Crystal chemistry of macfallite: Relationships to sursassite and pumpellyite

MARIKO NAGASHIMA,1,2,* NOURI-SAID RAHMOUN,1 EVGENY V. ALEKSEEV,1 CHARLES A. GEIGER,1 THOMAS ARMBRUSTER,2 AND MASAHIKE AKASAKA3

1Institut für Geowissenschaften, Abteilung Mineralogie, Christian-Albrechts-Universität zu Kiel, D-24098, Kiel, Germany
2Mineralogical Crystallography, Institute of Geological Sciences, University of Bern, Freiestrasse 3, CH-3012, Bern, Switzerland
3Department of Geoscience, Faculty of Science and Engineering, Shimane University, Matsue 690-8504, Japan

ABSTRACT

The crystal chemistry of macfallite from Keweenaw County, Michigan was studied using electron microprobe, thermogravimetry (TG), differential thermal analysis (DTA), powder Fourier transform infrared (FTIR) spectroscopy, and single-crystal X-ray diffraction methods. The chemical formula derived from the electron-microprobe measurements is

\[(\text{Ca}_{0.03}\text{Na}_{0.01})_{32.00}\text{Mn}^{3+}_{1.51}\text{Al}_{0.27}\text{Mg}_{6.09}\text{Cu}_{0.00}\text{O}^3_{6.00}\text{O}11_{0.01})_{32.00}\text{Si}_{3.05}\text{O}10.88\text{(OH)}_{3.12}\ (Z = 2).\]

An analysis using the intensities of the MnLβ and MnLα X-ray lines shows that most Mn is trivalent. The weight loss from TG measurement is 7.7 wt% at 1000 °C, most of which is interpreted to be due to the loss of structural OH groups. The crystal structure of macfallite \([\alpha = 8.959(3), b = 6.072(2), c = 10.218(4)\ \text{Å}, \beta = 110.75(3)^\circ, \text{space group } P2_1/m,\text{ which is isostructural with sursassite, was refined using } 1717\text{ unique reflections to } R = 4.1\%.\text{ The site populations at the three independent octahedral sites, Mn1, Mn2, and Mn3, are } \text{Mn}_{0.82}\text{Al}_{0.18}\text{Mg}_{0.09}\text{Cu}_{0.03}, \text{Mn}_{0.75}\text{Al}_{0.25}, \text{and } \text{Mn}_{0.95}\text{Al}_{0.05},\text{ respectively. In agreement with a bond-valence analysis, three crystal-chemically different OH groups are located at the O6, O10, and O11 positions. The site O7 is mostly occupied by oxygen, but minor amounts of hydroxyl may be located there as well. The powder FTIR spectrum in the region of the OH-stretching vibrations is characterized by three strong bands at 3413, 3376, and 3239 cm\(^{-1}\) and an additional broad absorption band around 2900 cm\(^{-1}\). The latter results from a relatively strong hydrogen bond, O6-H…O11, with a length of ~2.63 Å. Although there are three main hydroxyl groups occurring in macfallite, the exact number depends on the concentration of trivalent and divalent cations at the Mn1 site. If divalent cations occur at Mn1, a fourth OH group is necessary to maintain charge balance.}

Keywords: Macfallite, pumpellyite, sursassite, Mn silicates, hydrous silicates, FTIR spectroscopy, crystal structure

INTRODUCTION

Low-temperature silicates, many of which are hydrous, are often complex both structurally and compositionally. They often occur as fine-grained aggregates, and it is difficult to obtain good single crystals and/or sufficient amounts of material for experimental investigation. Macfallite is such a relatively low-temperature hydrous silicate with the ideal formula \(\text{Ca}_2\text{Mn}^3_+\text{Si}_6\text{O}_{15}\text{(OH)}_8.\) It has been reported from three localities: (1) fissures and lenses in Keweenaw basalt near Manganese Lake, Copper Harbor, Keweenaw County, Michigan (Moore et al. 1979); (2) metamorphosed and metasomatized manganese deposits in metacherts from the Cerchiara mine, eastern Liguria, Italy (Basso et al. 1989); and (3) low-grade metamorphic rocks from the Wakasa manganese mine, Hokkaido, Japan (Miyajima et al. 1998).

The crystal structure of macfallite was first determined by Moore et al. (1985) who proposed the general structural formula \(\text{Ca}_1\text{Ca}_2\text{Mn}_1\text{Mn}_2\text{Z}_3\text{O}_{15}\text{(OH)}_8,\) where Ca1 and Ca2 represent seven coordinated sites, Mn1, Mn2, and Mn3, octahedral sites, and Z, tetrahedral sites. However, the original structure determination \((R = 18.4\%)\) allowed only a description of the average crystal structure and not an examination of detailed crystal-chemical properties. Although little is known about the compositional variations in natural crystals, macfallite is compositionally similar to okhotskite \((\text{Mn}^{3+}-\text{analog of pumpellyite})\) of composition \(\text{Ca}_a(\text{Mn}^{2+}\text{Mg})_b(\text{Mn}^{3+}\text{AlFe}^{3+})_c\text{Si}_{2z}\text{O}_{56-4\alpha}\text{(OH)}_n\) (Togari and Akasaka 1987). A polysomatic description for isostructural macfallite and sursassite, as well as pumpellyite, ardenite and lawsonite, the so-called 6 × 9 Å structures with corresponding axial translations, was presented by Moore et al. (1985) and Ferraris et al. (1986).

In this study, we investigated, further and in detail, the crystal structure and crystal-chemical properties of macfallite. The following experimental methods were used: electron microprobe analysis, thermogravimetry (TG), differential thermal analysis (DTA), powder Fourier transform infrared spectroscopy (FTIR), and single-crystal X-ray diffraction. A bond-valence analysis was also carried out. Finally, we examined the structural and compositional relationships between macfallite, pumpellyite, and sursassite including the nature of their H-bonding systems.
EXPERIMENTAL METHODS

Specimen
The studied sample was collected from veins including macfallite, calcite, orientite, and brunitite that occur in basalt located near Mangarese Lake, Copper Harbor, in Keweenaw County, Michigan (Moore et al. 1979). Macfallite occurs as needle- and prism-like crystals that are elongated along the b axis and are up to 3 mm in length. The crystals are dark reddish-brown, and are pleochroic from reddish-brown to yellow in polarized light. Crystals are frequently twinned with {001} as a twin plane.

Chemical analysis
The chemical composition of macfallite was determined using a JEOL JXA-8900R electron probe microanalyzer at the University of Kiel. The abundances of Si, Ti, Al, Cr, V, Fe, Mn, Cu, Mg, Ca, Sr, Ba, K, and Na were measured using an accelerating voltage of 20 kV and a beam current of 20 nA, with a beam diameter of 1 μm. The following standards were used: wollastonite (Si), rutile (Ti), anorthite (Al, Ca), chromite (Cr), vanadium metal (V), fayalite (Fe), rhodonite (Mn), elemental copper (Cu), forsterite (Mg), a Ge-Ca-Al glass containing 10 wt% Sr (Sr), a Sr-Ge-Ca glass containing 10 wt% Ba (Ba), natural mullite (K), and anorthoclase (Na). The PKZ method (modified ZAF) was used for data correction for all elements.

Semi-quantitative Mn\(^{3+}\)/Mn\(^{2+}\) ratios in macfallite were determined using the ratio of the X-ray intensities of the Mn\(\beta\) and Mn\(\alpha\) lines, following the methods of Albee and Chodos (1970) and Kimura and Akasaka (1999). The X-ray intensities of the Mn\(\beta\) and Mn\(\alpha\) lines were measured using a JEOL JXA-8800M electron-probe microanalyzer at Shimane University, Japan. The measurements were done using an accelerating voltage of 15 kV; a beam current of 30 nA and beam diameter of 10 μm. Emission spectra of the Mn\(\beta\) and Mn\(\alpha\) lines were measured using a TAP monochromator crystal for a range of L-values between 206 and 213 mm with a step interval of 0.05 mm and a step counting time of 10 s. The L values are the distances (in millimeters) between the measuring spot on the sample and the point on the monochromator crystal where the X-ray beam impinges, and are related through the equation \(L = (Rd) \times n\), to the wavelength of X-rays (\(\lambda\) in Å), the radius of the Roland circle (\(R = 140\) mm for the JEOL JXA-8800M), the (001) \(d\)-spacings (\(\lambda\)) of the monochromator crystal and the order of Bragg reflection (\(n\)).

The measured spectra were fitted with Lorentzian curves by least-squares methods written by Kimura and Akasaka (1999). The equation, \(I(Mn\(\beta\))/I(Mn\(\alpha\)) = 0.46 \times Mn\(^{2+}\)/(Mn\(^{3+}\) + Mn\(^{2+}\)) + 0.554 (Kimura and Akasaka 1999), where \(I\) represents intensity, was employed for the estimation of the Mn\(^{3+}\)/Mn\(^{2+}\) ratio.

DTA/TG measurement
DTA/TG measurements were carried using a SHIMADZU DTG-60H instrument on 17.15 mg of material from room temperature to 1000 °C under a flow of N\(_2\) gas with a heating rate of 10 °C/min.

FTIR spectroscopy
The powder FTIR spectrum was measured with a BRUKER IFS/66v FTIR spectrometer, using the KBr pellet technique, in the wave region between 4000 and 3500 cm\(^{-1}\). A pellet of 13 mm diameter was prepared by pressing a mixture of 1 mg of powdered macfallite and 200 mg of dried KBr under vacuum. The pellet was kept in a vacuum desiccator for 24 h at 150 °C to minimize adsorbed water. The spectrum was recorded under vacuum at a resolution of 2 cm\(^{-1}\) by combining 512 scans. Spectral analysis, including baseline adjustment and peak fitting, was performed using the OPUS software package (Bruker).

Single-crystal structure analysis
X-ray diffraction data for a single crystal of macfallite were collected using a STOE STADI4 four-circle diffractometer. A prismatic crystal of size 0.32 × 0.07 × 0.06 mm was mounted on a glass fiber and intensity data were measured at room temperature using graphite monochromated MoK\(_\alpha\) radiation (\(\lambda = 0.71069\) Å). The reflection statistics and systematic absences were consistent with space groups P2\(_1\)/m and P2\(_1\)\(2\_1\)\(_1\)/m. Subsequent attempts to solve the structure indicated that the observed average structure is centrosymmetric and for this reason P2\(_1\)/m is the correct space group. Structure refinement was performed using the SHELXL-97 program (Sheldrick 1997). Scattering factors for neutral atoms were employed. Positions of the hydrogen atoms of the hydroxyl groups were derived from difference Fourier synthesis. Subsequently, hydrogen positions were refined assuming full occupancy at a fixed value of U\(_h\) = 0.05 Å\(^2\). The site populations for the Ca1, Ca2, Mn1, Mn2, and Mn3 were determined in accord with Hawthorne et al. (1995). The hydrogen positions were refined with a bond distance constraint of O-H = 0.98 Å (Franks 1973). Atom labeling, except for H sites, follows the nomenclature for iso-structural surassite (Mellini et al. 1984).

RESULTS

Chemical composition of macfallite
Macfallite crystals typically show compositional zoning with Al-rich cores and Mn-rich rims, and the total Mn\(_2\)O\(_3\) concentration varies between 28.0 and 39.4 wt%. The average chemical composition of our macfallite sample, based on 46 point analyses, is given in Table 1. The corresponding crystal-chemical formula is

\[(\text{Ca}_{2.03}\text{Na}_{0.01})_{22.04}\text{(Mn}_{1.51}\text{Al}_{0.12}\text{Mg}_{0.09}\text{Cu}_{0.03}\text{V}_{0.03})_{22.0}\text{Si}_{3.05}\text{O}_{8.68}\text{(OH)}_{3.12}\]

where the total number of cations, except H ions, was normalized to 8 and the amount of OH was calculated using charge balance considerations. The macfallite crystal used for the X-ray single-crystal refinement is from the same hand specimen, and the average chemical composition in Table 1 was, therefore, used to constrain Mg and Cu contents in the site population refinement.

DTA/TG measurement
The DTA/TG results are shown in Figure 1. The TG trace shows weight loss starting around 500 °C. The measured weight loss for the sample at the final temperature of 1000 °C corresponds to 7.7 wt%, most of which is related to structural OH groups. The TG results indicate that no loosely held molecular H\(_2\)O is present in macfallite. The change in the DTA signal starting around 300 °C is mainly related to increased atomic vibrations, much of it related to the OH groups followed by eventual dehydroxylation of the structure. At roughly 750 °C, the macfallite structure breaks down and various solid-state reactions begin to occur.

Infrared spectrum
The FTIR powder spectrum between 4000 and 1500 cm\(^{-1}\) is shown in Figure 2. The strongest absorption bands are located in the range between 3600 and 2700 cm\(^{-1}\). Three intense bands at 3413, 3376, and 3239 cm\(^{-1}\) and an additional broad absorption feature at about 2900 cm\(^{-1}\) were assigned to OH vibrations.

| Table 1. Composition of macfallite based on 46 point analyses |
|---------------|---------------|---------------|----------------|---------------|
| Oxide         | wt%           | s.d.          | Element        | apfu*         |
| SiO\(_2\)      | 33.14         | 0.41          | Si             | 3.05          | 0.03          |
| AlO\(_2\)      | 2.49          | 1.74          | Al             | 0.27          | 0.18          |
| V\(_2\)O\(_5\)† | 0.09          | 0.09          | V\(^{3+}\)     | 0.01          | 0.01          |
| FeO\(_2\)†     | 0.05          | 0.02          | Fe\(^{3+}\)    | 0.00          | 0.00          |
| MnO\(_2\)†    | 35.80         | 2.96          | Mn\(^{2+}\)    | 2.51          | 0.23          |
| CuO†          | 0.45          | 0.31          | Cu\(^{2+}\)    | 0.03          | 0.02          |
| MgO†          | 0.69          | 0.51          | Mg             | 0.09          | 0.07          |
| CaO†          | 20.53         | 0.31          | Ca             | 2.03          | 0.02          |
| SrO†          | 0.07          | 0.01          | Sr             | 0.00          | 0.00          |
| NaO\(_2\)      | 0.03          | 0.02          | Na             | 0.01          | 0.00          |
| Total         | 93.34         |               |                | 8.00          |               |

* Total cations = 8.
† Total V, Fe, Mn, and Cu as V\(_2\)O\(_5\), FeO\(_2\), MnO\(_2\), and CuO, respectively.
Crystal structural solution and refinements

Crystallographic data and refinement parameters are summarized in Table 2. The refined atomic positions and anisotropic displacement parameters are listed in Table 3. Interatomic distances and angles are listed in Table 4. Table 5 gives the site-scattering values and cation assignments. The crystal structure of macfallite is shown in Figure 3.

Both electron-microprobe data and the site-scattering values indicate that the Ca1 and Ca2 sites are occupied only by Ca. The site populations for Mn\(^{2+}\) and Al at the Mn1, Mn2, and Mn3 sites were calculated by use of the following procedure: (1) elements with less than 0.01 atoms per formula unit (apfu) were omitted;

### Table 2. Crystallographic data for macfallite and experimental details for the single-crystal X-ray diffraction analysis

<table>
<thead>
<tr>
<th>Parameter</th>
<th>Value</th>
</tr>
</thead>
<tbody>
<tr>
<td>Crystal size (mm)</td>
<td>0.32 x 0.07 x 0.06</td>
</tr>
<tr>
<td>Space group</td>
<td>(P2_1/m)</td>
</tr>
<tr>
<td>Cell parameters</td>
<td>(a (\AA)): 8.959(3)</td>
</tr>
<tr>
<td></td>
<td>(b (\AA)): 6.072(2)</td>
</tr>
<tr>
<td></td>
<td>(c (\AA)): 10.218(4)</td>
</tr>
<tr>
<td></td>
<td>(\beta (^\circ)): 110.75(3)</td>
</tr>
<tr>
<td>V (\AA(^3))</td>
<td>519.86(6)</td>
</tr>
<tr>
<td>Density (g/cm(^3))</td>
<td>3.454</td>
</tr>
<tr>
<td>Radiation</td>
<td>MoK(_\alpha)</td>
</tr>
<tr>
<td>Monochromator</td>
<td>Graphite</td>
</tr>
<tr>
<td>Diffractometer</td>
<td>STOE STADI4</td>
</tr>
<tr>
<td>Scan type</td>
<td>(\omega) scans</td>
</tr>
<tr>
<td>Refinement system used</td>
<td>SHELXL97 (Sheldrick 1997)</td>
</tr>
<tr>
<td>Refinement on</td>
<td>(F^2)</td>
</tr>
<tr>
<td>(R_e) (%)</td>
<td>4.12</td>
</tr>
<tr>
<td>(wR_F) (%)</td>
<td>9.41</td>
</tr>
<tr>
<td>No. of parameters</td>
<td>136</td>
</tr>
<tr>
<td>Weighting scheme</td>
<td>(w = 1/(\sigma^2(F^2) + (0.0490P)^2))</td>
</tr>
<tr>
<td>(</td>
<td>\Delta F</td>
</tr>
<tr>
<td>(D_{max}) (e Å(^{-2}))</td>
<td>0.958</td>
</tr>
<tr>
<td>(\Delta D_{max}) (e Å(^{-2}))</td>
<td>-0.913</td>
</tr>
</tbody>
</table>

### Table 3. Refined atomic positions, isotropic, and anisotropic displacement parameters

<table>
<thead>
<tr>
<th>Site</th>
<th>neq*</th>
<th>W†</th>
<th>x</th>
<th>y</th>
<th>z</th>
<th>U(_{11})</th>
<th>U(_{22})</th>
<th>U(_{33})</th>
<th>U(_{12})</th>
<th>U(_{13})</th>
<th>U(_{23})</th>
<th>U(_{13})</th>
</tr>
</thead>
<tbody>
<tr>
<td>Ca1</td>
<td>2 e</td>
<td>0.1670(1)</td>
<td>1/4</td>
<td>0.3134(1)</td>
<td>0.0076(4)</td>
<td>0.0116(5)</td>
<td>0.0084(4)</td>
<td>0.0018(3)</td>
<td>0.0095(3)</td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>Ca2</td>
<td>2 e</td>
<td>0.2949(1)</td>
<td>1/4</td>
<td>0.6813(1)</td>
<td>0.0119(4)</td>
<td>0.0105(5)</td>
<td>0.0081(4)</td>
<td>0.0057(3)</td>
<td>0.0096(2)</td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>Si1</td>
<td>2 e</td>
<td>0.3008(1)</td>
<td>3/4</td>
<td>0.1891(1)</td>
<td>0.0056(5)</td>
<td>0.0146(8)</td>
<td>0.0087(5)</td>
<td>0.0034(4)</td>
<td>0.0078(3)</td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>Si2</td>
<td>2 e</td>
<td>0.2074(1)</td>
<td>3/4</td>
<td>0.8040(1)</td>
<td>0.0052(5)</td>
<td>0.0146(6)</td>
<td>0.0087(5)</td>
<td>0.0034(4)</td>
<td>0.0078(3)</td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>Si3</td>
<td>2 e</td>
<td>0.1603(1)</td>
<td>3/4</td>
<td>0.0482(1)</td>
<td>0.0081(5)</td>
<td>0.0074(6)</td>
<td>0.0069(5)</td>
<td>0.0040(4)</td>
<td>0.0071(3)</td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>Mn1</td>
<td>2 b</td>
<td>0.6869(1)</td>
<td>1/2</td>
<td>0.4453(1)</td>
<td>0.0047(3)</td>
<td>0.0070(4)</td>
<td>0.0062(3)</td>
<td>0.0005(3)</td>
<td>0.0029(2)</td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>Mn2</td>
<td>2 a</td>
<td>0.5591(1)</td>
<td>1/2</td>
<td>0.2268(1)</td>
<td>0.0048(3)</td>
<td>0.0070(4)</td>
<td>0.0062(3)</td>
<td>0.0005(3)</td>
<td>0.0029(2)</td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>O1</td>
<td>4 f</td>
<td>0.1683(1)</td>
<td>3/4</td>
<td>0.3448(1)</td>
<td>0.0071(7)</td>
<td>0.0122(2)</td>
<td>0.0032(1)</td>
<td>0.0003(1)</td>
<td>0.0112(7)</td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>O2</td>
<td>4 f</td>
<td>0.1937(1)</td>
<td>3/4</td>
<td>0.3808(1)</td>
<td>0.0061(7)</td>
<td>0.0122(2)</td>
<td>0.0032(1)</td>
<td>0.0003(1)</td>
<td>0.0112(7)</td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>O3</td>
<td>4 f</td>
<td>0.3153(1)</td>
<td>3/4</td>
<td>0.8358(1)</td>
<td>0.0080(8)</td>
<td>0.0122(2)</td>
<td>0.0032(1)</td>
<td>0.0003(1)</td>
<td>0.0112(7)</td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>O4</td>
<td>2 e</td>
<td>0.4213(1)</td>
<td>3/4</td>
<td>0.0929(4)</td>
<td>0.0061(7)</td>
<td>0.0122(2)</td>
<td>0.0032(1)</td>
<td>0.0003(1)</td>
<td>0.0112(7)</td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>O5</td>
<td>2 e</td>
<td>0.4444(1)</td>
<td>3/4</td>
<td>0.3485(1)</td>
<td>0.0071(7)</td>
<td>0.0122(2)</td>
<td>0.0032(1)</td>
<td>0.0003(1)</td>
<td>0.0112(7)</td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>O6</td>
<td>2 e</td>
<td>0.0866(1)</td>
<td>1/4</td>
<td>0.9438(1)</td>
<td>0.0011(1)</td>
<td>0.0082(1)</td>
<td>0.0141(1)</td>
<td>0.0007(1)</td>
<td>0.0101(7)</td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>O7</td>
<td>2 e</td>
<td>0.4395(1)</td>
<td>1/4</td>
<td>0.3806(1)</td>
<td>0.0061(7)</td>
<td>0.0122(2)</td>
<td>0.0032(1)</td>
<td>0.0003(1)</td>
<td>0.0112(7)</td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>O8</td>
<td>2 e</td>
<td>0.0726(1)</td>
<td>3/4</td>
<td>0.8761(1)</td>
<td>0.0071(1)</td>
<td>0.0152(1)</td>
<td>0.0131(1)</td>
<td>0.0061(1)</td>
<td>0.0106(7)</td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>O9</td>
<td>2 e</td>
<td>0.0999(1)</td>
<td>3/4</td>
<td>0.6326(1)</td>
<td>0.0081(1)</td>
<td>0.0142(1)</td>
<td>0.0091(1)</td>
<td>0.0051(1)</td>
<td>0.0099(7)</td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>O10</td>
<td>2 e</td>
<td>-0.0071(1)</td>
<td>3/4</td>
<td>0.3610(1)</td>
<td>0.0121(1)</td>
<td>0.0212(1)</td>
<td>0.0031(1)</td>
<td>0.0041(1)</td>
<td>0.0120(8)</td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>O11</td>
<td>2 e</td>
<td>0.3983(1)</td>
<td>1/4</td>
<td>0.0631(2)</td>
<td>0.0071(1)</td>
<td>0.0252(2)</td>
<td>0.0192(1)</td>
<td>0.0061(1)</td>
<td>0.0168(8)</td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>H6</td>
<td>2 e</td>
<td>0.2032(1)</td>
<td>3/4</td>
<td>0.9784(1)</td>
<td>0.005</td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>H10</td>
<td>2 e</td>
<td>-0.0021(1)</td>
<td>3/4</td>
<td>0.2666(4)</td>
<td>0.05</td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>H11</td>
<td>2 e</td>
<td>0.4026(1)</td>
<td>1/4</td>
<td>0.1609(4)</td>
<td>0.05</td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
</tbody>
</table>

* Multiplicity.  
† Wyckoff letter.
The site-scattering values were determined from the site occupancies of calcium at Ca1 and Ca2 and manganese at Mn1, Mn2, and Mn3, refined without any constraints.

**Cation distributions at the octahedral sites**

Moore et al. (1985) determined site occupancies for the Mn2 site in macfallite as 0.61(2) Mn$^{2+}$ + 0.39 Al$^{3+}$, whereas the Mn1 and Mn3 sites were assumed to be completely filled with Mn$^{2+}$. The presence of Al at the Mn2 site is confirmed by the results obtained in this study giving 0.75 Mn$^{2+}$ + 0.25 Al$^{3+}$. However, our results indicate that a small amount of Al may also be assigned to Mn1 and Mn3 with 0.06 and 0.05 apfu Al, respectively. In addition, contrary to the proposal that the O7 position consists of an OH group (Moore et al. 1985), both oxygen atoms and OH groups can be assigned to this site (see below).

**DISCUSSION**

Crystal chemical relationships between macfallite, sursassite, and pumpellyite

The crystal structures of macfallite, sursassite, and pumpellyite ($Z = 1$, space group $P2_1/m$) of formula $\text{Mn}_3^2\text{Si}_3\text{O}_9\text{(OH)}$, and pumpellyite ($Z = 1$, space group $P2_1/m$) are closely related to each other (Allmann 1984; Mellini et al. 1984; Moore et al. 1985; Ferraris et al. 1986). Moore et al. (1985) analyzed the crystal chemistry of the group of so-called $6 \times 9$ Å sheet structures, namely lawsonite, sursassite, macfallite, and pumpellyite. All their structures contain a $\text{Mn}_3^2\text{Si}_3\text{O}_9\text{(OH)}$ sheet. In projections along their 6 Å axes (i.e., the $b$ axis in macfallite), these structures show two different infinite chains consisting of edge-sharing $\text{MnO}_6$ octahedra (labeled M2 and M3 in macfallite and sursassite and Y in pumpellyite). They are linked to one another along the 9 Å axis (i.e., $a$ axis in macfallite) by $\text{TO}_4$ tetrahedra (i.e., $\text{Si}_1$ and $\text{Si}_2$ in macfallite). The vacancy symbol $\square$ in the structural unit represents a third “empty structural column” running parallel to the former two chains of octahedra (Fig. 4). It is represented in projections onto (010) as narrow channels having a rhombohedral cross section.
Divalent cations are located at the Mn1 site in macfallite and at the X site in pumpellyite. The volume of the Mn1O6 octahedron (11.3 Å³) in macfallite is similar to that of the XO6 octahedron in pumpellyite (10.3–12.0 Å³; Galli and Alberti 1969; Allmann and Donnay 1973; Yoshihisa and Matsumoto 1985; Artioli and Geiger 1994; Artioli et al. 2003; Brigatti et al. 2006; Nagashima et al. 2006; Hatert et al. 2007; Nagashima and Akasaka 2007). In the case of sursassite, previous studies (e.g., Heinrich 1962; Allmann 1984; Mellini et al. 1984; Reinecke 1986; Minakawa and Momoi 1987) implied that small amounts of divalent cations, such as Mg, are distributed over the octahedral sites, although the possibility of intergrowths with pumpellyite (Mellini et al. 1984) among its three octahedral sites. On the basis of structural topology, M1 in macfallite (i.e., Mn1) and sursassite (A1) are the most closely related to the X site in pumpellyite, whereas M2 and M3 in macfallite (i.e., Mn2 and Mn3) and sursassite (A2 and A13) are related to the Y site in pumpellyite. This assignment is also supported by the cation distribution over the octahedral sites and the size of the MnO6 octahedra.

It should be noted, however, that the angular distortion for the Mn1O6 octahedron in macfallite is greater than that for the Mn2O6 and Mn3O6 octahedra (Table 7), which is opposite to the angular distortion found for corresponding octahedra in pumpellyite and sursassite. In addition, the degree of octahedral bond length distortion in macfallite is more pronounced than that in sursassite (Allmann 1984; Mellini et al. 1984). The reason may be related to a Jahn-Teller effect arising from the presence of Mn³⁺ in macfallite.

**Hydrogen bond systems in macfallite and sursassite**

The positions of the hydrogen atoms and directions of the hydrogen bonds in macfallite are shown in Figure 4. All O-H bonds are located in mirror planes that are parallel to (010). This is also the situation in pumpellyite (Yoshihisa and Matsumoto 1985; Brigatti et al. 2006; Nagashima and Akasaka 2007). The relationships between donor and acceptor O atoms and their hydrogen-bonding in macfallite can be summarized as follows: O10-H10···O6, O6-H6···O11, O11-H11···O2, and O11-H11···O7. The O10H10 hydroxyl group terminates the disilicate unit and represents a so-called silanol group. O6 acts as an acceptor of this weak hydrogen bond (Fig. 4). Following this, the general structural formula of macfallite can be presented as Ca3(Mn³⁺, Al)3[OH]3[SiO₄][SiO₃(OH)]₃.

The IR spectrum in the region of the OH-stretching vibrations is characterized by three OH bands at 3413, 3376, and 3239 cm⁻¹, and one broad band at 2900 cm⁻¹. The associated OH···O distances (Table 8) were calculated using the correlation between the observed OH-band wave number and the O···O separation in Å as given by Libowitzky (1999). The three strongest OH bands result from relatively weak hydrogen bonds associated with the O10 and O11 sites, in agreement with the H positions determined by X-ray and the bond valence sum analysis. The broad band at ca. 2900 cm⁻¹ is assigned to an OH stretch associated with the O6-H6···O11 bond with an O···O separation of about 2.63 Å. In macfallite, O6 and O11 (6-3 type sites) occur on the channel walls and they contain OH groups that compete to form a hydrogen bond toward the diagonally adjacent site. The resulting hydrogen bonding occurs inside the narrow channels, and it gives rise to the band at ca. 2900 cm⁻¹. The O6 and O11 site are associated with bond-valence values of 1.25 and 0.99 valence units (v.u.), respectively. This is due to development of Mn³⁺ Jahn-Teller distortion in the columns formed by Mn2 and Mn1 octahedra. Because the bond-valence value for O11 is less than that of O6 (Table 6), O11 can act as both a donor and an acceptor for a stronger hydrogen bond, whereas O6 acts as a donor and acceptor for a weaker hydrogen bond. Ferraris and Ivaldi (1988) showed that for short O-H···O donor-acceptor distances, the acceptor benefits from an additional ca. 0.25 v.u. Thus, O11, with the smaller bond-valence value, acts as an acceptor from O6-H6 to form a relatively strong hydrogen bond. The O11-H11 group has to adjust to this hydrogen bonding arrangement. This gives rise to a bifurcated hydrogen bond, where H11 is bonded to both O2 and O7. Similar broad IR bands at low wave number are present in pumpellyite.

![Figure 4](image-url)
also observed in sursassite (Reddy and Frost 2007), pumpellyite (Hatert et al. 2007), and lawsonite (Libowitzky and Rossman 1996). The origin of their OH bands is similar. These minerals also have relatively strong hydrogen bonds that cross narrow channels in their crystal structures.

Because of Mn\(^{2+}\) Jahn-Teller distortion, isostructural mozartite CaMn\(^{3+}\)SiO\(_4\)(OH) and vuagnatite CaAlSiO\(_4\)(OH) show a different arrangement of hydrogen-bond acceptors and donors (Nyfeler et al. 1997). Thus, it is not clear whether isostructural macfallite (M: Mn\(^{3+}\)) and sursassite (M: Al) have an identical system of hydrogen bonds. Mellini et al. (1984) suggested that in sursassite two different hydrogen-bond systems are present in equal proportions namely: O7·H3···O11·H2···O6·H1···O10\(^{3+}\) (Fig. 5a) and O7···H3′·O11·H2′·O6···H1′·O10\(^{3+}\) (Fig. 5b). In this model, the O6 and O11 sites always contain hydroxyl groups, whereas O7 and O10 can be occupied by either hydroxyl or oxygen anions on a local scale. In our reinterpretation of the hydrogen-bond system in sursassite, we propose that the occurrence of a specific set of hydrogen bonds depends on the degree of M\(^{2+}\) substitution at Al1. With only M\(^{3+}\) cations on Al1 there are only three hydrogen bonds labeled as H1′, H2′, and H3′ positions.

### Table 7. Angular and bond length distortions for the octahedral sites of macfallite and structurally related minerals

<table>
<thead>
<tr>
<th>Mineral</th>
<th>Mn1 site</th>
<th>Mn2 site</th>
<th>Mn3 site</th>
<th>Reference</th>
</tr>
</thead>
<tbody>
<tr>
<td>Macfallite</td>
<td>33.60</td>
<td>26.65</td>
<td>19.13</td>
<td>This study</td>
</tr>
<tr>
<td>Macfallite</td>
<td>31.58</td>
<td>26.39</td>
<td>18.04</td>
<td>Moore et al. (1985)</td>
</tr>
<tr>
<td>Sursassite</td>
<td>14.16</td>
<td>37.08</td>
<td>26.31</td>
<td>Allmann (1984)</td>
</tr>
<tr>
<td>Pumpellyite group</td>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>HR</td>
<td>21.08</td>
<td>31.63</td>
<td>30.16</td>
<td>Galli and Alberti (1969)</td>
</tr>
<tr>
<td>Julgoldite</td>
<td>14.55</td>
<td>29.59</td>
<td>16.30</td>
<td>Allmann and Donnay (1973)</td>
</tr>
<tr>
<td>Pumpellyite</td>
<td>18.83</td>
<td>37.46</td>
<td>30.16</td>
<td>Yoshiasa and Matsumoto (1985)</td>
</tr>
<tr>
<td>K1</td>
<td>26.64</td>
<td>34.01</td>
<td>31.63</td>
<td>Artioli and Geiger (1994)</td>
</tr>
<tr>
<td>BU</td>
<td>19.79</td>
<td>30.16</td>
<td>31.63</td>
<td>*</td>
</tr>
<tr>
<td>Julgoldite-(Fe(^{3+}))</td>
<td>5.98</td>
<td>16.30</td>
<td>31.63</td>
<td>Artioli et al. (2003)</td>
</tr>
<tr>
<td>MTS</td>
<td>14.41</td>
<td>36.39</td>
<td>31.63</td>
<td>*</td>
</tr>
<tr>
<td>KGH</td>
<td>18.41</td>
<td>35.25</td>
<td>31.63</td>
<td>*</td>
</tr>
<tr>
<td>Popplite</td>
<td>18.74</td>
<td>29.40</td>
<td>29.40</td>
<td>Brigatti et al. (2006)</td>
</tr>
<tr>
<td>Pumpellyite-(Al)</td>
<td>12.07</td>
<td>29.69</td>
<td>29.69</td>
<td>Hatert et al. (2007)</td>
</tr>
<tr>
<td>Chromian pumpellyite</td>
<td>13.56</td>
<td>37.69</td>
<td>37.69</td>
<td>Nagashima and Akasaka (2007)</td>
</tr>
</tbody>
</table>

Note: HR, K1, and BU are sample labels given by Galli and Alberti (1969) and Artioli and Geiger (1994). $\alpha_{\text{oct}}^{(i)} = \sum (l_i - l_j^2) / 6$, $l_i =$ center-to-vertex distance for an octahedron with $O_i$ symmetry, whose volume is equal to that of a distorted octahedron with bond lengths $l_i$ (Robinson et al. 1971), $D(\text{oct}) = 1 / 6 \sum R_i - \sigma_{\text{oct}}^{(i)} \sigma_{\text{oct}}^{(i)}$ (each bond length, $R_i =$ average distance for an octahedron) (Baur 1974), and $\sigma_{\text{oct}}^{(i)} = \sum (\theta_i - 90°) / 11$ (Robinson et al. 1971).

### Table 8. Hydroxyl stretching wave number and corresponding O–O distance in macfallite, as determined using the correlation from Libowitzky (1999)

<table>
<thead>
<tr>
<th>OH stretching wave number (cm(^{-1}))</th>
<th>d (O–O) (Å)</th>
</tr>
</thead>
<tbody>
<tr>
<td>3412</td>
<td>2.81</td>
</tr>
<tr>
<td>3376</td>
<td>2.78</td>
</tr>
<tr>
<td>3239</td>
<td>2.72</td>
</tr>
<tr>
<td>2900</td>
<td>2.63</td>
</tr>
</tbody>
</table>

Figure 5. Hydrogen-bond system in sursassite proposed by Mellini et al. (1984); (a) with H1, H2, and H3 positions; (b) with H1′, H2′, and H3′ positions.
of Allmann (1984) and Mellini et al. (1984), suggest that Mg occurs preferentially at Al1. This requires, for charge balance reasons, an OH group at O7, with O11 acting as the acceptor of the hydrogen bond. Therefore, the OH group at O11 cannot form a hydrogen bond to O7 as an acceptor (Fig. 5b), but instead must reorient toward the center of the narrow structural channels with O6 as an acceptor (Fig. 5a). Subsequently, the hydrogen bond associated with O6 as a donor must also be different from the corresponding one in macfallite. Mellini et al. (1984) suggest that O6-H1 develops a strong hydrogen bond to O10, the terminal anion of the SiO6 unit. Finally, O10 also represents an OH group and with O6 probably acting as an acceptor. However, the exact orientation of this bond is uncertain, because the H position of Mellini et al. (1984) is located only 1.58 Å from Mn2, which is very short indeed. Thus, a restudy of the hydrogen-bond system in sursasinate is mandatory.

**ACKNOWLEDGMENTS**

We thank P. Appel and B. Mader for their help with the EPMA analysis, H. Euerl for the TG/DTA measurements, U. Cornelissen for the FTIR spectroscopy, P. Kluge for technical assistance, and K. Makino of Shinsyu University for his valuable advice concerning single-crystal structural analysis and determination of site populations. We also thank D. Gatta, associate editor, A. Locock technical editor, M. Mellini, and two anonymous reviewers for their constructive comments on this manuscript. This research was supported by an Alexander von Humboldt Foundation Fellowship to M. Nagashima.

**REFERENCES CITED**


MANUSCRIPT RECEIVED JANUARY 18, 2008
MANUSCRIPT ACCEPTED MAY 20, 2008
MANUSCRIPT HANDLED BY G. DIEGO GATTA