## Crystal structure of tillevite: Refinement and coordination

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## Auszug

Die Atom-Parameter eines natürlichen Tilleyits wurden nach dem dreidimensionalen Ausgleichs-Verfahren verfeinert. Die Verformungen der Kolonnen von Calcium-Sauerstoff-Oktaedern lassen sich meistens durch Verkürzung der gemeinsamen Kanten erklären. In der Si<sub>2</sub>O<sub>7</sub>-Gruppe sind die Längen der (Si-O)-Brücken-Bindungen 1,656(7) und 1,678(7) Å, die der peripheren Bindungen 1,602(7) und 1,616(7) Å; der Si-O-Si-Brücken-Winkel beträgt 157°. Die inneren (O-O)-Kanten (Mittel: 2,58 Å) sind kürzer als die äußeren Kanten (Mittel: 2,71 Å); dies ist verträglich mit den Coulomb-Kräften in der Si<sub>2</sub>O<sub>7</sub>-Gruppe und mit einem  $\pi$ -Bindungs-Mechanismus.

## Abstract

The atomic parameters of a natural tilleyite  $[a = 15.108(3), b = 10.241(1), c = 7.579(1) \text{ Å}, \beta = 105.17(1)^{\circ}]$  were refined by least squares using threedimensional data. The distortions of the columns of calcium-oxygen octahedra are mostly explainable by shortening of shared edges. In the Si<sub>2</sub>O<sub>7</sub> group the Si-O bridge bonds are 1.656(7) and 1.678(7) Å and the peripheral bonds are 1.602(7) and 1.616(7) Å and the bridging Si-O-Si angle is 157°. The inner O-O edges (average 2.58 Å) are shorter than the outer edges (average 2.71 Å), consistent with Coulombic forces in the Si<sub>2</sub>O<sub>7</sub> group and with a  $\pi$ -bonding mechanism.

## Introduction

The crystal-structure determination by SMITH (1951, 1953) of tilleyite,  $Ca_5Si_2O_7(CO_3)_2$ , was incomplete because the primitive level of experimental and computational facilities made it necessary to use two-dimensional projections. Since tilleyite has a strong pseudo structure, the two-dimensional refinement resulted in averaging of some atomic coordinates which obscured details of the structure.

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A new three-dimensional refinement using intensities collected with an automatic diffractometer has now been made. The interatomic distances are discussed using recent ideas on the nature of chemical forces.

The original description of the crystal structure by SMITH was improved by BELOV (1963) who recognized tilleyite as the prototype of a series of structures. Common to these structures is a corrugated wall of linked calcium-oxygen octahedra and  $Si_2O_7$  groups (*corrugated* wall is geometrically more accurate than the term *ribbon* used in the English translation of Belov's work).

### Experimental

The cell dimensions and standard errors of tilleyite obtained for three specimens are as follows:

a	ь	c	eta	
15.111(5) Å	$10.242(2){ m \AA}$	$7.577(2)\mathrm{\AA}$	105.15(2)°	natural specimen from Barnavave, Carlingford, Ire- land.
15.079(10)	10.241(4)	7.573(3)	105.14(3)	synthetic specimen kindly supplied by Dr. R. I. HARKER of Tempres, Inc.
15.108(3)	10.241(1)	7.579(1)	105.17(1)	natural specimen from Crestmore, California, U.S.A.

X-ray powder diffractometer patterns calibrated with pure silicon (a = 5.43062 Å) as internal standard were made of the samples and the above cell dimensions were obtained by least-squares adjustment of  $1/d^2$  values using a program written by C. W. BURNHAM. A completely indexed powder pattern will be published in the Powder Diffraction File.

A single crystal was selected from Crestmore material kindly supplied by Dr. PAUL B. MOORE. A cleavage fragment of 0.007 mm<sup>3</sup> volume was mounted on the *b* axis. Fourteen levels of three-dimensional data up to a limit of  $2\theta = 80^{\circ}$  were collected with a PAILRED automatic diffractometer using monochromatic MoK $\alpha$  radiation. Of the 6067 independent reflections, 1920 had zero intensity, including

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## F. Fc Fe P 34765/42758/76677868461153178977306586750938255289807305169084809966733224575115918354417784024047681354412582222718175272321 Fe b k 227539969745607297794672807280728045112476150072734586116477914420971182345777766699178734727051657011603791926771463467870119 P 93447717591756654508404138578251219165275455615845910683724670725124982213382255845101577746714188501111857776469727374864 Fc 236631017455541124480955769602562561051243605281532662551031775629110451091016462286575825892461159816512140548572221154552222172 0216694016558818322775020669704405127770082273995176644711561177262434885164377631151177820243142001447595261582815851282615622870142179170 55551111358991501260001202678715292079935609716772158486152917169292921508723914226426917316792510725988209925511455899155167769251072595980254551455 1 12021176331313247883199803228532871961130733154797448562874135707538294144177952865475191811104226841934 7887391739709560692 72752161224071234973074961839058211565050515151671381174595129132118545812327112127113377992568899498994869104087111601573516608664145695021341165251121221045496562167221045498942566753214546046445695021341165257613817221045495656216725267522134116565621672574754403 20112345667789911223445567789912220112234556678990122312237455677990011111214678901220237455669244670112274556677

Table 1. Observed and calculated structure amplitudes in tilleyite

## Table 1. (Continued)

h k l	۴.	Fc	h k l	Fo	F.	h k l	F	F.	h k 1	F	Fe	h k l	Fo	Fc	h k l	F	Fc
23 2 2 -23	11.7	11.3	-126 2	27.5	28.0 74.4	729	15.5	13.9 23.3	241	14.7	14.4	1944 -19	7.6	7.9	14 4 8 -14	9.4 15.2	4.4 14.9
-24 -25 023	8.8 17.9	19.0* 18.4	3	4.8 23.0 9.5	21.3 4.9	-9 10	5.5	5.9	-3	18.5 8.3 9.2	18.5 2.7 9.3	-20 21 -21	25.4 7.2 6.3	27.5	-15 -16 -17	8.4 19.1 24.2	8.5 24.5 29.9
1 -1 2	21.7 14.0 20.3	22.2 14.9 20.0	5 -5 -6	7.3	6.0 21.5	11 -11 -12	18.5 4.8 3.2	14.9 5.3 4.8	5 -5 -7	14.2 15.7	11.9 15.2	-24 045	9.2 12.5	15.5* 13.2	-18 -19 -20	4.0 8.7	6.4 11.5 6 3
-4	9.7 12.4	9.1 11.9	7	5.6 29.0	6.0 31.4	-13	3.9 12.4	6.6 8.3	8	12.4	11.4	2	4.8	4.3	-21 -22	8.3 12.6	11.8 18.9*
-6 -7	8.0 9.1	9.5 10.8	9 -9	8.0 18.9	28.5 6.7 23.3	-14 15 -15	4.0 10.5 9.8	6.3 12.0	-9 -10	11.2 8.2 7.2	8.5 7.7	-3 4	11.0 21.1 18.8	13.3 22.7 14.1	-23 149 2	14.7 14.1 17.8	14.3
-8 9 -9	12.1 20.9 14.3	14.0 18.7 16.7	10 10 	11.0 11.2 19.3	8.2 12.3 16.0	-17 -20 -24	7.9 4.4 6.0	10.6 3.4 6.1	11 -11 12	18.9 7.6	19.7 5.0	5 -5 6	6.3 21.1 23.4	6.7 22.5 20.6	-2	9.6 4.6	10.6 5.2
-10	5.6	5.6	-11	35.7	34.4	0 2 10	42.6	40.9	13 -13	6.5 5.8	4.2	-6 7	19.5	21.8	-5	13,1	13.5 9.7
12 13	13.2	8.6 18.9	14	16.9 18.7	17.7	-3	16.2 11.5	13.4	15	10.4 11.2	9.7 10.6	-9	29.2 15.2	29.3 16.8	-0 -7	24.5	21.2
-15 -14 15	8.2 4.3 6.6	4.0	-15 -15 16	20.0 7.6 8.5	20.8 10.7 4.0	-4 5 -5	3.6 11.4 8.6	6.5 9.9 9.6	17 -17 -18	4.9 5.9 7.2	5.6 7.9 10.8	-10 11 13	4.2 6.6 10.9	4.4 4.3 7.7	-8 9 -9	7.5 5.9 14.4	12.7 5.1 17.4
-15 16 17	4.4 3.9 14.3	4.3 4.5 14.3	17 	8.4 4.1 11.7	6.8 6.0 10.5	-6 7	19.7 17.4 11.2	16.4 17.6 9.7	19 -19 -23	11.3 8.9 7.7	8.9 8.3 6.6	-13 -14 15	8.0 5.3 9.0	9.2 6.0 7.3	-10 -11 -15	8.5	9.9 11.9 5.7
-18 19 20	9.1 9.2	12.2 8.4	-18 19 20	16.0 8.2	24.9* 7.1	-7	22.9 17.4	24.3 15.4 27.6	-25 042	7.4	7.1	-15 -18	9.5	11.6	-17	9.7	11.4
21 -21	5.7	3.2	-21 -23	9.6 7.5	16.3*	10	11.4	8.4	-1 2	21.2 36.5	20.3 38.5	-19 20	8.9	12.9	0410	14.3	13.5
23 -23	6.8 7.9	10.4	-25 -26 -27	5.7 9.4 9.4	15.8* 17.8*	-12 13 -13	8.5 14.0 11.6	7.5*	-2 -3	22.3 75.3 7.3	24.2 69.4* 7.3	-21 -25 046	7.8 6.2 21.6	9.0 .11.1 21.4	-1 2 -2	16.7 19.5 15.5	16.5 18.4 15.7
24 -25 -27	5.7 5.5 5.5	4.3 8.5 9.4	027 -1 2	27.0 19.7 18.9	29.3 21.0 19.2	-14 -16 -17	4.1 19.0 8.0	1.9 22.1 10.4	-4 5 -5	11.8 59.6 69.4	18.5* 61.8 69.7	1 -1 2	39.2 13.6 6.8	37.9 13.8 5.1	-5	7.2 16.4 13.3	7.3 16.8 11.5
024 1	38.3 48.7 23.8	39.2 47.0 25.1	-2 3	18.6 7.0 21.0	19.9 7.0 23.1	-18 -20 -21	4.8 5.8 6 3	4.2	6 7	25.3 9.3	20.7 6.4	-2 3	20.3	21.0	5 -5	20.2	17.8
-2	9.7 27.3	10.7	-4 5	2.0	3.8 4.6	-22 0 2 11	11.2 14.6	15.9	8	24.0 34.8	25.0 35.2	-5 -5	27.8	27.3	-6 7	5.1	4.5
->	7.5	24.5 3.1 12.7	-6 7	10.3	11.3	-2 -3	4.3 5.8	0.7	-9 10 -10	6.3 15.5 33.9	9.2 13.2 32.8	-6 7	23.4 11.1 29.5	20.6 11.2 30.9	-9 10	14.3 7.5 12.2	9.6 6.1 8.7
-5 -6 -	43.4 74.7 7.1	42.7 72.4 8.2	-7 -8 9	4.4 6.9 24.7	4.6 11.2 24.9	4 5 -5	9.4 10.9 10.2	4.9 8.5 10.4	11 -11 12	47.8 35.7 8.9	49.2 35.6 5.9	-7 -8 9	20.1 3.8 27.0	23.7 6.0 21.2*	-10 -11 -12	6.7 6.6 17.8	8.0 6.5 20.1
-7 -7 8	12.4 22.7 18.9	7.9 29.4* 25.1*	~9 ~10 11	14.0 16.4 23.7	18.1 20.6	6 7 -8	5.5 21.0	4.8 17.3 8.7	-13	58.9 39.6	59.9 41.0	10 -10 -11	9.3 21.9	6.0 25.3	-14	7.7	8.3 18.8
9 -9	6.9 19.1	4.1 21.5	-11 -12	10.8	13.2	9 -9	8.8	6.1 8.5	-15	7.7	7.5	12	22.6	22.1	-17	14.7	18.7
-10 -11	56.4 38.8	58.6 38.1	-13	9.7	12.0	-11 -12	6.0 8.9	8.5	-17	7.2 12.3	10.3	-14	5.1	5.6	-20	14.1	17.5
-13	20.2	21.1 23.3*	15	10.9	6.3	- 14 - 19	4.5	1.5	-18 -19 -19	5.1	2.7	16 17	58.2 13.1 9.3	8.2 6.4	-2 3	10.1 13.9	5.4 11.7 13.7
-14	17.9 22.8	25.0* 20.3	-17 18	9.7 8.8	10.7 7.3	-1 2 12 2 -2	28.3 7.2	25.5 6.9	20 22 043	3.2 5.6 11.4	1.6 3.7 11.8	-17 -18 -19	11.6 31.7 18.2	13.8 32.9 23.1	-3 4 5	8.2 4.8 5.0	7.3 8.0 1.4
-15 -16 17	10.8 33.2 19.1	11.8 30.3 15.3	-19 -20 -21	4.3 6.6 5.8	5.2 7.3 7.1	3 -3 5	10.5 4.9 17.8	8.4 3.5 15.0	-1 2 -2	11.5 16.3 9.1	11.2 14.9 9.4	-20 -23 047	5.0 4.1 15.9	6.8 7.2	-5 6 -7	7.3 7.4 4.9	9.5 5.1 7.0
-17 18 19	14.2 24.8 21.7	20.4* 20.8 18.3	-23 -25 028	9.0 5.5 6.3	12.5 8.4 5.7	-5 6 7	7.2 8.8 7.0	8.4 5.7 5.3	-3	5.4 9.8 8.5	7.3	1 2	7.7	7.5	-15 -16	7.4	8.7
-19 -21 -22	6.6 6.4 27.8	10.8	1 -1 2	4.8	2.7 21.1	-7 -10	19.5 8.7 18.5	21.4	-6	5.0	5.4	-3	13.7	15.4	1 -1	10.8	10.9
-23 -26	13.7	23.9*	-2 -3	37.0	37.8	340 -3	4.6	6.3	8 -9	12.7	22.5* 11.6	6 7	21.0 6.0	19.2 5.4	-2 -3	11.9 10.3	3.3 12.7 9.9
125	11.6	12.5	-5	28.7	30.5	-4	43.9	41.9	-11	27.1 9.1	24.1 9.0	-8 9 -9	24.7 4.4	20.6	-4 -5 -6	9.7 13.0	9.7 13.3
-2	11.3	11.9 9.5	-7	15.5	13.1	-6	20.4 19.7	20.0 20.1 20.1	-13	25.8 6.9 9.6	7.9	-10 -11	8.9 5.3 21.1	7.0 4.5 22.1	-7 -8 -9	11.8 13.2 12.2	12.7 12.8 13.7
-) -4	20.1 39.4 16.9	28.6 36.6 18.2	-9 10	8.7 18.3 13.1	5.8 20.8 9.9	-7 -7 8	59.4 58.4 10.3	58.1 58.1 10.2	-18 -19 -21	5.0 5.5 10,4	6.6 2.6 14.4	-11 -13 14	10.5 11.8 6.9	12.7 13.4 6.3	-10 350 -3	4.0 34.3 34.0	0.3 31.6 31.6
-5 6	12.2 27.9 24.3	12.2 33.1 20.7	~10 11 -11	3.8 21.9 18.0	4.5 17.2 19.6	-8 9 -9	10.3 82.4 88.9	10.2 86.8 86.8	-22 -24 044	5.4 5.1 14.8	6.0 5.5	-14 15 -16	5.0	5.5	4 -4 5	16.6 16.8 28.5	16.6
-6 7 -7	15.4 22.8 14.1	18.5 19.3 15.7	13 -13 14	17.0 9.3 14.3	13.4	10 -10 11	28.5 27.6 10.2	26.6 26.6 7.4	1 -1	30.6 29.8	29.4 30.3	-21 -23 -24	4,4	6.3	-5	28.2	25.9
8 -8	5.5 4.9	4.8 5.3	-14 19 <sup>-1</sup>	11.1 8.1	12.4	-11 12 -12	9.3	7.4	-2	34.8 18.5	35.5 17.4	048	17.4	17.0	-7 8	51.7 34.5	49.8
-9 -10	12.2	15.1	16 17	13.9	10.3	13	20.3	19.7	-4	26.4 48.3	27.5	2	12.7	12.9	-8 9 -9	8.2 9.4	7.5
13 -13	23.0 16.5	20.7 19.5	-18 -21	15.4	19.3	-14	9.8 7.5 15.5	9.7 9.7 14.8	-5 6 -6	18.0 8.7 41.0	18.7 7.7 41.0	-3 -4	28.5 18.1 16.7	25.7 17.9 25.9*	10 -10 11	41.0 40.5 31.0	39.8 39.8 29.8
-15	5.5 10.2 11.6	4.1 11.4 10.1	-22 -24 -25	5.6 12.0 5.0	9.9 17.7 10.1	15 -16 16	15.1 4.7 4.6	14.8 5.4 5.4	7 -7 8	51.3 19.1 24.5	53.4 20.1 23.2	5 -5 6	19.1 6.5	17.9 5.9 14.8	-11 12 -12	30.5 13.2 13.5	29.8 12.3 12.3
-16 17 18	7.4 13.6 5.2	7.9 9.7 2.5	029 1 2	12.6 8.7 25.1	13.5 8.1 24.9	18 -18 20	21.8 19.6 6.2	21.1 21.1 5.2	-9 10 11	45.5 38.3 21.0	47.8 39.5 22.0	-6 7 -7	19.3 8.3 16.6	20.5 7.8 18.3	13 -13 15	10.2 9.6 9.6	9.8 9.8 9.2
-19 21 -22	6.2 12.5 7.8	9.3 9.0 7.0	-2 3 -3	4.0 9.8 4.6	4.7 8.6 2.8	-20 21 -21	5.9 14.4 12.9	5.2 14.5 14.5	-11 12 -13	59.9 12,2 7.7	60.3 2.9* 7.4	8 9 -9	16.0 20.5	18.1 16.1	-15 16 -16	10.0	9.2 12.9
-25 -26 0 2 6	10.6	14.7 3.7	4 5 -5	16.1 18.0	16.6	22 -22 24	19.2	20.3	-14	37.8	38.1 25.3*	-10	5.0	4.6	17	46.8	45.5
1	3.7	3.9	-6	11.2	11.3	-24	7.4	7.5	-17	12.9	18.1	-13	15.5	20.8	-19	22.6 21.9	22.1 22.1

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## Table 1. (Continued)

hki F	r.	h k 1	F.o	Fc	h k l	×.	F.e.	b k 1	F.	Fc	h k i	Fo	Fe	h k 1	Po	F.
$\begin{array}{cccccccccccccccccccccccccccccccccccc$	e 6.6 7.1 1.7 3 0.0 4 3 6 1 0 4 1 5 9 4 5 0 5 2 8 5 0 1 6 0 9 0 8 9 0 1 9 3 2 5 7 7 3 5 6 2 8 4 2 7 2 5 5 7 8 1 6 7 9 5 9 5 9 5 0 9 5 5 5 5 5 5 5 5 5 5 7 5 8 5 0 1 2 5 5 5 7 5 8 5 0 1 2 5 5 5 5 7 5 8 5 0 1 2 5 5 5 7 5 8 5 0 1 2 5 5 5 7 5 8 5 0 1 2 5 5 5 7 5 8 5 0 1 2 5 5 5 7 5 1 2 8 5 0 1 2 5 5 5 5 7 5 1 2 8 5 0 1 2 5 5 5 7 5 1 2 8 5 0 1 2 5 5 5 7 5 1 2 8 5 0 1 2 5 5 5 7 5 1 2 8 5 0 1 2 5 5 5 7 5 1 2 8 5 0 1 2 5 5 5 7 5 1 2 8 5 0 1 2 5 5 5 7 5 1 2 8 5 0 1 2 5 5 5 7 5 1 2 8 5 0 1 2 5 5 5 7 5 1 2 8 5 0 1 2 5 5 5 7 5 1 2 8 5 0 1 2 5 5 5 7 5 1 2 8 5 0 1 2 5 5 5 7 5 1 2 8 5 0 1 2 5 5 5 5 1 2 5 5 5 5 1 2 5 5 5 5 1 2 5 5 5 5	$ \begin{array}{c} h & k & j \\ h & k & j \\ - & 1 & 109 \\ - & 2122 & 30 \\ - & 223 & 0 \\ 1 & -1 & 2 \\ - & 2 & 2 \\ - & 2 & 2 \\ - & 2 & 3 \\ - & - & 8 \\ - & 8 & 9 \\ - & 101 \\ 1 & 11 \\ - & 1 \\ - & 1 \\ - & 2 \\ - & 1 $	ь 10620155411293121121773637105136911747850101273072878414010728151430341476411616361103161422405619721081181097391105316423402221618111422614955120630979797 08442275184227062861356194109333012294985155567699066052085294449727484106800739132818220777120732719991455512060590797	F 1347215719234111111675102092577134342757151525962862862431.600050326594261125912652570582652057059111302571525462257377111365810422757249247257242541224725724254625625717111365810422757247225720525712722527052571492459642256267177127222522500000000000000000000000000	b 4-5677899911115171-123745577911234575667899011112244455667789911122444555667901122546667790012254666777900002-2374556678990011 5 5 6 6 78990113575 6 6 6 789901135755 6 78990111224445556677899111224445556679011225466677900002-23745566778990011 1 0 0 1 1 0 0 1 1 0 0 0 1 1 0 0 0 0	2 6136279177610555981572477248314408155951281494399324404392131474111177550114234171114133817063914670811497796134652591232373246128312330 0 953552461282148339633724172483144081559512814494399324404332131474111177550114625417211413381706909485086384	2 11432387465154569101492275179282186810192841019211177755404230282866010177553111221515561843119499971126772212464564199142255576084235557608431941265255000162255000000000000000000000000	$ \frac{1}{2} + 1$	F 223114031319901510444085111444204895371141366535101728283711418108631061243710121819428718194780718051061293780914854184554366914455436691447257857081445543664775577081465418532898815587478572871487698840557909142928788521889455475577187069886095377	F 2540451215191016104277951126815510978891582725073125789518578978951812807591121423661114859154121065112516911765181975514525464219157725145414820206667512351691176518197854020554227749042717820738078404825570799512424066751235169117651819202055427749042211827290448255709994835570799462119580144148822206667575999078882000854222248225340920554227738078094883	h 1415567181922 0 1 2 2 3 3 5 6 6 7 9 9 9 0 0 1 12 3 3 5 6 6 7 7 8 9 9 0 0 0 1 1 1 1 1 1 1 1 1 1 1 1 1 1 1	F 179121516 1787874 56116 302 4210 781 91077944114 32 4118 4 22187156 8 4 436 1210 8 120 7 20 8 120 5 10 7 9 7 8 7 14 5 2 15 3 1 2 1 2 1 2 1 2 1 2 1 2 1 2 1 2 1 2 1	т 1512101611778879447105492418995705249107153162311652018926727310715910921783101613563737910612715005384541771626557731225155535514688926022 0010115422084338634899510524910715316231165201892672731071591092178310161356590129636854854185611804682377190005555544882220022	1 11714 11716 11818 199001 1223 232 37 4 5 6 6 7 7 8 8 9 7011 1714 1516 7 18 801 22 0 1 1 2 2 3 4 5 5 6 7 7 8 9 9 001 111 12 13 14 4 5 5 6 6 7 7 8 8 9 701 13 14 5 5 6 6 7 7 8 8 9 701 13 14 5 5 6 6 7 7 8 8 9 701 13 14 5 5 6 6 7 7 8 8 9 701 13 14 5 5 6 6 7 7 8 8 9 001 12 7 7 8 9 9 001 11 12 1 1 1 1 1 1 1 1 1 1 1 1 1 1 1	р 81111112541111498506789778997118849511298687175420488903688473265283551466445524332711208413285220187991113676874644697457786406401276 0 1.838515556912662812866007518177510610608209008903692704469340104576057173282834000858290407753540405141055462	F 81010212141121229944778811115954276786877572354775760697721501778287373757661595495709953141174511224234285106996013576883555457723557591111286 = 20022255511771144555555427755900237619775723547750906956956299421552242749500994314174511224242854278911227559002376195776001201
-15 8.2 17 23.7 -17 5.7	11.2 19.9 4.4	-1 -2 3	19.7 9.0 7.0	18.5 8.6 4.5	11 -11 -12	30.4 37.2	26.9 35.6 17.4	11 -11 12	15.7 15.0 20.9	12.3 17.9 21.1	-10 11 -11	15.5 20.3 20.1	15.2 20.3 20.3	12 -12 13	6.2 5.1 16.3	6,1 8,0 15,3
18 6.5	2.1	-3	22.8	23.5	13	12,4	13.0	13	7.8	6.2	12	7.2	8.2	-13	5.6	4.6

## Crystal structure of tilleyite

#### Table 1. (Continued)

## Fo F. F<sub>c</sub> x 8287297898521604616572561988540377.68859084158557745206647755878994403545746916230348099279651585754796519457565456461771092922420 F 514788907655466452920171592872776677764452746997714767545459704657124928139204511452485262886787287251275127606027551271 F 4447788004444445550066100066599977951554547841007713760545746546293888955188007445710255716440544428971841402 ۴. 0139181888531115542676482368625751822553904606677641069582442208832675935157844453523515882266551112222999779192022000199006783669676442352251188226582342435244445524444552455444456266554112222999779192022000199 +237+467899110111121513144555714920+227356749910144151671820+1225745566789101111217146718920+2274566781112171467 218449-10403299287262355616629975522199395216665539653582287654852400.1.077489482533099285679754729254038645577087544754729120120120012103

# Table 1. (Continued) F. F<sub>c</sub> F p 594454567545745754657567574575766456754575196759590691691050105767212292656751252541997564597454444675545699957807711512451105657165 p 59445946575545957549575661522664654941211511961955006916910500761212292656751252419917564597454444675575612 $\begin{array}{c} h & -10^{-1} 1 \\ -10^{-$ P 67111964601956578578571562799221666477565004302122765575476601478469954968721396743857375904568377444545826286722142486778409092456872244545886722414248427090245687234245672314424842200908895785075456075986872244548857535754624824867224144248242090 1785074652137571451712127930222523233337462221222229344438562035673785505725662842055186543119135407231624756257858209488295186542121222229344438562646485182114684374695186543521919135407316245745585740548420934844585578664648421051865445455581001445655781001457455557810000488400948864588400048865 71998545665956507479268287234661737547274574157326621747509598455945682736486732457245024991177546241595594028819777774761 4504453558464452710544565765876464876916487691668741681920167577457449991421023857776617109779241199065415584599101559159419845641559855 439357229864342514823153494222115626308197449821851597017221913239216986967987914990408827441409755166011897480505782254 470547523266771133300088884844413759915588891511110014412886585545415617944528756029192396756868555285441166605577787565005544522776553995891344155607545290708644488545529854411666005577787875660055445227765539958913445560778455290020328504270854512285042177878758854411675957657597877885

#### Table 1. (Continued)

<b>k</b> k 1	۴.	Fc	h k l	P <sub>o</sub>	F.	h k 1	F.	Fc	h k l	F.	Fe	h k l	Fo	Fe	h k 1	Fo	Fc
-8 12 9	12.5	13.0	-9 13 1	8.3	8.6	-12 13 3	5.0	5.8	6 13 6	9.9	8.0	8 14 0	6.5	6.2	10 14 2	6.1	5.2
-14	5.2	6.0	-10	7.9	8.1	-13	8.6	11.0	-7	0.0	4.0	-8	7+2	6.2 or 4	-10	12.3	10.5
0 12 10	32.5	51.0	15	5.7	4.7	14	0.3	4.9	-8	18.2	28.0*	-9	20.4	25.6	12	10.8	13 6
-2	9.0	7.8	17	5.7	2.0	-17	6.3	8.0	10	13.4	11.3	10	11.7	10.9	13	6.5	5.3
-,	6.6	6.1	-18	5.2	6.0	0 13 4	30.9	30.3	-10	5.7	7.9	-10	11.1	10.9	-13	8.7	12.7
-6	14.7	14.3	0 13 2	21.7	21.1	-1	5.3	4.5	-11	14.2	16.7	12	8.2	6.8	-14	9.3	12.3
-8	5.5	4.8	1	15.5	15.2	1	11.7	10.8	-16	5.7	7.5	-12	7.5	6.8	-15	8.9	8.7
-10	12.1	11.6	-1	13.4	14.2	2	11.4	11.0	-17	11.7	11,8	13	10.7	10_1	-17	6.6	7.2
2130	40.1	40.5	-2	26.1	26.7	-2	4.2	3.4	0 13 7	13.9	13.3	-13	9.2	10.1	0 14 3	5.9	6.0
-2	38.9	40.5	2	16.7	15.0	3	30.5	28.4	-1	6.5	6.3	14	5.1	3.8	1	3.9	4.3
2	14.1	13.4	-1	15.8	10.7	-9	36.3	35.0	2	5.0	2.4	15	4.9	4.0	-1	2.2	2.7
- 3	13.7	15.9	4	10.5	23.3"	2	14.1	12.1	-5	5,5	3.8	- 16	15.7	15.2	_1	10.7	10.5
	15.0	14.2	-5	6.7	6.4	-6	4.0	4.3	-7	13.3	15.0	17	5.0	h.5	6	13.8	11.8
	36.5	36.1	-6	11.0	9.1	-6	14.7	15.1	-9	6.4	5.7	-17	5.7	4.5	-6	8.2	8.8
-5	36.1	36.1	-6	24.2	28.1	7	6.8	4.6	-11	6.8	6.1	1 14 1	14.6	13.0	7	8.5	6.9
6	12.2	11.7	7	31.2	26.0	-7	22.5	25.1	-13	9.5	9.1	<b>⊷1</b>	2.2	4.9	-7	4.4	3.4
-6	11.7	11.7	-7	18.4	18,2	8	3.7	5.2	-14	5.8	5.9	4	7.8	10.7	-8	3.9	3.1
7	21.8	20.8	-8	9.1	9.7	-9	12.5	13.5	-15	12.9	15.3	-5	8.0	8.3	10	5.1	2.8
-7	21.4	20.8	-9	28.3	31.0	10	20.2	16.2	-16	4.4	0.3	6	5.7	4.1	11	11.9	9.8
-8	3.7	2.3	10	12.5	11.4	11	12.8	12.2	0 13 8	5.7	3.9	-6	4.0	4.2	-11	5.8	7.3
9	13.9	12.2	-10	7.4	8.7	-11	17.4	17.0	1	15.1	13.9	7	4.3	5.3	13	21.3	18.8
-9	14.1	12.2	11	0.0	3.9	12	9.7	10.9	-1	8.)	8.2	-8	2.2	15.2	-15	4.3	7.4
10	6.2	5.0	12	2.5	9.2	-12	13.1	2.2	-2	13.0	13.0	10	5.0	2.2	16	2.5	1111
-10	6.7	2.0	-15	26.1	22.6	-15	8.)	6.1	-5	9.2	1 2	-10	4.1	2.8	0 14 4	6.0	6.0
-11	7.1	2.3	-14	5.9	7.9	-1 13 #	10.5	10 6	-6	18.6	10.1	11	14.7	15.0	1	18.3	18.3
12	9.7	10.0	15	11.7	11.8	-2	3.2	2.7	-7	12.5	14.3	14	7.8	6.7	-1	12.5	11.7
-12	10.3	10.0	-15	5.1	3.4	-3	14.0	14.9	-8	10.5	10.2	15	7.5	7.8	-2	30.5	29.7
13	8.7	9.5	16	10.7	7.3	4	14.3	15.2	-9	7.5	9.6	-15	9.9	12.4	3	6.4	6.0
-13	9.4	9.5	-17	6.5	8.4	-5	5.2	5.2	-11	5.2	4.6	-16	2.9	6.4	4	19.2	17.6
14	7.9	8.5	0 13 3	3.6	6.1	7	9.1	7.4	-12	9.3	9-7	-17	5-8	4.6	-6	4.3	3.5
-14	7.8	8.5	1	8.9	7.9	-7	5.6	5.7	-13	10.4	10,1	1 14 2	3.5	1.2	7	22.1	19.3
16	2.9	5.9	-1	.5.3	5.2	9	8.8	6.7	-14	5.8	5.5	-1		6.8	-7	14.9	15.2
18	10.9	16.3	2	14.7	16.1	-9	12.7	13.2	-5 15 9	9.1	10.5	2	34.9	31.3	8	14.1	07.6
0 17 1	3.6	10.7	-2	1.0	. 2. )	-10	12.0	11.0	-0	6.7	6.7		26.1	10.0	-0	27.1	19.6
101	4.1	5.5	-3	7.1	6 1	-11	11 5	13.6	-0	4.5	6.1	-3	8.0	7.0	11	10.0	0.0
-1	12.7	13.5	4	6.3	6.0	-12	2.5	5.3	2 14 0	8.0	9.7	-4	37.0	36.6	-11	13.0	16.8
-2	7.2	11.3	-5	15.6	16.3	-14	5.1	6.0	-2	8.9	9.7	5	27.8	29.3	12	6.0	5.3
3	9.5	8.5	-6	8.1	9.5	-17	5.7	6.0	3	21.5	20.8	-5	16.1	15.1	13	5.0	8.2
-3	7.7	8.4	-7	9.3	8.6	0 13 6	5.1	1.8	-3	20.9	20.8	6	10.9	8.6	-14	9.7	11.3
4	8.6	4.0	-8	9.3	11.6	1	10.8	10.9	5	7.0	6.1	-6	12.0	13.3	-15	8.4	10.0
5	6.7	5.5	9	8.2	6.7	-1	8.9	9.4	-5	5.7	6,1	7	11.8	9.3			
þ	17.3	17.4	-9	3.8	2.0	2	12.5	11.9	6	26.7	26.3	-7	21.2	25.0			
-0	5.0	5.7	-10	2.9	3.4	-2	13.2	13.4	-6	25.4	26.3	8	0.3	11.0			
-/	9.6	9.7	11	2.9	- 2.5	- >	17.5	17.7	7	11.7	10.2	9	8.6	8.2			
-0	5.0	7.0		0.9	0.8	2	15.7	13.8	-1	11,2	10.2	-9	7+9	9.1			

the space-group extinctions for  $P2_1/a$ . The remaining 4147 data were corrected for Lorentz and polarization effects. An absorption correction for polyhedral transmission was applied using C. W. BURNHAM's subroutines ABSRP1 and ABSRP2.

The structure was refined with  $P2_1/a$ , using the SORFLS program. All hkl data with k = 3 were omitted from the calculations since an instrumental setting error was detected for this level. This left 3589 non-zero data for the 103 variable parameters comprising individual scale factors for each k level, 66 atomic positional and 22 isotropic thermal vibrational parameters and an extinction parameter (after ZACHARIASEN, 1963*a*). After six cycles of refinement with a single scale factor and unit weights for all data the agreement factor dropped to 0.18 from a starting value of 0.23. Introducing anomalous correction terms for Ca and Si (values of f' and f'' from CROMER, 1965) and individual scale factor for each k layer the agreement factor fell to 0.16 in three further cycles. The residuals correlated with  $F_{obs}$  and not with  $(\sin \theta)/\lambda$ . Omitting nine very strong and 130 medium to weak reflections that showed very large residuals from the beginning, and introducing

$$w(hkl) = 1.0/[20.0 + F_{\rm obs}(hkl) + 0.01F_{\rm obs}^2(hkl)]$$

	Table 2	
Fractional atomic coordinates	s and thermal vibrational parame	ters in tilleyite

Atom	x	<u>y</u>	z	<u> </u>
Ca(1)	0.0033(1)	0.9938(2)	0.7501(3)	0.75(2)
Ca(2)	0.1769(1)	0.2135(2)	0.0908(2)	0.69(2)
Ca(3)	0.1813(1)	0.2064(2)	0.6087(2)	0.79(2)
Ca(4)	0.1239(1)	0.5895(2)	0.0449(2)	0.62(2)
Ca(5)	0.1359(1)	0.5787(2)	0.5599(2)	0.72(2)
Si (1)	0.2032(1)	0.9209(2)	0.1381(3)	0.42(3)
Si (2)	0.2023(1)	0.9142(2)	0.5689(3)	0.57(3)
C (1)	0.0301(6)	0.3058(9)	0.3402(11)	1.14(12)
C (2)	0.0248(6)	0.3018(9)	0.8157(11)	0.84(11)
O (1)	0.0752(5)	0.2084(8)	0.2983(10)	1.10(10)
O (2)	0.0713(5)	0.2052(8)	0.7867(10)	1.08(10)
O (3)	0.0705(4)	0.3772(7)	0.4807(9)	1.08(10)
O (4)	0.0547(4)	0.3766(7)	0.9563(9)	1.01(9)
O (5)	0.0519(5)	0.6642(7)	0.7549(10)	1.37(10)
O (6)	0.0584(5)	0.6826(7)	0.2834(10)	1.25(11)
O (7)	0.2253(4)	0.7744(7)	0.0878(8)	0.73(12)
O (8)	0.2272(5)	0.7665(7)	0.6290(8)	0.81(9)
O (9)	0.1008(5)	0.9760(8)	0.0528(9)	0.98(9)
O (10)	0.1018(4)	0.9711(7)	0.5524(9)	0.83(9)
O (11)	0.2263(4)	0.5308(6)	0.8839(8)	1.01(8)
O (12)	0.2214(4)	0.5219(6)	0.3202(8)	0.77(8)
O (13)	0.2253(4)	0.9161(7)	0.3639(8)	1.18(8)

Table 3. Selected interatomic distances and bond angles

Si(1) - O(7)	1.605 Å	Si(2) - O(8)	$1.596 m \AA$
-O(9)	1.614	-O(10)	1.601
-O(11)	1.588	-O(12)	1.656
-O(13)	1.656	-O(13)	1.678
Mean	1.616	Mean	1.633
O−O distances in Si₂O7 g	roup:		
O−O distances in Si <sub>2</sub> O <sub>7</sub> g O(7)−O(9)	roup : 2.76 Å	O(8)-O(10)	2.78 Å
O−O distances in Si <sub>2</sub> O <sub>7</sub> g O(7)−O(9) O(7)−O(11)	roup : 2.76 Å 2.72	O(8) - O(10) O(8) - O(12)	2.78 Å 2.73
O-O distances in Si <sub>2</sub> O <sub>7</sub> g O(7)-O(9) O(7)-O(11) O(7)-O(13)	roup : 2.76 Å 2.72 2.55	O(8) - O(10) O(8) - O(12) O(8) - O(13)	2.78 Å 2.73 2.52
O-O distances in $Si_2O_7$ g O(7)-O(9) O(7)-O(11) O(7)-O(13) O(9)-O(11)*	roup : 2.76 Å 2.72 2.55 2.59	O(8) - O(10) O(8) - O(12) O(8) - O(13) $O(10) - O(12)^*$	2.78 Å 2.73 2.52 2.64
$\begin{array}{c} O-O \text{ distances in } Si_2O_7 \text{ g} \\ O(7)-O(9) \\ O(7)-O(11) \\ O(7)-O(13) \\ O(9)-O(11)^* \\ O(9)-O(13) \end{array}$	roup: 2.76 Å 2.72 2.55 2.59 2.67	$\begin{array}{c} O(8) - O(10) \\ O(8) - O(12) \\ O(8) - O(13) \\ O(10) - O(12)^* \\ O(10) - O(13) \end{array}$	2.78 Å 2.73 2.52 2.64 2.69

\* Shared edges.

Table 3. (Continued)										
C-O distances:										
C(1) - O(1)	$1.29~{ m \AA}$	C(2) - O(2)	1.27 Å							
-O(3)	1.30	-O(4)	1.29							
-O(5)	1.30	-O(6)	1.29							
Mean	1.30	Mean	1.28							
0-0 distances in CO <sub>3</sub> gro	oups:									
$O(1) - O(3)^*$	$2.23~{ m \AA}$	$O(2) - O(4)^*$	$2.23~{ m \AA}$							
O(1) - O(5)	2.27	O(2) - O(6)	2.21							
O(3) - O(5)	2.25	O(4) - O(6)	2.23							
0-Si-O angles:										
O(7) -Si(1) - O(9)	118.1°	O(8) -Si(2) -O(10)	$120.8^\circ$							
-O(11)	116.8	-O(12)	114.2							
-O(13)	102.9	-O(13)	100.6							
O(9) -Si(1) -O(11)	108.0	O(10) - Si(2) - O(12)	108.3							
-O(13)	109.5	-O(13)	110.2							
O(11) - Si(1) - O(13)	99.7	O(12) - Si(2) - O(13)	100.3							
Ca-O distances:										
Ca(1) - O(1)	$2.367~{ m \AA}$	Ca(2) - O(1)	$2.470~{ m \AA}$							
-O(2)	2.381	-O(2)	2.436							
-O(9)	2.457	-O(4)	2.499							
-O(9')	2.387	O(7)	2.335							
-O(10)	2.383	-O(8)	2.322							
O(10')	2.450	-O(9)	2.673							
		-O(11)	2.352							
Ca(3) - O(2)	$2.399~{ m \AA}$	Ca(4) - O(4)	$2.435~{ m \AA}$							
-O(1)	2.446	-O(5)	2.313							
-O(3)	2.440	-O(6)	2.468							
-O(8)	2.616	-O(7)	2.403							
-O(7)	2.463	-O(11)	2.287							
-O(10)	2.676	-O(12)	2.322							
-O(12)	2.368									
Ca(5)-O(3)	2.299 Å									
-O(6)	2.369									
-O(5)	2.354									
-O(8)	2.344									
-O(12)	2.557									
-O(11)	2.527									

\* Shared edges.



Fig. 1. Schematic diagrams showing the dimensions of the coordination polyhedra in tilleyite. Shared edges are shown by heavy lines. For Ca(2) and Ca(3) the triangular pyramids shown by dotted lines show the change caused by including the seventh-longest Ca-O distance

six more cycles of refinement gave the following agreement factors: R = 0.11 and wR = 0.06 for all reflections.

The observed and calculated structure amplitudes are given in Table 1. Data omitted from the least-squares refinement are marked by an asterisk. Table 2 gives the atomic positional and thermal vibrational parameters, while Table 3 and Fig. 1 give the principal interatomic distances. The standard errors are about 0.007 Å for Si-O and Ca-O distances and about 0.01 Å for C-O and O-O distances.

## Discussion

Figure 2 shows the c-axis projection of tilleyite in which the  $Si_2O_7$ groups are seen end-on as triangles, and each pair of  $CO_3$  groups superimposes as a single triangle. Each  $CaO_4$  column of octahedra superimposes in projection to yield the appearance of a single octahedron viewed down a diad axis. The corrugated walls of shared  $Si_2O_7$ and  $CaO_4$  groups lie parallel to {100}. Each wall consists of  $CaO_4$ columns sharing edges with adjacent pairs of columns and sharing corners with  $Si_2O_7$  groups. The corrugated walls are cross-linked by  $CO_3$  groups sharing corners and edges with the  $CaO_4$  columns and

Crystal structure of tilleyite



Fig.2. Idealized projection of the tilleyite structure viewed down the c axis. The Si<sub>2</sub>O<sub>7</sub> groups are represented end-on by the stippled triangles. The CO<sub>3</sub> groups are shown by the broken-line triangles. The CaO<sub>6</sub> octahedra are viewed down a diad axis; the octahedra for the Ca(2) and Ca(3) atoms were obtained by omitting the seventh-longest distance shown by the dotted lines (see text). One of the corrugated sheets identified by BELOV consists of the Si<sub>2</sub>O<sub>7</sub> and CaO<sub>6</sub> octahedra marked by Si, 2, 3, 4 and 5

by a further type of  $CaO_4$  column sharing corners with the other structural units.

This description of the structure is idealized since it implicitly ignores the actual geometry. To a first approximation, all atoms in tilleyite except O(13) occur in pairs displaced by about c/2 (see Fig. 8 of SMITH, 1953). The exception, O(13), is the bridging oxygen of the Si<sub>2</sub>O<sub>7</sub> group, and has no equivalent: in addition, the pairs of Si atoms are only 3.2 Å apart compared to c/2 = 3.8 Å. All of the CaO<sub>6</sub> octahedra occur in pairs sharing edges to produce columns of CaO<sub>4</sub> composition lying parallel to c with a pseudo repeat of c/2. In addition to the pseudo-halving along c, there is a pseudo-orthorhombic geometry (but not a pseudo-orthorhombic symmetry) as shown by Fig. 10 of SMITH (1953).

There is a further idealization. The coordination polyhedra of the Ca(2) and Ca(3) atoms are shown as distorted octahedra. Actually each of these Ca atoms has seven oxygen neighbors at 2.3–2.7 Å (Table 2). Although the six atoms ascribed to each octahedron are closer than the rejected ones, all seven are sufficiently close that they must be considered as first neighbors for bonding purposes.



Fig.3. Comparison of the positions of corresponding atoms as seen in the c-axis projection. This diagram corresponds to the upper-right part of Fig.1. Outlines of mean projections of the polyhedra are shown. The dots show the projections of individual atoms. Note that O(13) does not have a corresponding atom



Fig. 4. Comparison of regular  $Si_2O_7$  and  $CaO_6$  groups (Si-O = 1.62 and Ca-O = 2.40 Å) with those found in tilleyite. See text for explanation

The detailed distortions of the structure are very difficult to depict. Figure 3 shows the relative projections of atoms with respect to the idealized drawing of Fig. 2. Figure 1 shows schematically the distortions of the polyhedra in tilleyite. Figure 4 is a simplified diagram demonstrating schematically some of the major distortions in tilleyite.

There are two simple explanations of the distortions: (a) the misfit between the idealized shapes of the  $Si_2O_7$  group and the CaO<sub>4</sub> columns of octahedra, and (b) the shortening of edges shared between polyhedra as expected from electrostatic forces between charged ions. In Fig. 4*a* the length of an idealized  $Si_2O_7$  group (4.3 Å) is considerably greater than the repeat distance (3.4 Å) of undistorted calcium octahedra. The misfit is accommodated in two ways. First of all, sharing of edges between the  $CaO_6$  groups to form  $CaO_4$  columns results in a reduction from 3.4-2.9 Å with a concomitant expansion from 3.4-3.8 Å in the perpendicular direction. Secondly, there is a flexing of the  $Si_2O_7$  group coupled with distortion of the adjacent  $CaO_6$ polyhedra. The nature of these distortions is very complex (Figs. 3 and 4b). The Ca(1) polyhedron is relatively restricted since there are centers of symmetry at the center of each edge shared with an adjacent Ca(1) polyhedron down the column. The Ca(4) and Ca(5)polyhedra are suspended between two Si<sub>2</sub>O<sub>7</sub> groups and four CO<sub>3</sub> groups, while the Ca(2) and Ca(3) polyhedra are in a very awkward environment which results in 7-fold coordination. As is discussed later, the bridging Si-O-Si angle of an Si<sub>2</sub>O<sub>7</sub> group is quite variable in crystal structures suggesting that the amount of flexing makes only a small difference to the stability of a crystal structure. The net result of all the factors is that the  $Si_2O_7$  groups flex to yield an Si-O-Si angle of 158°. The Ca(1) polyhedra undergo only a little distortion additional to that caused by sharing of edges with other polyhedra. The flexing of the  $Si_2O_7$  group causes a severe distortion between the left and right sides of the Ca(4) and Ca(5) polyhedra shown in Fig. 4b. The Ca(2) and Ca(3) polyhedra also undergo complex distortion not shown on Fig. 4b.

The flexing of the  $Si_2O_7$  group causes subtle effects on all other structural units, especially the  $CO_3$  groups, causing them to lose the c/2 relation which would otherwise be possible.

From Fig.1, it may be seen that the distortions of the  $CaO_6$  polyhedra correlate strongly with sharing or non-sharing of edges. The Ca(1) polyhedron shares edges only with other Ca polyhedra: here the effect is very clear since the shared edges of 2.98, 3.03, 3.05, 3.06, 3.09 and 3.11 Å are quite distinct from the unshared edges of 3.65, 3.63, 3.66, 3.74, 3.74, 3.78 and 3.80 Å. For Ca(4) and Ca(5), the effect is less clear since the shared edges range from 2.90–3.41 Å while the unshared edges range from 3.2-4.25 Å. For Ca(2) and Ca(3), the situation is very complex since each polyhedron has one very short edge ( $\sim 2.25$  Å) shared with a CO<sub>3</sub> group, plus the distortion caused by the seventh-near oxygen atom. Packing problems are particularly severe for these two polyhedra but again almost all the shared edges are shorter than the unshared edges.

The two CO<sub>3</sub> groups in tilleyite are equilateral triangles within experimental error. The mean C—O distance of 1.29 Å indicates, as expected, a multiple-bond character for the C—O bond. The oxygen atoms of the carbonate and the pyrosilicate groups form the corners of the Ca polyhedra. This suggests that the C—O and Si—O  $\sigma$ -bonds utilize some  $sp^n$  ( $n \leq 3$ ) hybrid orbitals of the oxygen atoms. Such hydridizations would increase the partial covalent character in some of the Ca—O bonds. Thus some short Ca—O distances (less than the sum of ionic radii, 2.39 Å) could be explained by partial homopolar bonding.

The Si-O distances in the pyrosilicate group are consistent with the  $\pi$ -bonding model proposed by CRUICKSHANK (1961). The Si-O bridge bonds are longer (1.67) than the terminal Si-O distances (1.61). The inner O-O distances (mean 2.58) are shorter than the outer distances (2.71). The O–O distances of the  $Si_2O_7$  group in four other structures are shown in Fig. 5. The inner distances of the group tend to be shorter than the outer distances. This is well defined in thortveitite (2.60 vs. 2.70) and kornerupine (2.60 vs. 2.69) but is less noticeable for zoisite (2.62 vs. 2.67) and clinozoisite (2.61 vs. 2.67). A possible explanation (SMITH, 1953) is that the distortion is caused by the mean Coulombic forces between ions composing the Si<sub>2</sub>O<sub>7</sub> group: thus the cations tend to pull inwards the outer oxygen atoms, each of which is forced to move apart from the other two members of each outer triangular group in order to maintain a near-constancy of Si-0 distances. In addition there are forces from outside the Si<sub>2</sub> $0_7$ group, for example those involving edge sharing. In CRUICKSHANK's  $\pi$ -bonding model the short inner O–O distances arise from lesser electrostatic repulsions between the bridging and the non-bridging oxygen atoms as against the stronger repulsions between the nonbridging oxygen atoms (McDonald and CRUICKSHANK, 1967).

	Mean int dista	eratomic ance	$\Delta d$ (bridge	Si-O-Si	D.C.	
Structure	Si-O (bridge)	Si—O (peripheral)	— peripheral bond length)	angle	Reference	
I. Thortveitite	1.6045(5) Å	1.628(1) Å	— 0.023(1) Å	180.0(0)°	PREWITT (personal communication)	Crystal
II. $Ba_3Si_4Nb_6O_{26}$	1.599(6)	1.629(7)	- 0.030(9)	180°	SHANNON and KATZ (1970)	struc
III. Zoisite	1.633(10)	1.620(6)	+ 0.013(10)	172.6(8)	Dollase (1968)	ture c
IV. Clinozoisite	1.628(4)	1.617(4)	+ 0.011(5)	164.3(5)	Dollase (1968)	f tille
V. Tilleyite	1.667(4)	1.610(4)	+ 0.057(6)	157.3(4)	this work	yite
VI. Kornerupine	1.65(1)	1.61(1)	+ 0.04(1)	146.2(5)	Moore and Bennett (1968)	
VII. Hardystonite	1.649(3)	1.607(3)	+ 0.042(4)	138.5(3)	Louisnathan (1969)	

## Table 4. Selected dimensions of the Si<sub>2</sub>O<sub>7</sub> group in some structures

Standard errors given in brackets to same decimal level.

There has been considerable speculation on the size of the Si-O-Si angle. The observed variation is from 180° to about 135°. LIEBAU (1961) suggested that the Si-O-Si angle must be less than 180° in conformity with an expected partial covalent character of bonding in silicates. In CRUICKSHANK's  $\pi$ -bonding model an angle near 180°



Fig.5. Schematic diagrams showing the dimensions of the coordination polyhedra in four silicates. The Si-O-Si bridge angle is also shown



Fig. 6. Correlation between the Si–O–Si bridge angle and the difference  $(\Delta d)$  in the bridge and peripheral bond lengths. Errors in  $\Delta d \left( \sqrt{\sigma_{br}^2 + \sigma_{peri}^2} \right)$  are also shown in the diagram. Numerals on the plot refer to those in Table 4

would enable the bridging oxygen to participate in both  $\pi$  systems, and the Si-O bridge bond would tend to become shorter than the terminal bonds. LAZAREV'S (1964) investigation supported this hypothesis, and he suggested that for large Si-O-Si angles the repulsion between the two silicon atoms is minimized because of increased

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electron density in the  $\pi$  bond. In Table 4 and Fig.5 are given the dimensions of seven out of many structures that contain the pyrosilicate group. The variation of the Si-O-Si angle with the difference in the lengths of the bridge bond and mean peripheral bond is shown in Fig. 6. The decrease in the difference as the Si-O-Si angle increases is consistent with the  $\pi$ -bonding model, but the correlation factor is obviously low. BROWN and GIBBS (1970) reported a similar correlation in their analysis of several silicate structures. In tilleyite the observed difference of 0.057(6) Å for an angle of 157° is somewhat large.

A straightforward explanation for the size of the Si–O–Si angle can be given in terms of packing relationships between the Si<sub>2</sub>O<sub>7</sub> group and the neighboring polyhedra. For example, the angle of 157° in tilleyite can be justified plausibly by the geometrical requirements for linkage to the columns of Ca octahedra, while the higher angle in zoisite appears to result from the spatial requirements resulting from the suspension of the Si<sub>2</sub>O<sub>7</sub> group between the columns of AlO<sub>6</sub> octahedra.

Nevertheless it would be surely wrong to believe that the flexing of bridging angle has no effect on the contribution of the  $Si_2O_7$  group to the internal energy.

We have shown that the principal distortions in the calcium-oxygen octahedra and the silicon-oxygen tetrahedra are mostly explainable in terms of PAULING'S (1960) arguments regarding chemical bonding in complex ionic structures. Within the Si<sub>2</sub>O<sub>7</sub> group the variation of individual bond lengths can be satisfactorily accounted for in terms of CRUICKSHANK'S  $\pi$ -bonding theory, but valence-balancing (ZACHARIA-SEN, 1963b) could also provide an interpretation. As has been recognized by PANT and CRUICKSHANK (1968) the two theories are not exclusive. However it must be emphasized that neither the  $\pi$ -bonding theory nor the method of valence balancing examines anything more than the first-nearest neighbors in a structure for describing the chemical bond. Such an approach is certainly far from complete in a complex structure like tillevite.

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