\text{CaO}_2(\text{H}_2\text{O})_5$ à coordinence sept.

(Traduit par la Rédaction)

Mots-clés: magnésiopascoite, nouvelle espèce minérale, structure cristalline, décavanadate, groupe de la pascoite, mine Blue Cap, Utah.

[§] E-mail address: akampf@nhm.org

INTRODUCTION

Pascoite, $\text{Ca}_3(\text{V}_{10}\text{O}_{28}) \cdot 17\text{H}_2\text{O}$, was first described from the Ragra mine (Minasragra) near Cerro de Pasco, Pasco Department, Peru, by Hillebrand *et al.* (1914), who noted that it formed on the walls of an exploratory tunnel after its excavation. Subsequently, pascoite has been reported from several uranium mining districts in Colorado, Utah, Arizona and New Mexico, as well as from Příbram, Czech Republic and the Urcal deposit, La Rioja, Argentina.

Pascoite contains the decavanadate anionic complex, $(\text{V}_{10}\text{O}_{28})^{6-}$ (Hughes *et al.* 2005, Swallow *et al.* 1966), as do other members of the pascoite group: hummerite, $\text{K}_2\text{Mg}_2(\text{V}_{10}\text{O}_{28}) \cdot 16\text{H}_2\text{O}$ (Hughes *et al.* 2002), lasalite, $\text{Na}_2\text{Mg}_2(\text{V}_{10}\text{O}_{28}) \cdot 20\text{H}_2\text{O}$ (Hughes *et al.* 2008), and perhaps huemulite, $\text{Na}_4\text{MgV}_{10}\text{O}_{24} \cdot 24\text{H}_2\text{O}$ (Gordillo *et al.* 1966). As has been pointed out by Schindler *et al.* (2000) and others, the structures of these compounds are readily characterized as consisting of the anionic decavanadate structural unit and an interstitial hydrated cationic complex involving alkali or alkaline earth cations (or both).

In this study, we document a new member of the pascoite group, magnesiopascoite, $\text{Ca}_2\text{Mg}(\text{V}_{10}\text{O}_{28}) \cdot 16\text{H}_2\text{O}$, which is named as the Mg analogue of pascoite, Mg occupying one of the two nonequivalent Ca sites in the pascoite structure. The new mineral and name have been approved by the Commission on New Minerals, Nomenclature and Classification of the International Mineralogical Association (IMA 2007–025). The two cotype specimens are housed in the mineral collection of the Natural History Museum of Los Angeles County (catalogue numbers 58610 and 58611).

OCCURRENCE

Magnesiopascoite was discovered by Joe Marty at the Blue Cap mine, about 15 km east of La Sal, San Juan County, Utah, U.S.A. ($38^\circ 21' 14''\text{N}$, $109^\circ 03' 39''\text{W}$), in association with gypsum, rossite, pyrite, montroseite and another new mineral species, martyite (Kampf & Steele 2008) on a V–U oxide-replacement assemblage developed at the expense of organic material. Other minerals identified from the mine are brochantite, calcite, coffinite, corvusite, devilline, dickthomssenite, dolomite, lasalite, montrosite, navajoite, sherwoodite, uraninite, uranopilite and zeunerite. Magnesiopascoite was also identified on a sample collected by Joe Marty at the nearby Vanadium Queen mine.

Uranium–vanadium mineralization in the mines of the area occurs in certain layers of the sandstone paleochannels in the Salt Wash and Brushy Basin members of the Morrison Formation (Carter & Gualtieri 1965). Carbonaceous material in these layers is considered to have created a reducing environment that caused the precipitation of uranium and vanadium oxides (*e.g.*, uraninite, corvusite and montroseite) from solution.

Martyite, magnesiopascoite and rossite appear to have formed from post-mining leaching and oxidation of vanadium oxides by circulating groundwater.

The mode of occurrence and appearance of all pascoite-group minerals (pascoite, hummerite, lasalite and magnesiopascoite) are very similar, making visual identification problematic. It is, therefore, likely that magnesiopascoite exists in other deposits where it has been mistaken for other members of the group.

PHYSICAL AND OPTICAL PROPERTIES

The mineral forms as parallel stackings of lustrous, transparent, bright-orange crystals that vary in aspect from tabular to equant to prismatic. Complex intergrowths of crystals reach several mm in largest dimension (Fig. 1). The {001} form is prominent, and numerous small faces in the [110] and $[1\bar{1}0]$ zones (at approximately 120°) yield shapes and prominent striations that mimic a hexagonal symmetry. Forms in these zones include: {302}, {401}, {112}, {113}, {11 $\bar{1}$ }, {1 $\bar{1}$ 2}, {2 $\bar{2}$ 3} and {331}. No twinning was observed. We provide a crystal drawing in Figure 2.

The streak is yellow. Crystals are transparent with an adamantine luster. The Mohs hardness is about 2½. The tenacity is brittle, the fracture is conchoidal, and there is one perfect cleavage on {001}. Crystals exhibit curved fractures. The density measured by the sink–float method in sodium polytungstate solution is 2.43(2) g/cm³, and that calculated from the empirical formula is 2.442 g/cm³. Crystals dissolve slowly in H₂O and quickly in cold dilute HCl. When mildly heated, magnesiopascoite exhibits rapid decomposition, probably due to dehydration.

The mineral is optically biaxial (–), with the following optical constants measured in sodium light (589 nm): α 1.769(3), β 1.802(3), γ 1.807(3), $2V$ (meas.) = $45(5)^\circ$ and $2V$ (calc.) = 42° . Dispersion is $r < v$, crossed. The pleochroic scheme is: Y (orange) $> X$ and Z (yellow). The optical orientation is $X = b$, $Z \wedge a = 20^\circ$ (in β obtuse).

The Gladstone–Dale compatibility index $1 - (K_p/K_c)$, as defined by Mandarino (1981), provides a measure of the consistency among the average index of refraction, calculated density and chemical composition. For magnesiopascoite, the compatibility index is 0.005, indicating superior agreement among these data.

CHEMICAL COMPOSITION

Quantitative chemical analyses were performed with an electron microprobe (wavelength-dispersive spectrometry, 15 kV, 25 nA, 5 μm beam diameter). Four analyses gave the following averages (and ranges): CaO 7.78 (7.55–7.99), MgO 2.67 (2.62–2.68), ZnO 0.23 (0.12–0.28), CoO 0.05 (0.01–0.12), V₂O₅ 71.32 (69.53–73.05), H₂O 21.94 (calc.), total 103.99%. The amount of H₂O is calculated from the crystal-structure



FIG. 1. Intergrowth of magnesiopascoite crystals on V-U oxide replacement of organic material. Field of view is 2.2 mm across.

analysis. The high total may be due to rapid loss of H₂O in the electron beam. When mildly heated, magnesiopascoite exhibits rapid decomposition, probably due to dehydration.

The empirical formula (based on V = 10 atoms per formula unit (*apfu*), with O = 44 *apfu* and sufficient H for charge balance) is Ca_{1.77}(Mg_{0.85}Zn_{0.04}Co_{0.01})(H₂O)_{15.34}(H₃O)_{0.66}(V₁₀O₂₈). The end-member formula is Ca₂Mg(V₁₀O₂₈)•16H₂O, which requires: MgO 2.98, CaO 8.31, V₂O₅ 67.35, H₂O 21.35, total 100.00 wt.%.

X-RAY CRYSTALLOGRAPHY AND DETERMINATION OF THE STRUCTURE

X-ray powder-diffraction data (Table 1) were obtained using a Gandolfi camera (114.6 mm diameter,

Ni-filtered CuK α radiation). The data show good agreement with the pattern calculated from the structure.

The collection of diffracted-beam intensities was performed on a Bruker PLATFORM three-circle goniometer equipped with a 1K SMART CCD detector. A full sphere of three-dimensional data was collected. Fifty duplicate frames, acquired at the end of the data collection, indicated that no significant decay had taken place. The measured intensities were corrected for Lorentz and polarization effects using the program SAINT, and an empirical absorption-correction was applied using the program SADABS (Bruker 1997). The SHELXL97 software (Sheldrick 1997) was used for the refinement of the structure.

The positions of atoms other than hydrogen in the pascoite structure were used as a starting point. It was

TABLE 3. COORDINATES AND DISPLACEMENT PARAMETERS OF ATOMS IN MAGNESIOPASCOITE

atom	x	y	z	U_{eq}	U_{11}	U_{22}	U_{33}	U_{23}	U_{13}	U_{12}
V1	½	0.33296(10)	0	0.0154(2)	0.0169(5)	0.0131(4)	0.0195(5)	0	0.0116(4)	0
V2	0.54298(6)	½	0.79713(11)	0.0164(2)	0.0159(5)	0.0200(5)	0.0165(5)	0	0.0106(4)	0
V3	0.65679(6)	½	0.12750(11)	0.0166(2)	0.0142(5)	0.0210(5)	0.0168(5)	0	0.0095(4)	0
V4	0.38560(4)	0.34654(8)	0.67065(8)	0.0220(2)	0.0208(4)	0.0277(4)	0.0193(4)	-0.0077(3)	0.0114(3)	-0.0049(3)
O1	0.7265(3)	½	0.0902(5)	0.0233(9)	0.0189(21)	0.0305(24)	0.0248(23)	0	0.0142(18)	0
O2	0.5499(2)	½	0.1432(4)	0.0159(8)	0.0159(19)	0.0169(19)	0.0164(19)	0	0.0094(16)	0
O3	0.4160(2)	0.3755(3)	0.0321(3)	0.0159(6)	0.0171(14)	0.0156(13)	0.0187(14)	0.0004(11)	0.0117(11)	-0.0006(11)
O4	0.4498(2)	0.2316(3)	0.8553(3)	0.0214(6)	0.0249(16)	0.0177(14)	0.0254(16)	-0.0033(12)	0.0155(13)	-0.0029(12)
O5	0.6195(3)	½	0.7799(5)	0.0236(9)	0.0216(22)	0.0293(23)	0.0260(23)	0	0.0166(19)	0
O6	0.4828(2)	0.3650(3)	0.6821(3)	0.0209(6)	0.0218(15)	0.0245(15)	0.0205(15)	-0.0052(12)	0.0138(12)	-0.0026(12)
O7	0.6821(2)	0.6352(3)	0.2560(3)	0.0210(6)	0.0181(14)	0.0254(15)	0.0215(15)	-0.0045(12)	0.0115(12)	-0.0046(12)
O8	0.3495(2)	½	0.5634(5)	0.0242(9)	0.0200(21)	0.0353(25)	0.0181(21)	0	0.0101(17)	0
O9	0.3470(2)	0.2276(4)	0.5550(4)	0.0327(8)	0.0299(18)	0.0385(20)	0.0299(18)	-0.0162(16)	0.0152(15)	-0.0089(15)
Mg	0	½	½	0.0198(6)	0.0304(16)	0.0135(13)	0.0182(14)	0	0.0142(12)	0
Ca	0.76667(7)	½	0.91341(14)	0.0214(3)	0.0218(6)	0.0209(6)	0.0262(7)	0	0.0157(5)	0
O10	0.8799(4)	½	0.4285(7)	0.0415(13)	0.0388(31)	0.0368(32)	0.0476(35)	0	0.0209(27)	0
H10	0.870(6)	0.566(7)	0.468(10)	0.1173(390)						
O11	0.9785(4)	½	0.2880(6)	0.0423(14)	0.0569(38)	0.0444(35)	0.0251(28)	0	0.0202(27)	0
H11	0.966(6)	0.568(7)	0.239(9)	0.1169(400)						
O12	0.7190(2)	0.6838(5)	0.7473(5)	0.0433(10)	0.0371(22)	0.0465(24)	0.0366(22)	0.0135(20)	0.0115(18)	-0.0133(19)
H12a	0.748(5)	0.746(7)	0.739(10)	0.0995(324)						
H12b	0.713(5)	0.627(7)	0.683(7)	0.0659(260)						
O13	0	0.7037(5)	½	0.0462(16)	0.0978(51)	0.0178(24)	0.0595(39)	0	0.0667(41)	0
H13	0.001(5)	0.751(7)	0.441(7)	0.0660(245)						
O14	0.8423(2)	0.6719(4)	0.0825(4)	0.0361(9)	0.0496(23)	0.0300(19)	0.0425(22)	-0.0090(17)	0.0331(19)	-0.0164(17)
H14a	0.826(4)	0.700(7)	0.134(7)	0.0463(189)						
H14b	0.866(4)	0.729(6)	0.063(7)	0.0542(212)						
O15	0.8765(4)	½	0.8831(10)	0.0570(20)	0.0666(44)	0.0248(29)	0.1224(64)	0	0.0795(48)	0
H15	0.896(4)	0.564(6)	0.868(8)	0.0637(234)						

TABLE 4. SELECTED BOND-DISTANCES (Å) IN MAGNESIOPASCOITE

V1-		V3-		Mg-	
O4 (× 2)	1.689(3)	O1	1.623(4)	O13 (× 2)	2.023(5)
O3 (× 2)	1.911(3)	O7 (× 2)	1.804(3)	O11 (× 2)	2.090(6)
O2 (× 2)	2.134(3)	O3 (× 2)	2.012(3)	O10 (× 2)	2.104(6)
<V1-O>	1.956	O2	2.210(4)	<Mg-O>	2.072
		<V3-O>	1.968		
V2-		V4-		Ca-	
O5	1.620(4)	O9	1.600(3)	O15	2.361(6)
O6 (× 2)	1.805(3)	O8	1.824(3)	O12 (× 2)	2.387(4)
O3 (× 2)	2.010(3)	O7	1.876(3)	O14 (× 2)	2.389(4)
O2	2.236(4)	O6	1.880(3)	O1	2.397(4)
<V2-O>	1.973	O4	2.073(3)	O5	2.522(4)
		O2	2.316(3)	<Ca-O>	2.412
		<V4-O>	1.994		

(Table 4). The other two octahedra, V1(×2) in pascoite and magnesiopascoite have two short *trans* vanadyl bonds and four intermediate-length equatorial bonds. The marked bond-length variations are consistent with strong electrostatic repulsion between the V⁵⁺ cations. Furthermore, the bond-length variations lead to reasonably well-balanced bond-valences (Table 6).

In magnesiopascoite, the interstitial complex has the composition {Ca₂Mg(H₂O)₁₆}⁶⁺ and is formed of Mg(H₂O)₆ octahedra and seven-coordinated CaO₂(H₂O)₅ groups (including the two O atoms of the decavanadate

complex, to which the Ca is bonded), whereas in pascoite, the interstitial complex has the composition {Ca₂Ca(H₂O)₁₇}⁶⁺ and is formed of seven-coordinated Ca(H₂O)₇ and CaO₂(H₂O)₅ groups.

The only significant differences between the structures of magnesiopascoite and pascoite relate to the interstitial complex in each structure, and specifically to the coordinations of the two nonequivalent cation sites that form each complex. In the pascoite structure (Hughes *et al.* 2005, Swallow *et al.* 1966), the interstitial complex has the composition {Ca₂Ca(H₂O)₁₇}⁶⁺. The Ca1 site in pascoite at (0,½,½) is effectively coordinated to seven H₂O molecules, four in fully occupied sites, O10 (×2) and O11 (×2), and three in half occupied sites, O13A (×2) and O13B (×4). In magnesiopascoite, Mg is located at (0,½,½) and is coordinated to six H₂O molecules, all in fully occupied sites, O10 (×2), O11 (×2) and O13 (×2), forming a nearly perfect octahedron. In both cases, the cations bond only to H₂O molecules, which in turn link to O atoms of the decavanadate complex and the other portion of the interstitial complex only *via* H bonding.

The coordinations surrounding the other cation site, Ca2 in pascoite and Ca in magnesiopascoite, are both seven-fold and differ only in that one pair of ligated O sites (O12) is split in pascoite, but not in magnesiopas-

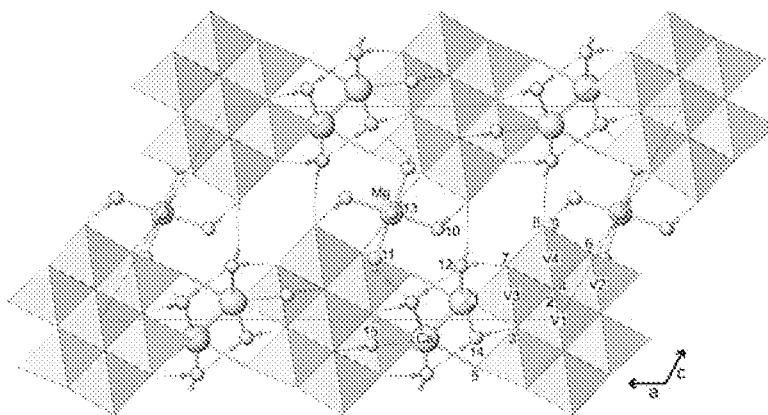


FIG. 3. Crystal structure of magnesiopascoite viewed down the *b* axis. The numbers refer to O atoms. H bonds are shown with dotted lines.

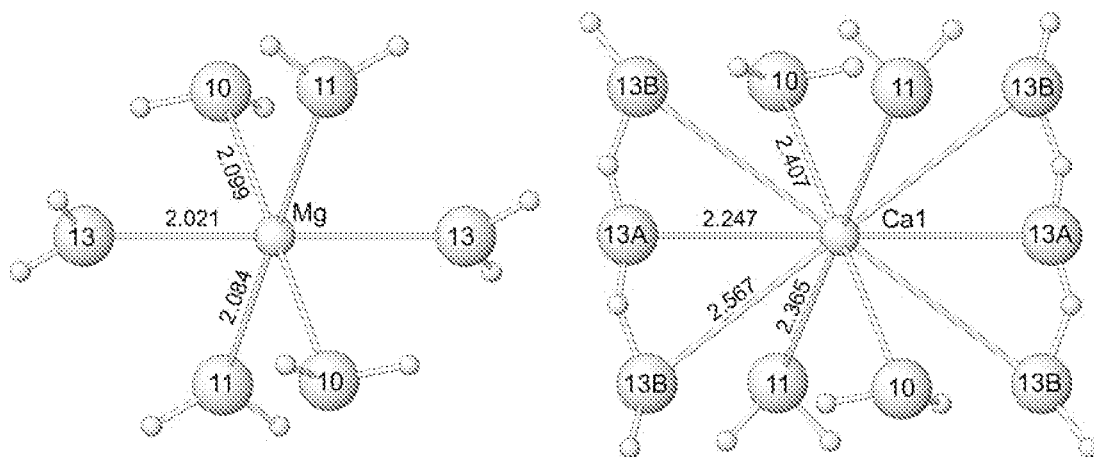


FIG. 4. Coordination around Mg in magnesiopascoite (left) and Ca1 in pascoite (right).

TABLE 5. HYDROGEN-BOND DISTANCES (Å) AND ANGLES (°) IN MAGNESIOPASCOITE

D-H	<i>d</i> (D-H)	<i>d</i> (H-A)	∠D-H-A	<i>d</i> (D-A)	A	∠H-D-H
O10-H10 (× 2)	0.85(6)	2.02(6)	178(10)	2.875(5)	O9	100
O11-H11 (× 2)	0.82(6)	2.39(8)	133(9)	3.011(7)	O14	112
O12-H12A	0.88(6)	1.80(6)	172(10)	2.676(5)	O7	101
O12-H12B	0.86(5)	2.63(8)	103(6)	2.951(5)	O9	
O13-H13 (× 2)	0.79(5)	1.89(5)	173(8)	2.676(4)	O6	107
O14-H14A	0.82(5)	2.21(5)	168(6)	3.017(7)	O12	113
O14-H14B	0.82(5)	1.89(5)	173(7)	2.706(5)	O3	
O15-H15 (× 2)	0.80(5)	2.02(5)	173(7)	2.817(4)	O4	105

D = donor O; A = acceptor O.

coite. In both cases, Ca bonds to two O atoms, O1 and O5, of the decavanadate complex and to five H₂O molecules, O12 (×2), O14 (×2) and O15. The coordinations of the cation sites of the interstitial complexes in magnesiopascoite and pascoite are compared in Figures 4 and 5.

ACKNOWLEDGEMENTS

We particularly thank Joe Marty for providing the specimens used in this study as well as for information regarding their occurrence. We thank Saeed Khan for

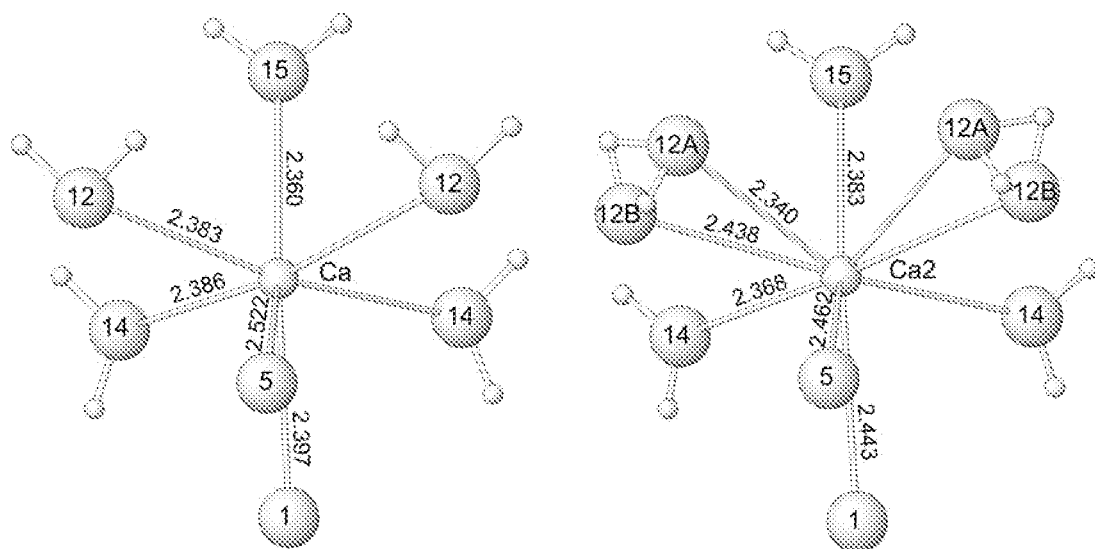


Fig. 5. Coordination around Ca in magnesiopascoite (left) and Ca2 in pascoite (right).

TABLE 6. BOND-VALENCE SUMMATIONS FOR MAGNESIOPASCOITE

	O1	O2	O3	O4	O5	O6	O7	O8	O9	O10	O11	O12	O13	O14	O15	Σ_{cv}
V1		0.409	0.747	1.361												5.034
		$\times 2!$	$\times 2 \rightarrow$	$\times 2 \rightarrow$	$\times 2 \rightarrow$											
V2		0.310	0.572		1.640	0.995										5.084
			$\times 2 \rightarrow$			$\times 2 \rightarrow$										
V3	1.627	0.333	0.568				0.997									5.090
			$\times 2 \rightarrow$				$\times 2 \rightarrow$									
V4		0.250		0.483		0.812	0.821	0.945	1.731							5.042
		$\times 2!$					$\times 2!$									
Mg										0.329	0.342		0.410			2.162
										$\times 2 \rightarrow$	$\times 2 \rightarrow$		$\times 2 \rightarrow$			
Ca	0.313				0.223							0.321		0.320	0.346	2.164
												$\times 2 \rightarrow$		$\times 2 \rightarrow$		
H10								0.160	1.680							1.000
								$\times 2 \rightarrow$								
H11											1.744			0.128		1.000
														$\times 2 \rightarrow$		
H12A							0.239					0.761				1.000
H12B									0.140			0.860				1.000
H13						0.239							1.522			1.000
							$\times 2 \rightarrow$									
H14A												0.127		0.873		1.000
H14B			0.224											0.776		1.000
H15				0.178											1.644	1.000
				$\times 2 \rightarrow$												
Σ_{av}	1.939	1.961	2.111	2.022	1.863	2.046	2.057	1.890	2.031	2.009	2.086	2.069	1.932	2.097	1.990	

Non-hydrogen bond-strengths from Brese & O'Keeffe (1991); hydrogen bond-strengths from Ferraris & Ivaldi (1988), based on O–O distances; valence summations are expressed in valence units.

his assistance with the collection of structure data in the X-Ray Crystallography Laboratory of the University of California, Los Angeles, Department of Chemistry and Biochemistry. We acknowledge Michael Schindler and John M. Hughes for their helpful reviews of the manuscript and Robert F. Martin for suggestions on improving the manuscript. This study was funded by the John Jago Trelawney Endowment to the Mineral Sciences Department of the Natural History Museum of Los Angeles County.

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Received January 4, 2008, revised manuscript accepted May 17, 2008.