

Structure and Stability of Carboxylate Complexes. Part I. The Crystal and Molecular Structures of Copper(II) Glycollate, DL-Lactate, 2-Hydroxy-2-methylpropionate, Methoxyacetate, and Phenoxyacetate

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The crystal structures of bis(glycollato)copper(II), aquobis-(DL-lactato)copper(II) hemihydrate, diaquobis-(2-hydroxy-2-methylpropionate)copper(II), diaquobis(methoxyacetato)copper(II), and diaquobis(phenoxyacetato)copper(II) have been determined by three-dimensional X-ray structure analysis. All the complexes contain *trans*-chelate rings, which, apart from the bow-shaped lactate, are coplanar. The copper ions are in elongated tetragonally distorted octahedral environments with two exceptions: in the lactate, copper ions are in a nearly square-pyramidal five-co-ordinate environment; in the methoxyacetate, the octahedron is compressed. Details of the structures may be rationalised in terms of Jahn-Teller distortions and hydrogen-bonding.

Most bis(monocarboxylato)copper(II) complexes in which the ligand is unequivocally unidentate have low magnetic moments (~ 1.45 B.M. at room temperature) in the solid state and in some non-aqueous solvents,¹ and may be assumed to have dimeric or polymeric structures. However, in aqueous solution the complexes are mononuclear,² up to a total copper ion concentration of at least 0.1M. The known crystal structures of aminocarboxylatocuppper(II) complexes³ indicate that 2- and 3-aminocarboxylates are invariably bidentate ligands in the solid state. The high thermodynamic stability and hydration⁴ of these copper complexes give little reason to doubt that they are also chelates in aqueous solution. However, for 2- and 3-hydroxy-, -alkoxy-, and -aryloxy-carboxylates the situation is less clear. No crystallographic information was available when the present programme was initiated. Copper complexes of the 2-substituted carboxylates $R^1R^2C(OR^3)CO_2^-$ have magnetic moments⁵ of about 1.9 B.M. and are therefore monomeric, though not necessarily chelates. However, complexes of the analogous 3-substituted ligands have low magnetic moments,⁵ indicative of the dimeric copper acetate structure.⁶ Systematic studies of differential proton relaxation⁷ in bound ligands indicate that 2- and 3-hydroxycarboxylates and 2-alkoxycarboxylates form chelates with copper(II) in aqueous solution, but that the 3-alkoxycarboxylates are unidentate ligands. Infrared data in heavy-water solution,⁸ and thermodynamic data² are consistent with these deductions. There is a fine balance between the different structures of the complex and between the different functions of the ligand, and we have set out to investigate the structures of hydroxy- and alkoxy-carboxylates in the solid state. In aqueous solution, there is a good linear free-energy relationship² of the form

$$\log \beta_1 = a \log {}^H K_1 + b \quad (1)$$

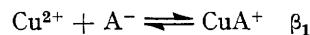
¹ M. Kato, H. B. Jonassen, and J. C. Fanning, *Chem. Rev.*, 1964, **64**, 99.

² F. J. C. Rossotti, J. D. E. Carson, J. J. Clark, K. D. Dillon, D. L. Martin, and V. Moxham, unpublished work.

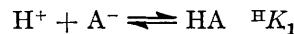
³ H. C. Freeman, 'The Biochemistry of Copper,' ed. P. Aisen, W. Blumberg, and J. Peisach, Academic Press, New York, 1966, p. 77; *Adv. Protein Chem.*, 1967, **22**, 257.

⁴ F. J. C. Rossotti and H. Sunshine, *Chem. Comm.*, 1968, 447.

between the equilibrium constants for the reactions



and



with unequivocally unidentate ligands. In this paper we report the crystal structures of a series of complexes (phenoxyacetate, methoxyacetate, glycollate, lactate, and 2-hydroxy-2-methylpropionate) which are increasingly stabilised² in aqueous solution in terms of equation (1). Diaquobis(methoxyacetato)copper(II) dihydrate and bis(glycollato)copper(II) have been the subject of a preliminary communication.⁹

RESULTS

All five structures were determined by three-dimensional X-ray techniques (see Experimental). In each compound the copper atom is surrounded by six oxygen atoms at the corners of a tetragonally distorted octahedron. However, in the lactate one of the copper-oxygen contacts is so long that it probably should not be regarded as a bond. The carboxylate ligands $O_2C-C-OR$ form two *trans*-five-membered chelate rings with the metal atom, and lie more or less in an equatorial plane of the co-ordination octahedron. The co-ordination sphere may be completed by two aquo-ligands (phenoxyacetate, methoxyacetate, 2-hydroxy-2-methylpropionate), two carboxyl oxygen atoms of neighbouring complexes (glycollate), or one aquo-ligand and one (carboxy-oxygen) lactate. In all crystals the complexes are bound together by a complex series of intermolecular hydrogen-bonds.

Diaquobis(phenoxyacetato)copper(II).—There are two crystallographically different units of the above composition in the crystal. One is located about copper atoms at inversion centres, and the other about copper atoms in general positions (Figure 1). Both units have approximately the same orientation in the crystal, and very similar geometries. Their best planes are parallel to within 1°.

⁵ C. Lea, F. J. C. Rossotti, and D. H. Schärer, unpublished work.

⁶ J. N. van Niekerk and F. R. L. Schoening, *Acta Cryst.*, 1953, **6**, 227; G. A. Barclay and C. H. L. Kennard, *J. Chem. Soc.*, 1961, 5244; F. Hanic, D. Stempelova, and K. Hanicova, *Acta Cryst.*, 1964, **17**, 633; B. H. O'Connor and E. N. Maslen, *ibid.*, 1966, **20**, 824.

⁷ K. D. Dillon and F. J. C. Rossotti, *Chem. Comm.*, 1966, 768.

⁸ F. J. C. Rossotti and A. Willson, unpublished work.

⁹ J. G. Forrest, C. K. Prout, and F. J. C. Rossotti, *Chem. Comm.*, 1966, 658.

Each copper atom is co-ordinated to six oxygen atoms: two phenoxy-oxygen atoms at 2.46—2.50, two carboxy-oxygen atoms at 1.94—1.95, and two aquo-ligands at 1.97—1.99 Å. The copper atoms have therefore the elongated tetragonal distortion of the co-ordination octahedron generally observed. The four-fold axis is through a pair of bonds in the chelate ring, as also occurs in copper croconate.¹⁰ In the centrosymmetric molecule, the angle

The configuration at the phenoxy-oxygen atom is approximately trigonal-planar. The plane of the phenyl groups in the three independent phenoxyacetate residues are inclined to the planes of the chelate rings (Figure 3). The angles of inclination are 8½° for the phenyl group at O(3), 6° for that at O(13), and 9° for that at O(23). In the complex, at a general position in the unit-cell the directions of tilt of the benzene rings are such that the near inversion

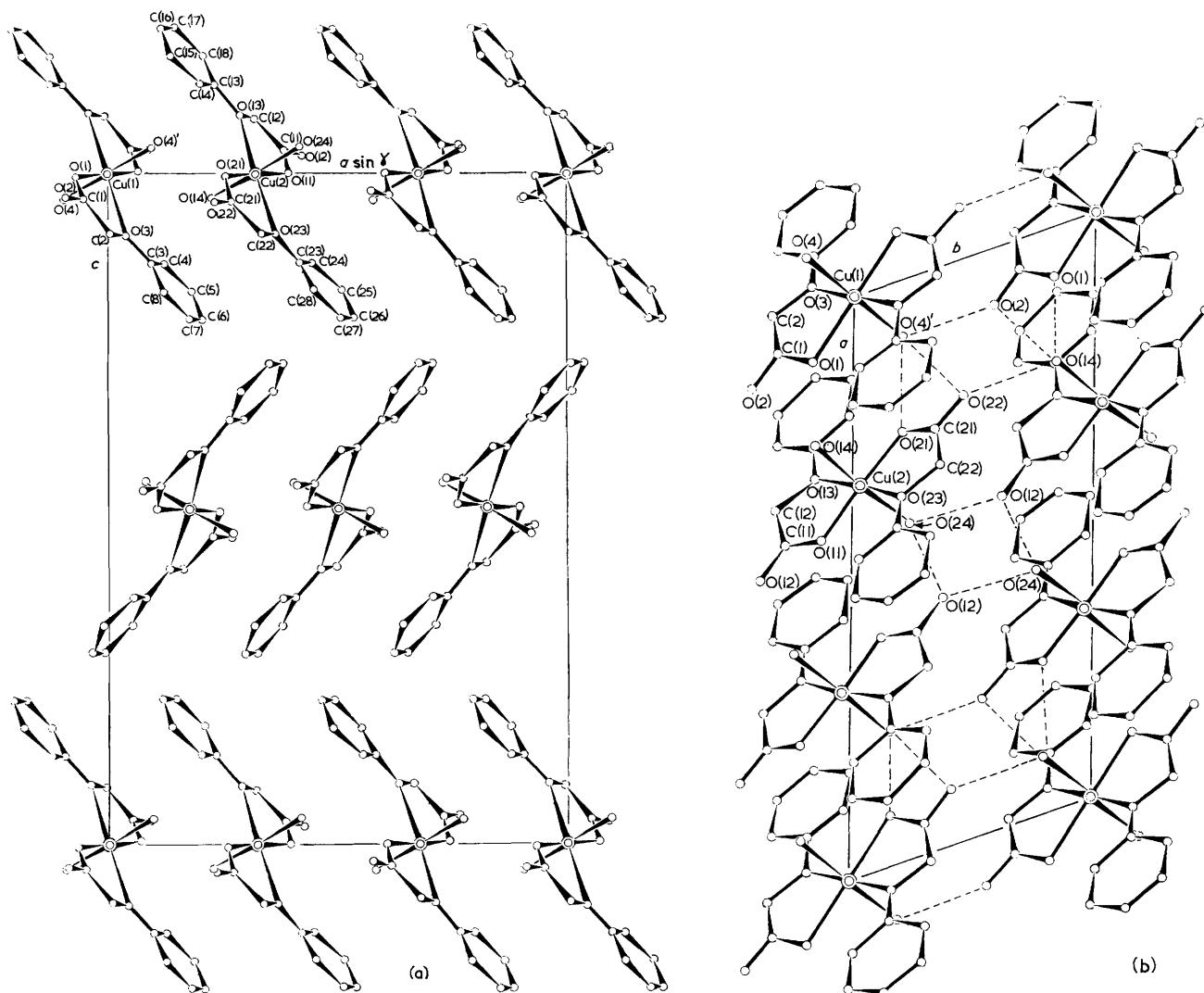


FIGURE 1 Diaquobis(phenoxyacetato)copper(II): (a) Projection of the structure along [010]; the molecules are hydrogen-bonded to give sheets parallel to [010]; (b) one of the sheets projected along [001]

O(1)CuO(3) of 73.6° in the chelate ring is, within experimental error, the same as the corresponding angles O(11)CuO(13) 74.3° and O(21)CuO(23) 75.1° in the other complex. However, in the latter complex the external angles O(13)CuO(21) 110.1° and O(11)CuO(23) 100.5° are different, a feature reflected by the differing behaviour of O(12) and O(22) with respect to intermolecular hydrogen-bonding. In both types of complex the aquo-ligands lie, to a very good approximation, along the perpendiculars to the chelate ring planes at the copper atom. Interatomic distances and interbond angles are summarized in Figure 2.

centre is preserved. As the phenyl group turns out of plane, the carbon atom adjacent to the carbonyl group in the chelate ring is pushed up to 0.15 Å out of the plane in the same direction. Angular distortions at the phenoxy-oxygen and the phenyl ring carbon attached to it are consistent with overcrowding. Each water molecule forms two hydrogen-bonds to oxygen atoms of a carboxyl group. Hydrogen bonds O(4)O(2) of 2.67, O(12)O(24) of 2.70, and O(14)O(22) of 2.71 Å hold the molecules together in chains

¹⁰ M. D. Glick, G. L. Downs, and L. F. Dahl, *Inorg. Chem.*, 1964, **3**, 1712.

parallel to b . The chains are linked to form sheets by the hydrogen-bonds O(4)O(21) of 2.87, O(14)O(1) of 2.84, and O(24)O(12) of 2.72 Å (Figure 1). These three hydrogen-bonds are coplanar with the carboxy-group at which they are formed. In contrast to the other structures described, there are two hydrogen-bonds to carboxy-oxygen atoms which are bound to the metal atoms. For each of these bonds the contacts to the neighbouring non-chelating atoms

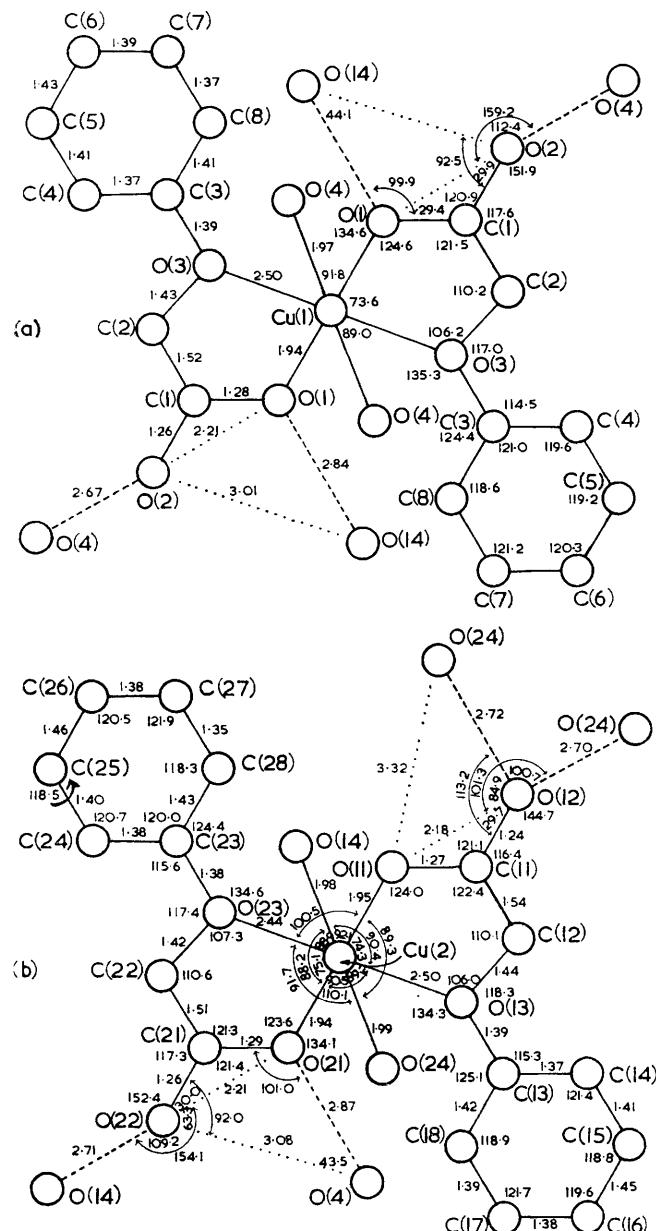


FIGURE 2 Diaquobis(phenoxyacetato)copper(II): interatomic distances and angles (a) in the centrosymmetric molecule about Cu(1); (b) in the molecule containing Cu(2)

[O(2)O(14) 3.01; O(4)O(22) 3.08 Å] suggests that these may be examples of bifurcated hydrogen-bonds. The arrangement of hydrogen-bonds at O(12) is more typical of the behaviour of the non-chelating carboxy-oxygen atom in the compounds described here.

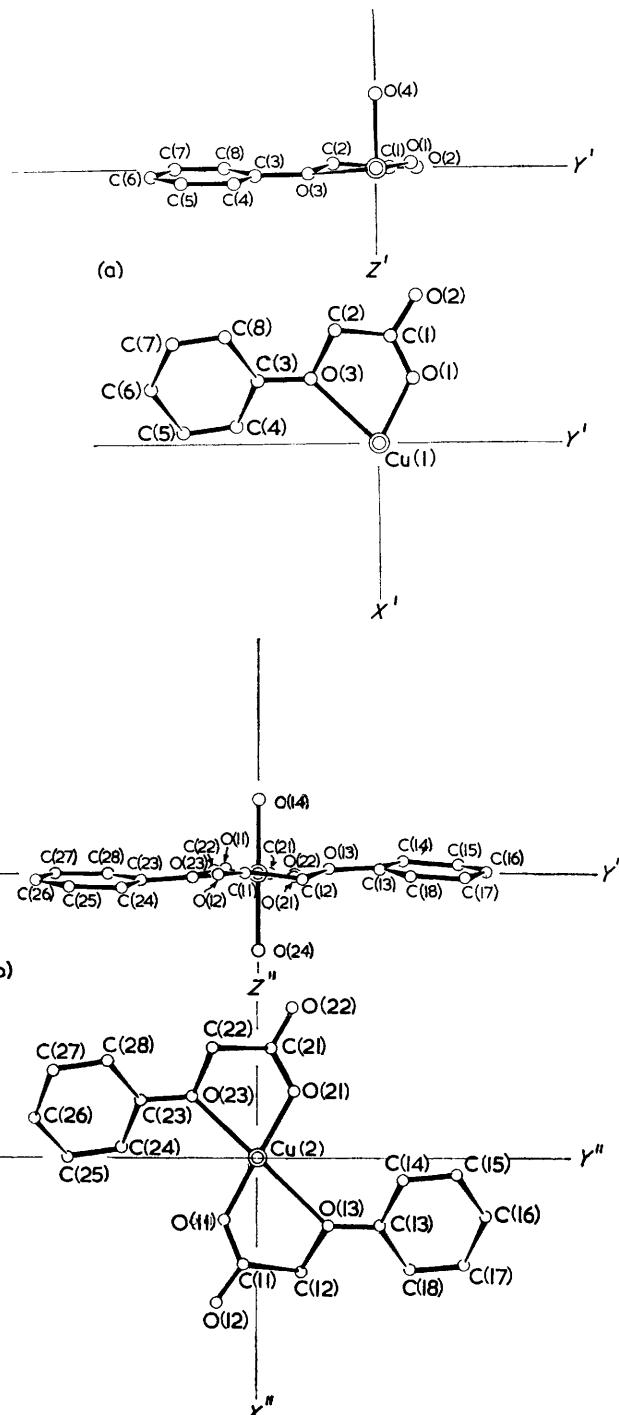


FIGURE 3 Diaquobis(phenoxyacetato)copper(II): molecules (a) about Cu(1), and (b) about Cu(2) projected parallel and perpendicular to the least-squares best planes through the atoms of their chelate rings. The transformation matrices from the orthogonal axes a , b , and c^* to the axes $X'Y'Z'$ and $X''Y''Z''$ are

$$\begin{array}{cccccc} 0.502 & -0.754 & -0.425 & \text{and} & 0.519 & -0.749 & -0.412 \\ -0.653 & -0.008 & -0.757 & & -0.644 & -0.026 & -0.764 \\ 0.567 & 0.657 & -0.497 & & 0.561 & 0.662 & -0.496 \end{array}$$

Diaquobis(methoxyacetato)copper(II).—The crystals are built up from units of this composition located about copper atoms at inversion centres (Figure 4). The co-ordination is

superficially similar to that of the phenoxyacetate but there are interesting differences in the Cu–O contact distances (Figure 5). The co-ordination octahedron has two methoxy-oxygen atoms at 2.13, and two carboxy-oxygen atoms at 1.93 Å in the chelate rings, and is completed by two aquo-ligands at 2.16 Å. The tetragonal distortion therefore takes on the unusual compressed form of four longer and two shorter bonds. The angle O(3)CuO(1) within the ring is 79.9°, and the water molecules lie very near to the perpendicular to the chelate ring plane at the copper atom. The atoms forming the two chelate rings

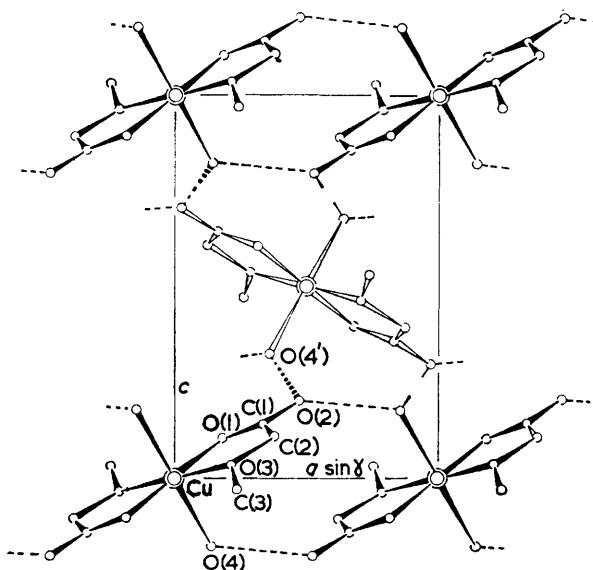


FIGURE 4 Diaquobis(methoxyacetato)copper(II): projection of the structure along [010]. Molecules with copper atoms at y co-ordinates of 0 and 1 are represented by open and full lines respectively

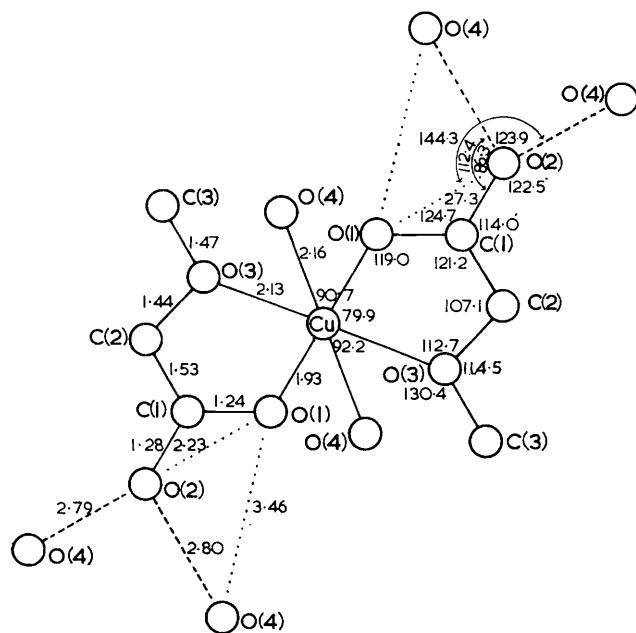


FIGURE 5 Diaquobis(methoxyacetato)copper(II): interatomic distances and angles

atoms at 1.93 Å in the chelate rings, and is completed by two aquo-ligands at 2.16 Å. The tetragonal distortion therefore takes on the unusual compressed form of four longer and two shorter bonds. The angle O(3)CuO(1) within the ring is 79.9°, and the water molecules lie very near to the perpendicular to the chelate ring plane at the copper atom. The atoms forming the two chelate rings

have a maximum deviation of 0.027 and a mean deviation of 0.018 Å from their least-squares best plane. The methyl carbon C(3) is 0.29 Å above the plane and the oxygen O(2) 0.08 Å above it. As in the phenoxyacetate, the oxygen O(3) is nearly coplanar with C(2), C(3), and Cu, and the angle C(3)O(3)C(2) of 114.5° is greater than the expected angle (109°) at an ether oxygen. The bonded distance O(1)C(1) within the chelate ring is unexpectedly shorter than C(1)O(2), but this is without statistical significance. However, a similar short bond is found in bis(methoxyacetato)nickel(II) dihydrate.¹¹ Other bonds have

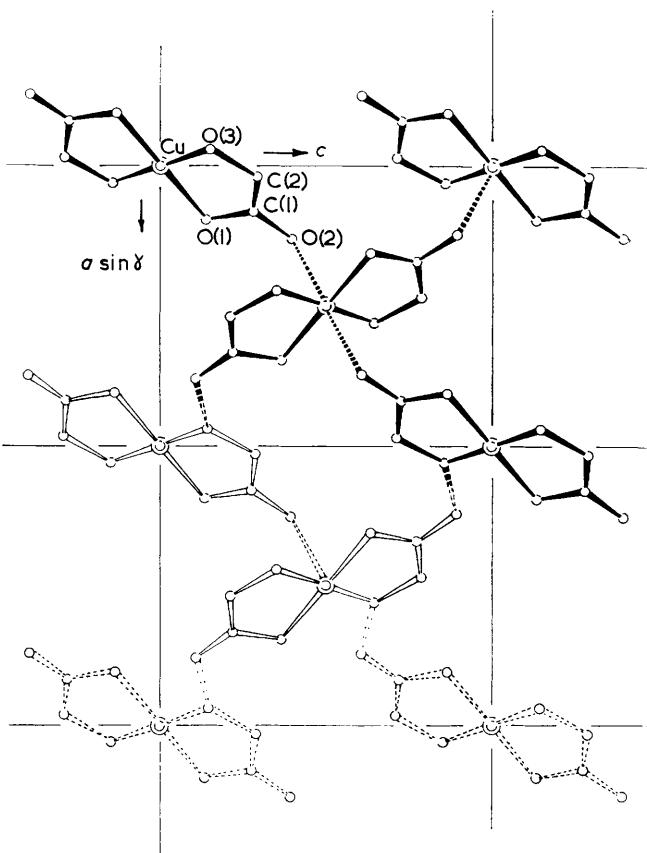


FIGURE 6 Bis(glycollato)copper(II): projection of the structure along [010]. Molecules with copper atoms at y co-ordinates of 0, 1, and 2 are represented by dotted, open, and full lines respectively. Weak Cu–O(2) bonds join molecules at the same y co-ordinates to form sheets perpendicular to [010], and hydrogen-bonds between O(2) and O(3) join the molecules to form sheets perpendicular to [110]

the expected lengths. The carboxyl oxygen O(2) not involved in co-ordination is hydrogen-bonded to two water molecules O(4) of neighbouring complexes (Figure 4). One of these water molecules is in the plane of O(2), C(1), O(1), and the hydrogen-bond 2.80 Å long is approximately parallel to the C(1)C(2) vector. The contact distance O(4)O(1) at 3.46 Å is too long to suggest any tendency to form a bifurcated hydrogen-bond. The second hydrogen-bond 2.79 Å long is also very close to the chelate ring plane, and approximately parallel to the bond C(1)O(1). Each water molecule O(4) has three near neighbours, the two

¹¹ C. K. Prout, F. J. C. Rossotti, and C. Walker, unpublished work.

hydrogen-bond oxygen atoms O(2) of two other complexes and the copper atom in its own complex, forming a shallow pyramid. There are no other intermolecular interactions other than van der Waals contacts.

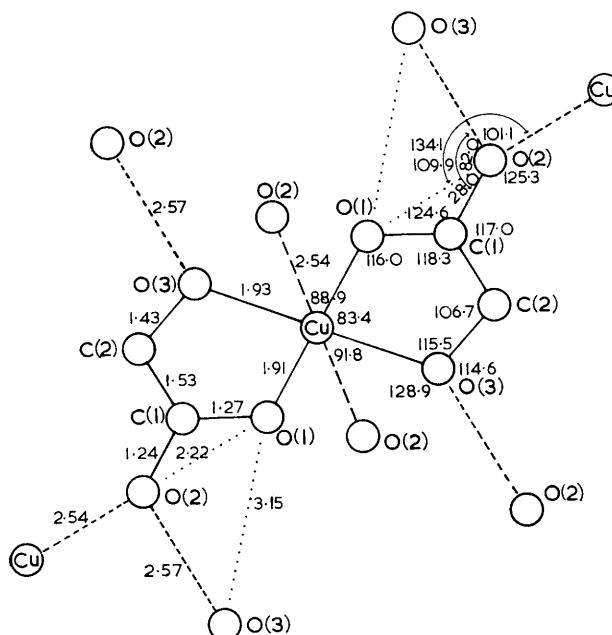


FIGURE 7 Bis(glycollato)copper(II): interatomic distances and angles

Bis(glycollato)copper(II).—Here also, the copper atoms at inversion centres are co-ordinated to two glycollate ligands forming coplanar chelate rings (Figure 6). The copper to hydroxy-oxygen contact $\text{CuO}(3)$ is 1.93 Å and the copper carboxy-oxygen $\text{CuO}(1)$ 1.91 Å, with the angle $\text{O}(1)\text{CuO}(3)$ 83.4° (Figure 7). The copper atom and the atoms of both chelate rings are coplanar within experimental error. The maximum and mean deviations from the least-squares best plane are 0.013 and 0.008 Å respectively. The co-ordination octahedron is completed by the two carboxy-oxygens O(2) of two neighbouring complexes so that the $\text{Cu} \cdots \text{O}(2)$ vectors, 2.54 Å long, are perpendicular to the plane of the copper chelate containing the copper atom. In addition, each oxygen atom O(2) forms a remarkably short intermolecular hydrogen-bond 2.57 Å long to the hydroxy-group O(3) of a second neighbouring complex. This hydrogen-bond is coplanar with the chelate ring associated with O(2), and is approximately parallel to the C(1)C(2) bond of that ring. If the hydrogen-bond is linear, then O(3) has a trigonal-planar configuration. The $\text{CuO}(2)$ contact rises 1.86 Å above the $\text{O}(2)\text{C}(1)\text{O}(1)\text{O}(3)'$ plane but is in a very similar direction to that of the second hydrogen-bond to the carboxy-oxygen in the methoxyacetate structure.

Aquobis-(DL-lactato)copper(II) Hemihydrate.—The crystals are formed from aquobis(lactato)copper(II) groups with the copper atom in general positions in the space-group $A\bar{2}$, and isolated water molecules O(5) at general positions in the unit cell. These water molecules are located in channels parallel to c . There is space for each site to be occupied by water. In the crystal used for the X-ray analysis there were approximately half as many water molecules as sites. Analytical results and density measurements indicate that

the amount of water varies with the age and previous handling of the crystals.

The copper co-ordination sphere has two crystallographically independent chelate rings, one formed from a D-lactate and the other from an L-lactate ligand. The absolute configuration has not been determined. The methyl groups on the 2-carbon atoms appear on opposite sides of the chelate rings (Figure 8). The co-ordination sphere is completed by a water molecule O(4) 2.30 Å from the copper at the apex of an approximately tetragonal prism (Figure 9). The copper atom is 0.15 Å above the best basal plane of the prism taken through the ring oxygen atoms (Figure 10). The oxygen atoms O(1) and O(11) are 0.05 Å above the plane, and the hydroxy-oxygen atoms O(3) and O(13) are 0.05 Å below the plane. The chelate rings as a whole are bent away from the water molecule O(4) in a bow shape. The individual rings are far less planar than those found in the methoxyacetate and glycolate complexes. In the ring containing O(1) the maximum

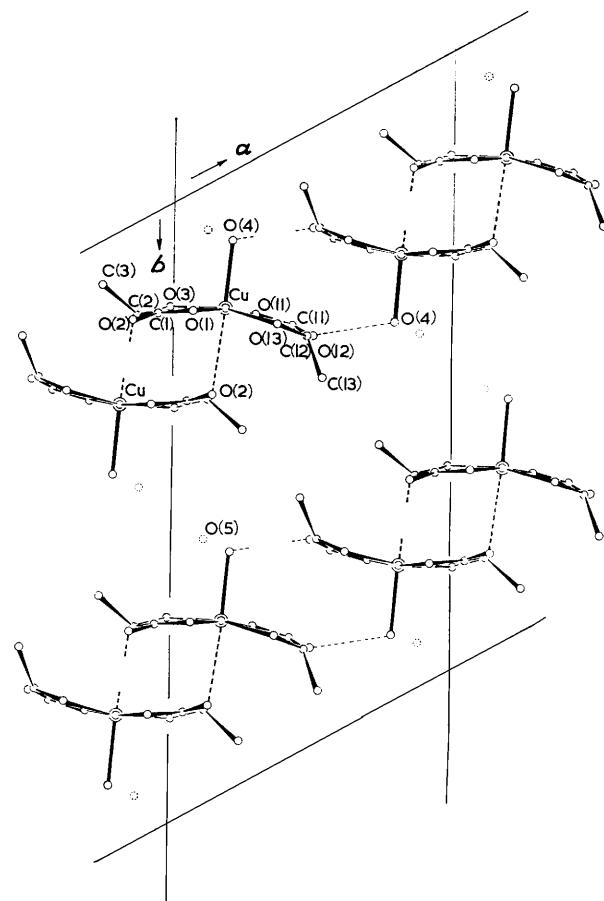


FIGURE 8 Aquobis-(DL-lactato)copper(II) hemihydrate: projection of the structure along [001]. The short contacts $\text{Cu}-\text{O}(2)$ and $\text{O}(4)-\text{O}(12)$ are shown by discontinuous dotted lines when they are to molecules at heights other than those drawn. The water molecule O(5), shown as a broken circle, has an occupation number of about 0.5

deviation from the best plane is 0.082 and the mean deviation 0.055 Å. For the ring containing O(11) the corresponding values are 0.070 and 0.046 Å. The aquobis(lactato)copper(II) complexes are joined by hydrogen-bonds

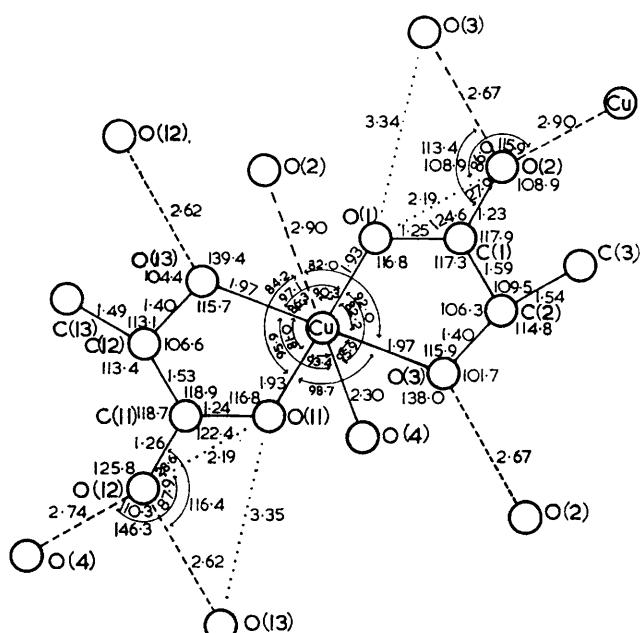


FIGURE 9 The aquobis-(DL-lactato)copper(II) molecule: interatomic distances and angles

to form ribbons parallel to c , and these ligands are associated in pairs by a weak interaction of the copper atoms and the non-chelating carboxy-oxygen atom O(2) at a separation of 2.9 Å (Figure 11). This oxygen atom O(2) forms one hydrogen-bond to the hydroxy-oxygen O(3) of a neighbouring complex, as well as the short contact to the copper atom in a complex related to that containing O(2) by rotation about a two-fold screw axis. The oxygen atom O(12) is

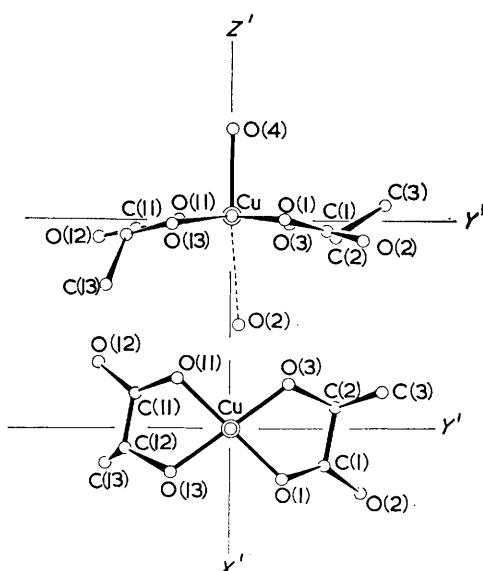


FIGURE 10 The aquobis-(DL-lactato)copper(II) molecules projected parallel and perpendicular to the least-squares best plane through O(1), O(3), O(11), and O(13). The transformation matrix from the orthogonal axes a , b^* , and c to the axes X' , Y' , and Z' is

$$\begin{array}{ccc} 0.179 & 0.175 & 0.968 \\ 0.791 & 0.560 & -0.247 \\ -0.586 & 0.810 & -0.038 \end{array}$$

involved in two hydrogen-bonds. One is to the hydroxy-group O(13) in another complex (Figure 11). The other is to the water molecules O(4) of a neighbouring complex and joins the double ribbons of complexes to form sheets (Figure 8). The hydrogen-bonds about O(12) are very similar to those about O(2) in the methoxyacetate. If it is assumed that the hydrogen-bonds are linear, the hydroxy-oxygen atoms O(3) and O(13) have a planar configuration.

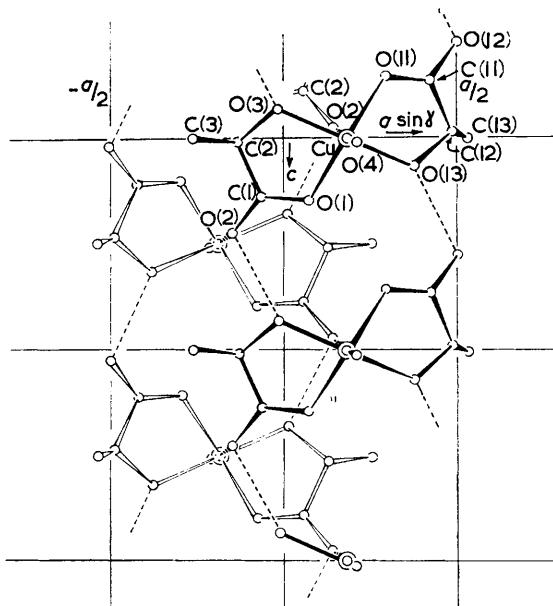


FIGURE 11 Aquobis-(DL-lactato)copper(II) hemihydrate: the double ribbon formation projected along [010]

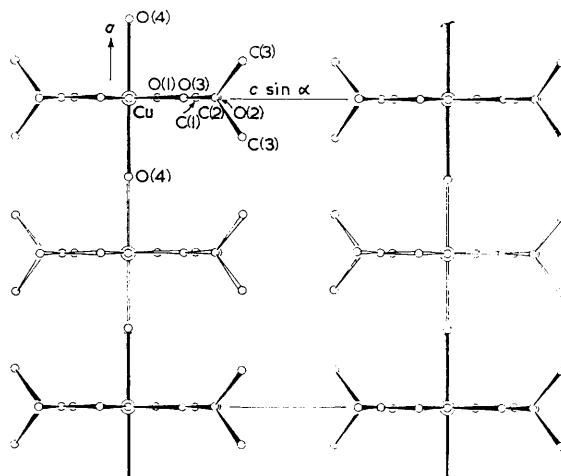


FIGURE 12 Diaquobis-(2-hydroxy-2-methylpropionato)-copper(II): projection of the structure along [010]

Diaquobis-(2-hydroxy-2-methylpropionato)copper(II).—This structure is the least well determined. In the isolated complex the copper atom appears to lie at a point of symmetry $2/m$ in a unit cell of space-group $C2/m$. The atoms of the chelate rings lie in the mirror plane (Figure 12). The co-ordination of the copper is distorted octahedral with the copper-water contact $\text{CuO}(4)$ 2.56, copper-hydroxy-group $\text{CuO}(3)$ 2.01, and the copper-carboxy-oxygen

$\text{CuO}(1)$ 1.89 Å (Figure 13). The chelate rings and the copper atom are necessarily coplanar, but difference maps indicate a remarkably high anisotropic thermal motion along the $\text{O}(4)\text{CuO}(4)'$ vector. This appears better explained as a statistical distribution of copper atoms about the mirror plane (see Experimental) than as true anisotropic thermal

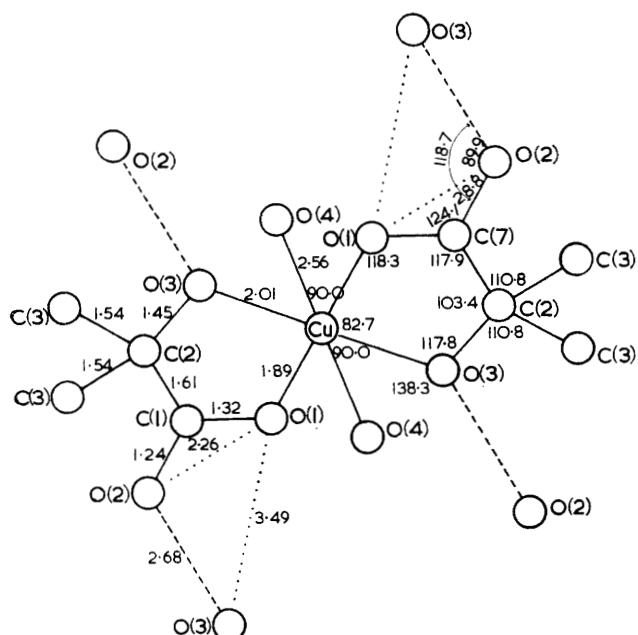


FIGURE 13 Diaquobis-(2-hydroxy-2-methylpropionato)-copper(II): interatomic distances and angles

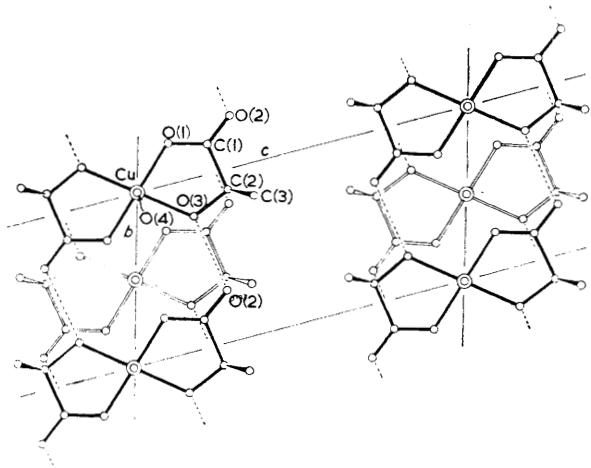


FIGURE 14 Diaquobis-(2-hydroxy-2-methylpropionato)copper(II): projection of the structure along [100] showing the hydrogen-bonded ribbons

motion, and may represent a tendency to form a copper lactate-like structure. The structures of other modifications derived from the original triclinic twinned crystal may yield further information on this point. The complexes are joined together to form ribbons (Figure 14) by hydrogen-bonding of the hydroxy-group O(3) to O(2) of another molecule in the mirror plane, and the sheets are held together by hydrogen-bonding through the water molecules.

DISCUSSION

The only other example of a hydroxycarboxylato-copper(II) structure in the literature appears to be that of the diaquobis(salicylato)-dihydrate.¹² The salicylate ligands, which may be regarded as 3-hydroxycarboxylates, are unidentate and bound to the copper ion through the carboxy-groups with Cu–O bonds of 1.91 Å. Chelation by the hydroxy-group is prevented by strong intramolecular hydrogen-bonding to the carboxylate oxygen unattached to copper. The copper ions are also co-ordinated to two aquo-ligands at a Cu–O distance of 1.95 Å, and form an asymmetric $[\text{Cu}(\text{OH}_2)_2]$, chain with longer Cu–O contacts of 2.8 Å. The interatomic distances in this compound have been calculated from a revised set of atomic parameters obtained from the data in the literature¹² by four cycles of least-squares refinement.

In Tables 1 and 2 the formally equivalent bond lengths and bond angles for the compounds reported here and for diaquobis(salicylato)copper(II) dihydrate are collected together. For all compounds each independent measurement of a particular bond is given. The mean values are the simple arithmetic means, and the standard deviations of the mean values were computed from the formula

$$\sigma_m = [\sum_i (d_i - \bar{d})^2 / n(n-1)]^{1/2}$$

where d_i is the i th of n measurements for which \bar{d} is the mean value. In parentheses are the standard deviations of individual bond lengths and angles (for bond lengths $\sigma_i \times 10^2$) calculated from the standard deviations of atomic parameters in Table 3 and neglecting the covariances. Tables 1 and 2 are in the same form as the analogous Tables for copper-2-aminocarboxylates and copper peptides given by Freeman.³

The copper to carboxy-oxygen bond lengths represent a spread about a mean of 1.92 Å, which is somewhat shorter than the corresponding mean length of 1.97 Å found in the aminocarboxylate and peptide complexes³ and in the copper carboxylate dimers.⁶ The mean Cu–OR bond length is longer (2.16 Å), but the spread is much bigger, and the lengths increase in the order R = H < Me < Ph. The mean of the Cu–OH₂ bond lengths is also 2.16 Å, but again the spread is significantly larger than for the Cu–OOC bond lengths. As the bond length Cu–OR increases, so the Cu–OH₂ bond length decreases. It is noteworthy that the short Cu–OH₂ contact of 1.97 Å in the phenoxyacetate is equal to the mean short Cu–OH₂ contact in the hydroxycarboxylates. Similarly, the longer Cu–OH₂ contacts approximate to the longer Cu–OR contacts.

In the three diaquobiscarboxylate complexes, the CuO₆ co-ordination polyhedra all have symmetry D_{2h} , but may be considered to approximate to D_{4h} . In the 2-hydroxy-2-methylpropionate, the four-fold axis lies along the copper–water vector, and there are four short and two long bonds. In the methoxyacetate, the four-fold axis is along the copper to carboxylate oxygen bond,

¹² F. Hanic and J. Michalov, *Acta Cryst.*, 1959, **13**, 299.

and there are four long and two short contacts. In the phenoxyacetate, the four-fold axis is along the copper to phenoxy-oxygen bond, and there are four short and two long contacts. These tetragonal distortions may be attributed to the Jahn-Teller effect, provided that the oxygen atoms around the central copper atom neither have different radii nor have markedly different intrinsic donor powers due to their different chemical environments. We find¹¹ that diaquobis(methoxyacetato)-nickel(II), which is isomorphous with the copper complex, has three nearly equal independent pairs of Ni-O bonds (of which the Ni-OMe bond is the shortest at 1.99 Å).

form, a later¹⁸ MO treatment predicts a difference of about 0.2 Å between the shorter and longer Cu-F bonds, but is unable to define the direction of the distortion, which appears to be very sensitive to the amount of covalency in the bonds.¹⁹ In the present CuO₆ polyhedra, the three pairs of oxygen donors are nearly, but not quite, equivalent. The free energies² and enthalpies²⁰ of the copper complexes in solution indicate that the strengths of the bonds from Cu to CO₃²⁻, OH, OMe, and OPh differ remarkably little from the strength of the copper-water bonds, but decrease in that order, which is also the order of increasing bond lengths. It

TABLE 1
Collected bond lengths in the complexes (in Å) with standard deviations ($\times 10^2$) in parentheses

Compound	Organic ligand										'Axial' bonds	
	CO ₃ ²⁻			C-O-R							Cu-OH ₂	Cu-O
(a) Diaquobis(phenoxyacetato)copper(II)	1.26(2) 1.26(2)	1.28(2) 1.29(2)	1.94(1) 1.94(1)	2.21(2) 2.21(2)	1.43(2) 1.42(2)	— —	1.39(2) 1.38(2)	— —	2.50(1) 2.44(1)	2.84(3) 2.79(3)	1.52(2) 1.51(2)	1.97(1) 1.98(1)
(b) Diaquobis(methoxyacetato)copper(II)	1.24(2) 1.28(2)	1.27(2) 1.24(2)	1.94(1) 1.93(1)	2.23(2) 2.18(2)	1.44(2) 1.44(2)	— 1.47(2)	1.39(2) —	— —	2.50(1) 2.57(2)	2.76(3) 2.57(2)	1.54(2) 1.53(2)	1.99(1) 2.18(1)
(c) Bis(glycollato)copper(II)	1.24(2) 1.24(2)	1.27(2) 1.24(2)	1.91(1) 1.93(1)	2.22(2) 2.19(2)	1.43(2) 1.40(4)	— —	— —	— —	2.57(2) 2.67(3)	1.93(1) 1.97(2)	2.54(3) 2.56(3)	2.54(1) 2.30(2)
(d) Aquobis-(DL-lactato)copper(II) hemihydrate	1.23(3) 1.26(3)	1.25(3) 1.24(3)	1.93(2) 1.93(2)	2.19(3) 2.19(3)	1.40(4) 1.40(4)	— —	— —	— —	2.62(3) 2.68(3)	1.97(2) 2.01(2)	2.56(3) 2.58(3)	1.59(4) 1.61(4)
(e) Diaquobis(2-hydroxy-2-methylpropionato)-copper(II)	1.26(3) 1.24(4)	1.29(3) 1.32(4)	1.93(2) 1.89(2)	2.19(3) 2.26(4)	1.40(4) 1.45(4)	— —	— —	— —	2.68(3) 2.01(2)	1.95 (Cu-OH ₂)	2.56(3) 2.01(2)	— —
(f) Diaquobis(salicylato)copper(II) dihydrate	1.27	1.28	1.91	2.22	[1.32]	—	—	—	—	[1.43]	[2.80]	—
Mean	1.254	1.271	1.924	2.212	1.426	1.47	1.39	2.64	2.160	2.656	1.545	2.160
Standard deviation of the mean	0.006	0.009	0.006	0.008	0.007	—	0.003	0.026	0.084	0.042	0.013	0.096
Range	1.23	1.24	1.89	2.18	1.40	1.47	1.38	2.57	1.93	2.54	1.51	1.97
Number in sample	9	9	9	9	8	1	3	4	9	8	9	7

TABLE 2
Collected bond angles in the complexes (in degrees) with standard deviations in parentheses

Compound	CO ₃ ²⁻				C-O-R				'Axial' angles				
	O-Cu-O	Cu-O-C	O-C=O	O-C-C	O-C-C	C-C-O	C-OH···O	C-O-R	Cu-O-C	Cu-O-R	Cu-OH···O	Cu-O-Cu	
(a) Diaquobis(phenoxyacetato)copper(II)	73.6(0.8) 75.1(0.3)	124.6(0.9) 123.6(0.9)	120.9(1.2) 121.4(1.2)	121.5(1.2) 121.3(1.2)	117.6(1.2) 116.6(1.2)	110.2(1.1) —	— —	117.0(1.0) 117.4(1.0)	106.2(0.7) 107.3(0.7)	135.3(0.8) 134.6(0.7)	— —	—	
(b) Diaquobis(methoxyacetato)copper(II)	74.3(0.3) 79.9(0.4)	124.0(0.9) 119.0(0.9)	122.4(1.2) 124.7(1.2)	121.1(1.2) 121.2(1.2)	116.4(1.2) 114.0(1.3)	110.1(1.1) 107.1(1.1)	— —	118.3(1.0) 114.5(1.1)	106.0(0.7) 112.7(0.8)	134.3(0.7) 130.4(0.9)	— —	—	
(c) Bis(glycollato)copper(II)	83.4(0.4) 83.4(0.4)	116.0(0.9) 116.0(0.9)	124.6(1.3) 124.6(1.3)	118.3(1.2) 117.0(1.2)	117.0(1.2) 106.7(1.1)	106.7(1.1) 114.6(1.2)	— —	— 115.5(0.8)	— —	— 128.9(0.9)	— —	—	
(d) Aquobis-(DL-lactato)copper(II) hemihydrate	82.2(0.4) 81.0(0.7)	116.8(1.5) 118.1(1.5)	124.6(2.2) 122.4(2.2)	121.7(2.2) 118.7(2.2)	117.3(2.1) 118.9(2.2)	106.3(2.2) 106.6(2.2)	101.7(2.0) 104.4(2.0)	— —	— 115.7(1.6)	— —	— 139.4(1.9)	— —	—
(e) Diaquobis(2-hydroxy-2-methylpropionato)-copper(II)	82.7(0.8) 82.7(0.8)	118.3(1.7) 118.3(1.7)	124.1(2.8) 124.1(2.8)	117.9(2.5) 118.1(2.2)	103.4(2.2) 103.4(2.2)	104.3(2.8) 104.3(2.8)	— —	— 117.8(1.8)	— —	— —	— —	— 138.3(1.8)	—
(f) Diaquobis(salicylato)copper(II) dihydrate	[88.5] 79.0 73.6 82.7	121.8 120.1 116.0— 124.6	121.4 122.9 120.9— 124.6	121.2 119.9 117.3— 121.5	117.4 117.1 114.0— 118.7	— 107.6 103.4— 110.6	— 111.5 101.7— 118.3	— 111.5 101.7— 117.8	— 112.1 106.0— 8	— 112.1 106.0— 8	— 134.9 128.9— 8	— 1.3 — 139.4	—
Mean	88.5	121.8	121.4	121.2	117.4	—	—	—	—	—	—	—	—
Standard deviation of the mean	1.4°	1.2	0.5	0.5	0.5	0.9	2.4	—	1.6	—	—	1.3	—
Range	73.6— 82.7	116.0— 9	120.9— 9	117.3— 9	114.0— 9	103.4— 8	110.6— 8	118.3— 8	117.8— 8	106.0— 8	128.9— 8	139.4— 8	—
Number in sample	8	9	9	9	9	8	8	8	8	8	8	8	—

Accordingly, the tetragonal distortions do appear to be primarily due to the metal ion rather than to the ligands.

It is widely believed that the CuF₆ polyhedron shows compressed tetragonal distortion in KCuF₃ and K₂CuF₄. However, the earlier report¹³ of four long and two short Cu-F bonds in KCuF₃ appears to have been erroneous.¹⁴ Similarly, the original structure¹⁵ of K₂CuF₄ has been questioned.¹⁶ Our diaquobis(methoxyacetato)copper(II) complex appears to be a rare example of the compressed tetragonal distortion. Although Liehr and Ballhausen's electrostatic model¹⁷ for CuF₆ suggested that the compressed octahedron is more stable than the elongated

appears that the four-fold axis lies along the two bonds that differ most in strength from the remaining four, i.e., along the weak Cu-OH₂ bonds in the 2-hydroxy-2-methylpropionate, and along the weak Cu-OPh bonds in the phenoxyacetate, but along the strong Cu-OOC bonds in the methoxyacetate.

In accommodating the varying Cu-OR bond lengths, distortions in the chelate ring are necessary. Equivalent bonds within the various ligands have lengths randomly spread about their means which are acceptable values for each particular type of bond. The separation between the two ring oxygen atoms increases as R is changed

¹³ A. J. Edwards and R. D. Peacock, *J. Chem. Soc.*, 1959, 4126.

¹⁴ A. Okazaki and Y. Suemune, *J. Phys. Soc. Japan*, 1961, **16**, 176.

¹⁵ K. Knox, *J. Chem. Phys.*, 1959, **30**, 991.

¹⁶ K. Oelkrug, *Z. phys. Chem. (Leipzig)*, 1967, **56**, 325; D. Babel, *Structure and Bonding*, 1967, **3**, 1.

¹⁷ A. D. Liehr and C. J. Ballhausen, *Ann. Phys.*, 1958, **3**, 304.

¹⁸ C. J. Ballhausen and H. Johansen, *Mol. Phys.*, 1966, **10**, 183.

¹⁹ M. H. L. Pryce, K. P. Sinha, and Y. Tanabe, *Mol. Phys.*, 1965, **9**, 33.

²⁰ J. D. Pocock and F. J. C. Rossotti, unpublished work.

from H to Me to Ph with a concomitant opening of the angles at the carbon atoms. The angle C-C-OR is less than the tetrahedral angle for R = H, and greater for R = Ph. The angles at the carboxy-group are symmetrical with respect to the C-C bond for R = H, but unsymmetrical for R = Me and Ph. If it is assumed

TABLE 3

Atomic co-ordinates ($\times 10^4$) and their estimated standard deviations ($\times 10^4$)

	x/a	$\sigma(x/a)$	y/b	$\sigma(y/b)$	z/c	$\sigma(z/c)$	U_{iso}	$\sigma(U_{iso})$
(a) Diaquobis(phenoxyacetato)copper(II)								
Cu(1)	0	0	0	0	0	0	351	7
O(1)	681	5	-1738	10	-51	3	363	18
O(2)	-929	6	4345	11	363	3	416	21
O(3)	414	5	1788	10	904	3	309	17
O(4)	-968	5	-2018	10	384	3	321	17
C(1)	-530	8	3147	16	397	4	324	25
C(2)	131	8	3440	15	867	4	289	24
C(3)	961	8	1754	14	1344	4	300	25
C(4)	1246	9	196	16	1350	5	378	28
C(5)	1858	10	56	18	1762	5	465	33
C(6)	2110	11	1478	20	2175	5	541	36
C(7)	1820	10	3040	20	2171	5	520	35
C(8)	1237	9	3214	16	1757	4	363	27
Cu(2)	3292	1	344	2	4	1	371	5
O(11)	4004	6	-1330	11	5	4	447	20
O(12)	4312	6	-3870	11	-272	3	368	19
O(13)	2922	5	1588	10	-889	3	297	17
O(14)	2316	6	-1673	10	389	3	329	18
C(11)	3892	8	-2747	15	-325	4	300	25
C(12)	3240	8	-3182	14	-817	4	283	24
C(13)	2370	8	-1653	14	-1327	4	277	23
C(14)	2049	9	-155	17	-1350	5	392	28
C(15)	1403	10	-114	18	1763	5	483	33
C(16)	1165	11	-1569	20	-2162	6	542	36
C(17)	1507	11	-3070	21	-2146	6	551	36
C(18)	2112	9	-3150	17	-1729	5	409	29
O(21)	2626	5	2098	10	55	3	364	18
O(22)	2369	6	4648	12	415	3	471	22
O(23)	3725	5	2034	10	890	3	327	18
O(24)	4251	6	2392	10	-393	3	353	17
C(21)	2774	9	3453	16	426	4	352	27
C(22)	3414	8	3634	15	897	4	305	24
C(23)	4266	8	1920	14	1320	4	283	23
C(24)	4544	9	341	17	1302	5	417	29
C(25)	5173	10	150	19	1709	5	485	34
C(26)	5431	10	1510	20	2134	5	529	35
C(27)	5154	10	3060	20	2149	6	551	37
C(28)	4556	9	3310	18	1744	5	437	31
H(1)	691		4738		782		400	
H(2)	-163		3622		1252		400	
H(3)	-954		-3382		376		400	
H(4)	1008		-962		1037		400	
H(5)	2114		-1127		1752		400	
H(6)	2553		1365		2504		400	
H(7)	2038		4148		2493		400	
H(8)	1013		4448		1757		400	
H(9)	-1520		-2505		159		400	
H(11)	3569		-3337		-1192		400	
H(12)	2702		-4500		-735		400	
H(13)	2334		-3034		397		400	
H(14)	2277		1028		-1048		400	
H(15)	1126		1031		-1755		400	
H(16)	702		1539		2487		400	
H(17)	1303		-4211		-2461		400	
H(18)	2363		-4338		-1726		400	
H(19)	1759		-2175		169		400	
H(21)	3955		4970		850		400	
H(22)	3087		3636		1285		400	
H(23)	4271		3765		-349		400	
H(24)	4300		-762		981		400	
H(25)	5429		-1041		1678		400	
H(26)	5469		1341		2459		400	
H(27)	5401		4140		2477		400	
H(28)	4333		4542		1769		400	
H(29)	4776		2929		-149		400	

TABLE 3 (Continued)

x/a	$\sigma(x/a)$	y/b	$\sigma(y/b)$	z/c	$\sigma(z/c)$	U_{iso}	$\sigma(U_{iso})$
(b) Diaquobis(methoxyacetato)copper(II)							
Cu	0	0	0	0	0	0	376
O(1)	1842	14	-1154	11	1051	7	345
O(2)	4731	14	-837	11	2039	8	365
O(3)	2096	14	2341	12	353	7	346
O(4)	1423	16	-708	13	-1794	8	435
C(1)	3356	19	-205	15	1411	10	268
C(2)	3739	21	1866	16	1082	11	337
C(3)	2398	25	4157	20	-312	13	472
(c) Bis(glycollato)copper(II)							
Cu	0	0	0	0	0	0	273
O(1)	1575	12	2063	18	1604	12	207
O(2)	2572	12	1676	19	3954	12	217
O(3)	-614	12	-3055	18	1469	11	203
C(1)	1612	17	681	25	2809	17	186
C(2)	335	18	-2403	26	2907	17	209
(d) Aquobis(lactato)copper(II) hemihydrate							
Cu	1816	2	2158	1	0	0	436
O(1)	736	13	1906	7	2923	29	405
O(2)	-1418	15	1548	7	4455	32	500
O(3)	-154	12	1647	6	-1362	32	371
O(4)	2090	13	1118	6	198	40	465
O(5)	1184	29	5713	15	270	79	606
C(1)	-606	17	1652	9	2769	41	320
C(2)	-1292	18	1493	9	198	56	416
C(3)	-2576	20	688	10	75	72	546
C(11)	4186	18	3027	9	-2722	42	339
C(12)	4808	17	3313	8	-265	51	356
C(13)	5290	19	4118	10	68	66	524
O(11)	2918	12	2505	4	-2896	28	367
O(12)	4967	14	3336	7	-4502	32	473
O(13)	3745	14	2855	7	1355	34	431

[The occupation number of O(5) is 0.503 with an e.s.d. of 0.053.]

(e) Diaquobis-(2-hydroxy-2-methylpropionato)copper(II)

(i) With anisotropic temperature factors for the copper atom giving $R = 0.169$.

x/a	$\sigma(x/a)$	y/b	$\sigma(y/b)$	z/c	$\sigma(z/c)$	U_{iso}	$\sigma(U_{iso})$
Cu	0	0	0	0	0	0	—
O(1)	0	0	-2364	35	882	16	378
O(2)	0	0	-3069	40	2775	18	509
O(3)	0	0	2295	34	1712	15	351
O(4)	2534	27	0	0	0	0	736
C(1)	0	0	-1656	55	2128	24	398
C(2)	0	0	1297	52	2750	22	334
C(3)	1236	25	1954	41	3545	19	554

The anisotropic temperature factors ($\times 10^4$) for the copper atom with their e.s.d. in parentheses are U_{11} , 879(44); U_{22} , 260(37); U_{33} , 71(18); U_{12} , 0(0); U_{13} , 0(0); U_{23} , 47(35).

(ii) With disordered copper atoms giving $R = 0.129$.

Cu(1)	0	0	0	0	0	0	156	—
Cu(2)	550	—	0	0	0	0	156	—
O(1)	0	0	-2353	29	881	13	439	38
O(2)	0	0	-3007	30	2774	13	467	41
O(3)	0	0	2261	27	1714	12	366	36
O(4)	2520	22	0	0	0	0	771	56
C(1)	0	0	-1631	41	2098	19	364	51
C(2)	0	0	1209	38	2749	16	281	44
C(3)	1229	20	1946	33	3545	15	573	47

The occupation numbers of Cu(1) and Cu(2) are 0.702 and 0.149 respectively. Parameters for which there are no e.s.d.'s were not refined in the final cycles.

that the hydrogen-bonds formed by the hydroxy-groups in the hydroxycarboxylates are linear, the oxygen atom of all Cu-O(R)-C groups appears to be trigonal-planar to a high degree of approximation, although the angles about the oxygen may deviate widely from 120° . By contrast, oxygen atoms in bridging hydroxy-groups²¹ or

²¹ C. K. Prout, *J. Chem. Soc.*, 1962, 4429.

TABLE 4

Observed structure amplitudes and calculated structure factors for diaquobis(phenoxyacetato)copper(II)

<i>h</i>	<i>k</i>	<i>l</i>	<i>F</i> _o	<i>F</i> _c	<i>h</i>	<i>k</i>	<i>l</i>	<i>F</i> _o	<i>F</i> _c	<i>h</i>	<i>k</i>	<i>l</i>	<i>F</i> _o	<i>F</i> _c	<i>h</i>	<i>k</i>	<i>l</i>	<i>F</i> _o	<i>F</i> _c		
0	0	8	566	635	0	4	17	98	113	1	-1	19	48	-48	1	4	17	55	-57		
0	0	10	667	701	0	4	18	163	179	1	-1	20	67	73	1	4	19	51	-60		
0	0	12	414	392	0	4	20	97	114	1	-1	22	82	86	1	4	21	48	-54		
0	0	14	391	322	0	4	21	68	64	1	-1	23	52	65	1	4	23	49	-54		
0	0	16	78	56	0	4	22	132	124	1	-1	24	24	-22	1	4	25	42	-44		
0	0	18	235	183	0	4	23	73	75	1	-1	25	36	55	1	4	1	221	231		
0	0	20	182	149	0	4	24	88	90	1	-1	26	18	-14	1	4	2	35	-24		
0	0	22	599	682	0	4	25	38	35	1	-1	27	32	38	1	4	3	269	286		
0	0	24	505	568	0	4	26	91	87	1	2	1	69	54	1	4	5	341	363		
0	0	26	61	104	0	5	1	469	490	1	2	2	37	-25	1	4	6	85	77		
0	0	28	226	245	0	5	2	167	172	1	2	3	61	-69	1	4	7	303	303		
0	0	30	134	165	0	5	3	265	263	1	2	4	57	-49	1	4	8	64	57		
0	1	4	254	199	0	5	5	155	132	1	2	5	151	-145	1	4	9	250	249		
0	1	6	626	710	0	5	6	23	16	1	2	6	100	-82	1	4	11	204	203		
0	1	6	171	-136	0	5	7	377	382	1	2	7	172	-176	1	4	13	175	169		
0	1	7	820	1112	0	5	8	179	-198	1	2	8	77	-67	1	4	14	65	-55		
0	1	8	613	552	0	5	9	271	229	1	2	9	154	-172	1	4	15	181	181		
0	1	9	739	928	0	5	10	59	-54	1	2	10	35	-36	1	4	16	87	-81		
0	1	10	78	-41	0	5	11	182	175	1	2	11	126	-139	1	4	17	167	166		
0	1	11	652	671	0	5	12	170	146	1	2	12	41	-40	1	4	18	87	-71		
0	1	12	42	-51	0	5	13	109	96	1	2	13	138	-153	1	4	19	130	-124		
0	1	13	650	693	0	5	14	159	172	1	2	14	45	-44	1	4	20	58	-52		
0	1	14	341	-317	0	5	15	34	56	1	2	15	81	-96	1	4	21	102	96		
0	1	15	636	680	0	5	16	100	85	1	2	16	28	-26	1	4	22	41	-36		
0	1	16	259	-250	0	5	17	39	-34	1	2	17	81	-92	1	4	23	63	58		
0	1	17	441	483	0	5	18	22	21	1	2	18	62	-58	1	4	24	36	27		
0	1	18	23	-27	0	5	19	154	172	1	-2	1	108	106	1	4	25	43	36		
0	1	19	268	292	0	5	20	44	-46	1	-2	2	110	-90	1	4	27	56	60		
0	1	20	76	70	0	5	21	132	140	1	-2	3	161	158	1	5	0	175	-192		
0	1	21	299	311	0	5	22	86	76	1	-2	4	160	-142	1	5	1	163	-173		
0	1	22	197	161	0	5	23	482	556	1	-2	5	175	187	1	5	3	54	-48		
0	1	23	304	316	0	5	24	142	104	1	-2	6	193	-182	1	5	4	108	-106		
0	1	24	231	-225	0	5	25	3	65	1	-2	7	193	165	1	5	5	41	-35		
0	1	25	122	118	0	5	26	149	145	1	-2	8	119	-108	1	5	6	102	-87		
0	1	26	48	-71	0	5	27	115	136	1	-2	9	99	110	1	5	7	24	-21		
0	1	27	102	90	0	5	28	141	129	1	-2	11	54	57	1	5	8	108	-99		
0	1	28	26	33	0	5	29	8	112	99	1	-2	12	42	39	1	5	10	86	-84	
0	2	29	153	153	0	5	30	10	160	171	1	-2	13	69	75	1	5	12	74	-68	
0	2	30	228	-147	0	5	31	12	174	156	1	-2	14	109	117	1	5	13	35	-26	
0	2	31	222	239	0	5	32	13	33	1	-2	15	106	99	1	5	14	69	-68		
0	2	32	299	227	0	5	33	14	130	149	1	-2	16	108	105	1	5	15	61	-56	
0	2	33	480	220	0	5	34	15	58	1	-2	17	49	39	1	5	16	88	-89		
0	2	34	522	455	0	5	35	16	73	71	1	-2	18	78	84	1	5	20	43	-34	
0	2	35	6	797	799	0	5	36	17	122	-125	1	-2	20	52	65	1	5	22	35	-44
0	2	36	7	762	684	0	5	37	2	179	-223	1	-2	22	76	80	1	5	24	49	-62
0	2	37	8	575	543	0	5	38	5	125	156	1	-2	24	44	46	1	5	3	344	334
0	2	38	9	205	142	0	5	39	7	87	72	1	-2	25	35	28	1	5	5	88	80
0	2	39	10	231	247	0	5	40	9	65	75	1	-2	26	21	-19	1	5	6	316	289
0	2	40	12	452	451	0	5	41	10	60	-68	1	-2	27	57	-22	1	5	7	160	140
0	2	41	13	235	189	0	5	42	11	38	51	1	-2	28	18	-39	1	5	8	306	285
0	2	42	14	290	301	0	5	43	12	35	36	1	-2	29	13	-139	1	5	9	115	95
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0	2	44	16	571	547	0	5	45	1	7	54	1	-2	31	3	20	1	5	11	30	30
0	2	45	17	219	-215	0	5	46	8	34	21	1	-2	32	126	-141	1	5	12	224	209
0	2	46	18	288	311	0	5	47	10	9	59	1	-2	33	88	-99	1	5	13	60	-51
0	2	47	19	76	-81	0	5	48	11	56	-47	1	-2	34	138	-139	1	5	14	210	226
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0	2	53	25	102	98	0	5	54	25	35	-47	1	-2	40	12	29	1	5	20	143	152
0	2	54	26	81	91	0	5	55	26	23	-79	1	-2	41	13	32	1	5	22	129	146
0	2	55	27	28	91	0	5	56	27	39	-39	1	-2	42	14	30	1	5	23	30	-31
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0	2	57	29	226	259	0	5	58	29	37	1	-2	44	2	57	1	6	1	91	-103	
0	2	58	30	3	142	0	5	59	61	1	-2	45	3	57	1	6	2	39	37		
0	2	59	31	426	-211	1	1	0	1	45	40	1	-2	46	1	6	3	68	-70		
0	2	60	32	536	547	1	1	1	2	78	82	1	-2	47	5	59	1	6	6	59	-50
0	2	61	33	7	417	408	1	1	1	4	54	-50	1	-2	48	6	7	30	-31		
0	2	62	34	8	320	-282	1	1	1	7	40	-29	1	-2	49	7	8	81	-87		
0	2	63	35	9	4	42	67	47	1	-2	50	65	66	1	-2	51	10	110	-118		
0	2	64	36	10	170	-156	1	1	1	8	87	-76	1	-2	52	10	115	-94			
0	2	65	37	11	11	412	412	412	1	1	9	62	-68	1	-2	53	12	154	-148		
0	2	66	38	12	11	412	412</														

TABLE 4 (Continued)

<i>h</i>	<i>k</i>	<i>l</i>	<i>F</i> _o	<i>F</i> _c	<i>h</i>	<i>k</i>	<i>l</i>	<i>F</i> _o	<i>F</i> _c	<i>h</i>	<i>k</i>	<i>l</i>	<i>F</i> _o	<i>F</i> _c	<i>h</i>	<i>k</i>	<i>l</i>	<i>F</i> _o	<i>F</i> _c			
2	3	4	56	-42	2	-5	23	56	-58	3	-1	16	279	264	3	4	1	276	246			
2	3	5	172	183	2	6	0	207	194	3	-1	18	136	142	2	-6	7	17	97	-81		
2	3	6	137	-122	2	6	1	27	-49	3	-1	19	110	-64	18	96	-96	18	96	-96		
2	3	7	123	126	2	6	2	189	156	3	-1	20	253	249	20	-6	7	17	90	106		
2	3	8	135	-134	2	6	4	166	139	3	-1	21	46	34	21	141	150	19	64	42		
2	3	9	46	51	2	6	5	63	58	3	-1	22	271	287	20	-6	7	17	90	-114		
2	3	10	86	-91	2	6	6	173	147	3	-1	24	179	197	21	-6	7	17	90	73		
2	3	11	16	23	2	6	7	90	82	3	-1	26	117	112	22	-6	7	17	90	109		
2	3	12	69	-72	2	6	8	213	205	3	-1	27	127	124	23	-6	7	17	90	-53		
2	3	14	50	-54	2	6	9	83	71	3	-1	28	43	57	24	-6	7	17	90	-57		
2	3	17	57	48	2	6	10	256	283	3	-1	2	1	180	-70	25	-6	7	17	90	70	
2	3	18	50	-53	2	6	11	54	49	3	-1	2	2	90	70	26	-6	7	17	90	38	
2	3	19	46	43	2	6	12	255	273	3	-1	2	3	482	504	27	-6	7	17	90	43	
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2	3	25	74	82	2	6	10	57	-69	3	-1	2	9	671	781	32	-6	8	3	95	-89	
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2	-3	1	41	9	2	6	14	51	-57	3	-1	2	12	73	-58	35	-6	8	1	8	61	-43
2	-3	3	26	-35	2	6	16	38	-36	3	-1	2	13	587	657	36	-6	4	1	11	50	-37
2	-3	4	40	-34	2	7	1	180	141	3	-1	2	15	476	529	37	-6	4	1	12	85	-55
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2	-3	7	110	-118	2	7	3	212	228	3	-1	2	17	504	563	39	-6	4	1	19	51	-57
2	-3	9	117	-119	2	7	4	23	-38	3	-1	2	18	124	66	40	-6	4	1	21	63	-50
2	-3	11	126	123	2	7	5	213	240	3	-1	2	19	487	520	41	-6	4	1	3	103	108
2	-3	13	124	18	2	7	7	206	192	3	-1	2	20	71	-52	42	-6	4	1	12	59	56
2	-3	13	112	-110	2	7	8	27	28	3	-1	2	21	96	89	43	-6	4	1	13	74	57
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2	-3	17	86	-94	2	7	11	130	117	3	-1	2	24	54	33	45	-6	4	2	0	42	-35
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2	-3	21	57	-53	2	7	15	119	111	3	-1	2	26	46	28	47	-6	4	2	9	51	39
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2	-4	6	121	111	2	8	5	76	-79	3	-1	2	31	311	367	52	-6	4	2	14	28	84
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2	-4	9	52	-35	2	8	8	106	82	3	-1	2	34	119	70	55	-6	4	2	17	36	37
2	-4	10	69	61	2	8	9	109	-98	3	-1	2	35	431	-440	56	-6	4	2	18	86	56
2	-4	12	55	50	2	8	10	512	524	3	-1	2	36	307	-217	57	-6	4	2	19	35	35
2	-4	14	42	47	2	8	11	416	482	3	-1	2	37	174	163	58	-6	4	2	20	42	-35
2	-4	15	66	78	2	8	12	533	705	3	-1	2	38	219	170	59	-6	4	2	21	74	-63
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2	-4	18	79	100	2	8	14	1056	1209	3	-1	2	40	111	-82	61	-6	4	2	23	126	-73
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2	-4	24	67	78	2	8	17	405	412	3	-1	2	43	130	-217	64	-6	4	2	26	125	-23
2	-4	26	50	53	2	8	18	275	-286	3	-1	2	44	174	163	65	-6	4	2	27	228	-123
2	-4	28	123	-139	2	8	19	175	165	3	-1	2	45	185	201	66	-6	4	2	28	177	-80
2	-4	4	205	-199	2	8	20	302	-279	3	-1	2	46	158	162	67	-6	4	2	29	109	-87
2	-4	5	41	-36	2	8	21	306	283	3	-1	2	47	122	120	68	-6	4	2	30	99	-87
2	-4	6	231	-241	2	8	22	109	51	3	-1	2	48	68	-59	69	-6	4	2	31	22	-65
2	-4	7	63	-51	2	8	23	215	160	3	-1	2	49	94	111	70	-6	4	2	32	103	-53
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2	-4	10	144	-156	2	8	26	241	196	3	-1	2	52	384	413	73	-6	4	2	35	285	-304
2	-4	12	122	-132	2	8	27	243	134	3	-1	2	53	724	808	74	-6	4	2	36	220	-229
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2	-4	16	115	-119	2	8	29	266	65	3	-1	2	55	462	659	76	-6	4	2	38	311	-311
2	-4	17	46	40	2	8	30	216	123	3	-1	2	56	61	-13	77	-6	4	2	39	270	-264
2	-4	18	85	-96	2	8	31	154	206	3	-1	2	57	164	124	78	-6	4	2	40	370	-382
2	-4	19	38	21	2	8	32	241	196	3	-1	2	58	304	326	79	-6	4	2	41	58	64
2	-4	20	78	-84	2	8	33	267	186	3	-1	2	59	274	195	80	-6	4	2	42	216	195
2	-4	22	39	-34	2	8	34	246	143	3	-1	2	60	34	47	81	-6	4	2	43	66	47
2	-4	5	126	299	2	8	35	196	143	3	-1	2	61	55	45	82	-6	4	2	44	133	108
2	-4	6	81	71	2	8	36	171	119	3	-1	2	62	176	-153	83	-6	4	2	45		

TABLE 4 (Continued)

<i>h</i>	<i>k</i>	<i>l</i>	<i>F</i> _o	<i>F</i> _c	<i>h</i>	<i>k</i>	<i>l</i>	<i>F</i> _o	<i>F</i> _c	<i>h</i>	<i>k</i>	<i>l</i>	<i>F</i> _o	<i>F</i> _c	<i>h</i>	<i>k</i>	<i>l</i>	<i>F</i> _o	<i>F</i> _c		
4	-3	13	223	243	5	-1	10	92	-103	5	-4	17	137	-152	6	-1	11	522	546		
4	-3	14	51	-41	5	-1	11	49	-45	5	-4	19	116	-127	6	-1	12	149	-142		
4	-3	15	171	186	5	-1	12	101	-103	5	-4	21	55	-64	6	-1	13	415	445		
4	-3	17	130	145	5	-1	14	41	-39	5	-5	6	50	40	6	-1	14	76	-73		
4	-3	19	132	138	5	2	1	84	80	5	-5	8	109	87	6	-1	15	124	139		
4	-3	21	99	100	5	2	3	133	117	5	-5	10	152	155	6	-1	16	83	84		
4	-3	23	90	89	5	2	4	39	-43	5	-5	12	174	176	6	-1	17	70	77		
4	-3	25	83	92	5	2	5	151	142	5	-5	14	127	130	6	-1	18	108	-113		
4	-3	27	58	-156	5	2	6	95	-88	5	-5	16	99	92	6	-1	19	253	247		
4	-4	2	126	-98	5	2	7	91	87	5	-5	18	62	55	6	-1	20	64	-70		
4	-4	4	105	-74	5	2	8	59	-50	5	-5	20	51	46	6	-1	21	123	130		
4	-4	6	96	-68	5	2	11	99	86	5	-5	21	112	91	6	-1	23	124	120		
4	-4	8	69	-52	5	2	12	45	36	5	-5	4	122	-133	6	-1	25	69	87		
4	-4	10	43	-46	5	2	13	121	120	5	-5	5	96	70	6	-1	27	86	74		
4	-4	12	51	-34	5	2	14	115	113	5	-5	6	116	-127	6	-2	0	20	55		
4	-4	2	323	307	5	2	15	150	176	5	-5	10	100	-121	6	-2	1	222	196		
4	-4	3	68	26	5	2	16	117	116	5	-5	12	140	-159	6	-2	2	2	707		
4	-4	4	444	443	5	2	17	96	111	5	-5	14	88	-104	6	-2	3	64	54		
4	-4	6	438	458	5	2	18	59	56	5	-5	15	59	56	6	-2	4	596	614		
4	-4	7	41	-32	5	2	19	82	91	5	-5	16	65	-69	6	-2	5	149	-106		
4	-4	8	327	386	5	2	20	40	47	5	-5	18	116	-112	6	-2	6	708	803		
4	-4	10	284	328	5	2	21	63	76	5	-5	20	74	-61	6	-2	8	644	695		
4	-4	12	270	319	5	2	23	46	56	5	-5	21	39	-52	6	-2	9	217	-218		
4	-4	14	235	285	5	2	25	46	54	5	-5	22	134	128	6	-2	10	239	215		
4	-4	16	231	277	5	2	27	42	49	5	-5	24	90	-91	6	-2	12	279	239		
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4	-4	26	61	69	5	2	5	135	-119	5	-5	29	131	122	6	-2	17	172	132		
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4	-4	7	45	-33	5	2	12	50	-47	5	-5	34	41	-44	6	-2	24	177	158		
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4	-4	21	257	261	5	2	25	72	-78	5	-5	40	157	124	6	-2	30	132	-113		
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4	-4	9	291	334	5	3	3	136	105	5	-5	45	128	-71	6	-2	35	215	263		
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4	-4	13	292	351	5	3	5	12	106	5	-5	47	107	155	6	-2	37	238	486		
4	-4	15	106	-85	5	3	6	8	175	182	5	-5	48	103	89	6	-2	38	303	270	
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4	-4	17	205	250	5	3	9	13	49	48	5	-5	51	9	454	431	6	-2	41	121	-118
4	-4	19	155	173	5	3	10	14	111	95	5	-5	52	10	647	779	6	-2	42	13	118
4	-4	20	52	55	5	3	11	59	-76	5	-5	53	11	91	-63	6	-2	43	15	41	
4	-4	21	89	81	5	3	12	56	-59	5	-5	54	12	691	772	6	-2	44	238	255	
4	-4	23	57	54	5	3	13	47	-41	5	-5	55	13	143	-156	6	-2	45	17	166	
4	-4	6	71	-53	5	3	14	56	-59	5	-5	56	14	463	533	6	-2	46	281	-85	
4	-4	6	35	-24	5	3	15	87	76	5	-5	57	15	281	-253	6	-2	47	303	326	
4	-4	6	149	145	5	3	16	207	-226	5	-5	58	16	237	282	6	-2	48	19	311	
4	-4	6	129	122	5	3	17	90	90	5	-5	59	17	211	205	6	-2	49	219	236	
4	-4	6	9	82	78	5	3	18	163	-178	5	-5	60	18	145	-173	6	-2	50	24	252
4	-4	6	10	139	147	5	3	19	104	100	5	-5	61	19	128	200	6	-2	51	24	185
4	-4	6	11	46	46	5	3	20	128	-137	5	-5	62	20	170	165	6	-2	52	25	55
4	-4	6	12	131	133	5	3	21	33	34	5	-5	63	22	200	102	6	-2	54	27	32
4	-4	6	13	36	48	5	3	22	103	-116	5	-5	64	23	109	46	6	-2	55	28	17
4	-4	6	14	56	52	5	3	23	104	-104	5	-5	65	24	43	31	6	-2	56	3	174
4	-4	7	1	29	-17	5	3	24	18	-87	5	-5	66	25	817	82	6	-2	57	19	-139
4	-4	7	3	54	-46	5	3	25	77	-75	5	-5	67	26	171	45	6	-2	58	3	270
4	-4	7	7	53	-53	5	3	26	63	-65	5	-5	68	27	298	215	6	-2	59	3	253
4	-4	8	10	39	45	5	3	27	64	-69	5	-5	69	28	125	-81	6	-2	60	3	248
4	-4	8	11	47	-37	5	3	28	49	-51	5	-5	70	29	374	364	6	-2	61	3	246
4	-4	9	12	44	47	5	3	29	155	159	5	-5	71	30	214	229	6	-2	62	3	245
4	-4	9	13	46	-44	5	3	30	138	123	5	-5	72	31	45	31	6	-2	63	3	244
4	-4	9	14	59	9	5	31	5	114	106	5	-5	73	32	12	75	6	-2	64	3	243
4	-4	9	15	50	-42	5	3	32	7	105	90	5	-								

TABLE 4 (Continued)

<i>h</i>	<i>k</i>	<i>l</i>	F_o	F_c	<i>h</i>	<i>k</i>	<i>l</i>	F_o	F_c	<i>h</i>	<i>k</i>	<i>l</i>	F_o	F_c	<i>h</i>	<i>k</i>	<i>l</i>	F_o	F_c		
6	6	16	35	33	7	-4	1	263	280	8	6	4	67	63	9	1	12	290	-273		
6	-6	6	63	61	7	-4	2	98	66	8	6	6	74	76	9	1	13	129	110		
6	-6	7	176	147	7	-4	3	292	314	8	6	8	60	53	9	1	14	334	349		
6	-6	8	79	-71	7	-4	5	338	362	8	6	10	40	32	9	1	16	253	240		
6	-6	9	168	175	7	-4	7	116	-103	8	-1	3	71	-86	9	1	17	148	-145		
6	-6	10	182	177	7	-4	7	361	380	8	-1	5	45	-45	9	1	18	199	207		
6	-6	12	257	286	7	-4	8	109	-107	8	-1	7	86	-83	9	1	22	113	107		
6	-6	14	68	-59	7	-4	9	301	317	8	-1	9	125	-123	9	1	24	97	100		
6	-6	12	162	-160	7	-4	10	51	-59	8	-1	11	110	-108	9	1	26	64	87		
6	-6	16	48	-36	7	-4	11	201	189	8	-1	13	76	-80	9	2	1	492	486		
6	-6	19	52	-36	7	-4	13	202	201	8	-1	15	82	-76	9	2	2	158	156		
6	-7	1	80	72	7	-4	14	33	32	8	-1	17	58	-57	9	2	3	485	399		
6	-7	3	55	27	7	-4	15	238	254	8	-2	0	153	-159	9	2	4	270	220		
6	-7	5	107	119	7	-4	16	68	46	8	-2	4	138	-150	9	2	6	117	103		
6	-7	7	120	139	7	-4	17	222	225	8	-2	8	166	-186	9	2	7	371	340		
7	0	5	55	63	7	-4	19	175	170	8	-2	10	178	-182	9	2	9	254	269		
7	0	7	65	82	7	-4	21	149	140	8	-2	12	152	-166	9	2	10	313	-367		
7	0	9	75	80	7	-4	23	110	104	8	-2	14	128	-133	9	2	11	124	115		
7	0	11	83	86	7	-4	24	59	62	8	-2	16	106	-103	9	2	12	283	275		
7	0	12	78	-63	7	-4	25	71	75	8	-2	17	41	53	9	2	14	264	270		
7	0	15	83	77	7	-5	2	262	244	8	-2	18	60	-55	9	2	15	163	177		
7	1	1	30	-31	7	-5	3	217	-188	8	-2	20	42	-46	9	2	17	128	114		
7	1	7	54	-56	7	-5	4	310	307	8	-2	22	64	-52	9	2	19	117	108		
7	1	8	62	-55	7	-5	5	265	-232	8	-3	2	284	-320	9	2	21	116	135		
7	2	1	113	-107	7	-5	6	293	310	8	-3	2	80	-72	9	2	23	126	140		
7	2	2	112	-113	7	-5	7	182	-170	8	-3	3	138	-118	9	2	24	167	157		
7	2	2	5	92	-88	7	-5	8	239	238	8	-3	4	153	-167	9	2	14	244	244	
7	2	2	7	56	-53	7	-5	10	147	143	8	-3	5	163	-167	9	2	15	85	87	
7	2	2	10	53	67	7	-5	12	89	88	8	-3	6	94	-79	9	2	16	213	215	
7	2	3	67	-63	7	-5	13	84	-66	8	-3	7	154	-168	9	2	17	170	-174		
7	3	2	74	-67	7	-5	14	149	142	8	-3	8	61	-53	9	2	18	87	84		
7	3	3	4	43	-30	7	-5	15	126	170	8	-3	9	170	-162	9	2	19	106	-99	
7	3	6	77	-61	7	-5	16	100	114	8	-3	11	119	-115	9	2	20	148	140		
7	3	10	124	-126	7	-5	18	113	114	8	-3	13	128	-122	9	2	22	160	180		
7	3	12	127	-140	7	-5	19	90	90	8	-3	14	68	-41	9	3	13	86	90		
7	3	13	59	-64	7	-5	20	68	53	8	-3	15	120	-113	9	3	14	281	265		
7	3	14	103	-104	7	-5	21	95	95	8	-3	17	88	-95	9	3	16	260	246		
7	3	15	59	-72	7	-5	22	38	33	8	-3	18	80	-85	9	3	18	153	138		
7	3	18	146	-52	7	-5	24	64	58	8	-3	19	64	-74	9	3	20	141	129		
7	4	1	106	-108	7	-6	2	52	52	8	-3	21	74	-81	9	3	21	104	95		
7	4	4	101	-105	7	-6	3	158	130	8	-3	22	70	-81	9	3	22	104	116		
7	4	4	56	-83	7	-6	7	96	79	8	-3	23	90	-98	9	3	23	13	59		
7	4	4	47	-49	7	-6	8	229	201	8	-3	24	85	-105	9	3	24	275	256		
7	4	7	91	-86	7	-6	9	196	171	8	-3	25	151	-131	9	3	25	281	259		
7	4	9	91	-78	7	-6	10	161	152	8	-3	26	143	-135	9	4	1	229	236		
7	4	11	90	-88	7	-6	11	254	263	8	-3	27	72	-66	9	4	17	216	212		
7	4	13	81	-64	7	-6	13	197	194	8	-3	28	152	-170	9	4	18	141	146		
7	4	15	45	-34	7	-6	14	81	-72	8	-3	29	35	-41	9	4	19	218	209		
7	4	4	45	54	7	-6	15	149	148	8	-3	30	83	-86	9	4	20	180	164		
7	4	1	25	-32	7	-6	17	129	120	8	-3	31	47	-44	9	4	21	141	146		
7	4	6	43	50	7	-6	19	91	91	8	-3	32	14	-85	9	4	22	249	259		
7	4	8	120	133	7	-6	22	70	-57	8	-3	33	15	-39	9	4	23	121	117		
7	4	9	38	39	8	0	0	43	-72	8	-3	34	16	-124	9	4	17	124	-75		
7	4	10	145	156	8	0	0	2	82	-82	8	-3	35	17	-109	9	4	18	124	-116	
7	4	12	113	116	8	0	0	4	88	-88	8	-3	36	12	-109	9	4	19	168	160	
7	4	14	72	73	8	0	0	5	82	-82	8	-3	37	72	-51	9	4	20	176	170	
7	4	16	64	61	8	0	0	8	82	-82	8	-3	38	116	-123	9	4	21	285	295	
7	4	18	67	69	8	0	0	10	64	-53	8	-3	39	94	-85	9	4	22	126	121	
7	4	20	50	63	8	0	0	12	101	109	8	-3	40	159	-153	9	4	23	124	119	
7	4	22	58	66	8	0	0	2	121	109	8	-3	41	112	-124	9	4	24	229	236	
7	4	21	211	246	8	0	0	2	124	108	8	-3	42	7	-98	9	4	25	104	-116	
7	4	2	41	31	8	0	0	6	61	-48	8	-3	43	9	-65	9	4	26	81	55	
7	2	3	223	235	8	0	0	12	103	92	8	-3	44	12	-44	9	4	27	180	190	
7	2	4	112	104	8	0	0	13	151	165	8	-3	45	2	242	9	4	28	300	321	
7	2	5	214	215	8	0	0	15	105	114	8	-3	46	4	306	9	4	29	323	313	
7	2	6	38	30	8	0	0	3	7	131	105	8	-3	47	6	335	9	4	30	135	137
7	2	7	201	211	8	0	0	3	19	72	70	8	-3	48	6	431	9	4	31	119	136
7	2	8	38	35	8	0	0	11	66	78	8	-3	49	14	-56	9	4	32	161	120	
7	2	9	239	254	8	0	0	11	183	204	8	-3	50	7	-89	9	4	33	847	487	
7	2	10	35	25	8	0	0	12	103	92	8	-3	51	2	240	9	4	34	252	255	
7	2	11	196	222	8	0	0	13	151	165	8	-3	52	3	156	9	4	35	153	154	
7	2	12	38	-47	8	0	0	15	105	114	8	-3	53	11	-97	9	4	36	300	321	
7	2	13	183	191	8	0	0	17	88	94	8	-3	54	12	162	9	4	37	323	313	
7	2	14	51	-56	8	0	0	19	72	70	8	-3	55	13	-158	9	4	38	452	468	
7	2	15	152	166	8	0	0	21	55	55	8	-3	56	14	145	9	4	39	305	367	
7	2	16	73	-88	8	0	0	23	173	173	8	-3	57	15	231	9	4	40	447	487	
7	2	17	151	150	8</td																

TABLE 4 (Continued)

<i>h</i>	<i>k</i>	<i>l</i>	<i>F</i> _o	<i>F</i> _c	<i>h</i>	<i>k</i>	<i>l</i>	<i>F</i> _o	<i>F</i> _c	<i>h</i>	<i>k</i>	<i>l</i>	<i>F</i> _o	<i>F</i> _c	<i>h</i>	<i>k</i>	<i>l</i>	<i>F</i> _o	<i>F</i> _c								
10	-1	9	140	124	10	-4	2	196	211	10	-6	16	184	186	11	-4	5	107	-119	12	-1	5	477	479			
10	-1	10	60	-67	10	-4	3	39	48	10	-6	18	113	118	11	-4	7	91	-98	12	-1	7	439	390			
10	-1	11	135	126	10	-4	4	212	212	10	-6	20	81	87	11	-4	8	93	75	12	-1	9	291	239			
10	-1	13	110	102	10	-4	5	41	49	10	-7	6	117	-89	11	-4	10	99	85	12	-1	11	278	237			
10	-1	15	82	77	10	-4	6	236	260	10	-7	7	232	233	11	-5	2	71	-79	12	-1	13	207	199			
10	-1	17	76	72	10	-4	7	31	-38	10	-7	8	63	-61	11	-5	4	90	-94	12	-1	15	214	201			
10	-1	19	76	77	10	-4	179	194		10	-7	9	214	229	11	-5	5	52	-55	12	-1	17	251	214			
10	-1	23	55	68	10	-4	9	162	-133	10	-7	10	70	-55	11	-5	6	118	-125	12	-1	19	156	179			
10	-2	2	207	229	10	-4	10	143	-147	10	-7	11	202	231	11	-5	7	100	-81	12	-1	21	165	174			
10	-2	3	81	-84	10	-4	12	132	129	10	-7	13	181	195	11	-5	8	95	-78	12	-1	23	193	200			
10	-2	4	242	234	10	-4	13	91	91	10	-7	15	147	141	12	0	1	310	-165	12	-2	4	321	375			
10	-2	5	48	-64	10	-4	14	93	89	10	-7	17	120	121	12	0	2	707	485	12	-2	5	100	-102			
10	-2	6	258	246	10	-4	16	85	90	11	-1	2	125	-119	12	0	4	454	459	12	-2	6	374	379			
10	-2	8	250	247	10	-4	18	91	76	11	-1	4	118	-113	12	0	5	185	152	12	-2	8	278	237			
10	-2	10	245	249	10	-4	20	85	80	11	-1	5	43	40	12	0	6	382	340	12	-2	9	248	173			
10	-2	12	233	241	10	-4	22	69	70	11	-1	6	124	-116	12	0	7	142	90	12	-2	10	273	247			
10	-2	14	186	200	10	-4	24	62	2	11	-1	8	96	-88	12	0	8	357	338	12	-2	12	357	286			
10	-2	16	130	145	10	-5	1	290	313	11	-1	10	90	-67	12	0	9	129	105	12	-2	15	137	151			
10	-2	18	126	125	10	-5	3	282	297	11	-1	12	90	-64	12	0	10	200	197	12	-2	18	161	157			
10	-2	20	98	96	10	-5	5	281	288	11	-2	3	131	-146	12	0	12	160	157	12	-2	20	125	137			
10	-2	22	69	59	10	-5	6	100	83	11	-2	4	30	37	12	0	14	165	215	12	-2	32	136	115			
10	-3	1	228	216	10	-5	7	225	227	11	-2	5	140	-152	12	0	16	211	208	12	-3	4	85	88			
10	-3	2	72	62	10	-5	8	96	70	11	-2	7	114	-110	12	0	18	156	150	12	-3	6	200	-154			
10	-3	3	183	183	10	-5	11	130	-103	11	-2	9	134	-123	12	0	20	165	170	12	-3	9	347	339			
10	-3	4	149	136	10	-5	12	93	84	11	-2	11	148	-143	12	1	3	378	242	12	-3	11	433	460			
10	-3	5	161	159	10	-5	15	65	77	11	-2	13	130	-126	12	1	5	262	274	12	-3	13	384	381			
10	-3	6	185	181	10	-5	17	102	94	11	-2	15	88	-80	12	1	6	151	121	12	-3	15	136	138			
10	-3	7	179	190	10	-5	19	88	87	11	-2	17	69	-56	12	1	7	257	290	12	-3	19	181	175			
10	-3	8	105	95	10	-5	21	83	94	11	-2	19	68	-58	12	1	9	131	151	12	-4	0	227	216			
10	-3	9	234	238	10	-5	23	76	99	11	-3	0	59	-62	12	1	11	251	281	12	-4	3	107	95			
10	-3	10	46	42	10	-6	2	118	117	11	-3	2	40	-42	12	1	13	165	182	12	-4	4	295	304			
10	-3	11	270	300	10	-6	3	124	-108	11	-3	4	43	-49	12	1	15	160	170	12	-4	5	207	182			
10	-3	12	34	26	10	-6	4	119	105	11	-3	5	71	-68	12	1	17	217	268	12	-4	7	265	-239			
10	-3	13	257	276	10	-6	5	106	-92	11	-3	6	57	-57	12	1	19	142	176	12	-4	9	233	-212			
10	-3	15	190	194	10	-6	6	200	179	11	-3	8	55	-57	12	3	1	286	226	12	-4	10	131	156			
10	-3	17	132	147	10	-6	8	228	234	11	-3	10	143	-156	12	3	5	131	131	12	-4	15	136	140			
10	-3	19	130	145	10	-6	10	224	201	11	-3	12	158	-184	12	3	6	160	-209	12	-4	16	156	167			
10	-3	21	110	119	10	-6	11	53	-62	11	-3	14	117	-134	12	3	9	156	225	12	-4	19	120	-121			
10	-3	23	87	93	10	-6	12	224	228	11	-3	16	69	-70	12	-1	1	727	516	12	-5	5	283	276			
10	-3	25	60	65	10	-6	13	63	-81	11	-4	1	79	-84	12	-1	2	312	-186	12	-5	6	273	268			
10	-4	0	212	233	10	-6	14	224	249	11	-4	3	80	-76	12	-1	3	425	440	12	-5	7	227	191			
																							12	-5	18	120	-129

in aquo-ligands appear to be pyramidal. In the chelate ring, there is a steady decrease in the angle Cu-OR-C for R = H, Me, Ph, accompanied by an increase in CuOC angle at the carboxylate oxygen. The C-O-R angles follow a pattern consistent with expected intramolecular repulsions for the phenoxyacetate, methoxyacetate, and glycolate, but have widely different values for the lactate and 2-hydroxy-2-methylpropionate, and these variations are necessarily reflected in the variations in the CuOR angle.

In the three 2-hydroxycarboxylates, each hydroxyl group is not only bound to the copper atom, but is also hydrogen-bonded to the non-chelated carboxylate oxygen atom of another complex. Although variations in the O-Cu and O-H bond lengths could be attributed to experimental error, it is noteworthy that the hydrogen bonds appear to shorten as the Cu-OH bonds shorten. A similar effect is observable for the water molecules in the hexa-aquocopper(II) ion in Tutton's salt.²² Such a co-operative strengthening of bonds would be expected from an electronic viewpoint.

The carboxylate group usually forms hydrogen-bonds only at the non-chelating oxygen atom. When this happens there is one hydrogen-bond approximately parallel to the carbon-carbon bond of the chelate ring and coplanar with it. Some non-chelating oxygen atoms (in phenoxyacetate, methoxyacetate, and lactate) form a second hydrogen-bond which is near to the carboxylate plane and roughly parallel to the O ··· O vector in the carboxylate group, but others (in lactate and glycolate) form a long contact to a copper atom in another molecule.

The hydration of the 2-hydroxycarboxylate complexes

increases from 0 to 1 to 2 in the sequence glycolate, lactate, 2-hydroxy-2-methylpropionate, i.e., in parallel with the substitution of methyl groups on the 2-carbon atom. In the 2-hydroxy-2-methylpropionate, adjacent chelate rings are linked into a ribbon structure by hydrogen-bonds from O(2) to O(3). The carboxy-group does not form a second hydrogen-bond (see above). The inclination of the two methyl substituents in each chelate ring keeps superimposed planes of chelate rings sufficiently far apart to accommodate two *trans* water molecules on each copper ion (see Figure 12). The hypothetical conversion of the 2-hydroxy-2-methylpropionate into the lactate involves the removal of one methyl group from each chelate ring, and the formation of large holes in the structure. The consequent closer approach of superimposed planes of chelate rings is such that there is now only room for one aquo-ligand on each copper ion. One lactate ligand only of each complex forms the second type of hydrogen-bond (see above), and a double hydrogen-bond ribbon is formed. The hypothetical conversion of lactate into glycolate involves the removal of the remaining two methyl substituents from the rings, and again the formation of holes. If the original ribbon structure were to be retained, a reasonably close packed structure could not be formed with axial Cu-O contacts of less than about 2.9 Å. The original O(2) ··· O(3) two-dimensional sheet with shorter Cu-O(2) contacts of 2.54 Å results.

The formation of two hydrogen-bonds by the non-chelating oxygen atom suggests *sp*² hybridisation of the

659.

²² H. Montgomery and E. C. Lingafelter, *Acta Cryst.*, 1966, **20**,

TABLE 5

Observed structure amplitudes and calculated structure factors for diaquobis(methoxyacetato)copper(II)

<i>h k l</i>	<i>F</i> _o	<i>F</i> _c	<i>h k l</i>	<i>F</i> _o	<i>F</i> _c	<i>h k l</i>	<i>F</i> _o	<i>F</i> _c	<i>h k l</i>	<i>F</i> _o	<i>F</i> _c	<i>h k l</i>	<i>F</i> _o	<i>F</i> _c		
0 0 2	268	300	1 -2 5	107	111	1 5 6	90	97	2 -7 8	21	24	3 -3 1	1	46	40	
0 0 4	122	127	1 -2 6	183	172	1 5 7	19	-24	2 -8 0	40	36	3 -3 2	2	239	218	
0 0 6	186	204	1 -2 7	110	116	1 5 8	28	31	2 -8 1	11	-5	3 -3 3	3	48	33	
0 0 8	92	96	1 -2 8	31	39	1 5 10	34	41	2 -8 2	45	39	3 -3 4	4	95	96	
0 0 10	84	96	1 -2 9	69	72	1 6 10	40	-41	2 -8 3	14	-17	3 -3 5	5	56	49	
0 0 12	10	5	1 -2 10	13	-10	1 6 11	70	61	2 -8 4	28	26	3 -3 6	6	89	84	
0 1 12	54	-32	1 -2 11	52	57	1 6 12	34	28	2 -8 5	8	9	3 -3 7	7	19	-18	
0 1 13	156	-111	1 -2 12	17	23	1 6 13	115	115	2 -8 6	42	41	3 -3 8	8	93	98	
0 1 14	49	-54	1 -2 13	0	63	1 6 14	115	115	2 -9 1	38	35	3 -3 9	9	12	-9	
0 1 15	195	187	1 -3 1	204	169	1 6 15	32	39	2 -9 2	9	-6	3 -4 10	1	189	179	
0 1 16	36	-31	1 -3 2	218	193	1 6 16	10	-15	2 -9 3	15	16	3 -4 2	2	29	-27	
0 1 17	114	111	1 -3 3	24	-23	1 6 17	9	21	2 -1 1	1	120	3 -4 3	3	109	103	
0 1 18	41	-48	1 -3 4	223	219	1 6 18	2	71	2 -1 1	1	131	3 -4 4	4	41	31	
0 1 19	89	105	1 -3 5	6	97	1 7 1	4	100	2 -1 1	1	154	3 -4 5	5	84	71	
0 1 20	11	43	1 -3 6	23	15	1 7 2	5	37	2 -1 1	1	176	3 -4 6	6	78	76	
0 2 20	298	333	1 -3 7	8	76	1 7 3	6	55	2 -1 1	1	186	3 -4 7	7	97	89	
0 2 21	39	52	1 -3 8	4	-46	1 7 4	8	19	2 -1 1	1	197	3 -4 8	8	54	58	
0 2 22	2	39	41	1 -3 9	10	51	1 8 1	1	58	2 -1 1	1	200	3 -4 9	9	160	143
0 2 23	3	45	-21	1 -3 11	24	29	1 8 2	2	22	2 -1 1	1	211	3 -5 0	10	21	19
0 2 24	128	131	1 -3 12	12	29	1 8 3	3	49	2 -1 1	1	219	3 -5 1	11	60	-53	
0 2 25	29	-17	1 -3 13	1	15	1 8 4	4	-13	2 -1 1	1	228	3 -5 2	12	48	-50	
0 2 26	192	193	1 -4 2	2	84	1 8 5	5	27	2 -1 1	1	236	3 -5 3	13	94	-78	
0 2 27	7	67	-69	1 -4 3	204	200	1 8 6	6	-7	2 -1 1	1	244	3 -5 4	14	13	21
0 2 28	8	78	80	1 -4 4	90	-78	1 9 1	0	41	2 -1 1	1	253	3 -5 5	15	59	59
0 2 29	9	25	-28	1 -4 5	136	140	1 9 2	1	10	2 -1 1	1	262	3 -5 6	16	55	64
0 2 30	47	62	1 -4 6	6	23	1 9 3	2	52	2 -1 1	1	271	3 -5 7	17	21	-17	
0 2 31	27	42	1 -4 7	4	42	1 9 4	3	68	2 -1 1	1	280	3 -5 8	18	55	-23	
0 2 32	120	131	1 -4 8	8	75	1 9 5	4	208	2 -1 1	1	289	3 -5 9	19	132	131	
0 2 33	194	151	1 -4 9	9	39	1 9 6	5	102	2 -1 1	1	298	3 -5 10	20	63	-59	
0 2 34	129	138	1 -4 10	16	-15	2 0 0	0	118	2 -1 1	1	307	3 -5 11	21	29	26	
0 2 35	123	103	1 -4 11	28	37	2 0 0	0	127	2 -1 1	1	316	3 -5 12	22	6	19	
0 2 36	194	196	1 -5 1	1	62	2 0 0	0	134	2 -1 1	1	325	3 -5 13	23	1	17	
0 2 37	59	46	1 -5 2	194	173	2 0 0	0	143	2 -1 1	1	334	3 -5 14	24	13	11	
0 2 38	9	28	35	1 -5 3	18	15	2 0 0	0	152	2 -1 1	1	343	3 -5 15	25	7	13
0 2 39	50	56	1 -5 4	208	197	2 0 0	0	161	2 -1 1	1	352	3 -5 16	26	6	72	
0 2 40	11	19	27	1 -5 5	28	14	2 0 0	0	170	2 -1 1	1	361	3 -5 17	27	6	62
0 2 41	12	13	16	1 -5 6	76	69	2 0 0	0	179	2 -1 1	1	370	3 -5 18	28	5	56
0 2 42	147	128	1 -5 7	8	39	2 0 0	0	188	2 -1 1	1	379	3 -5 19	29	4	46	
0 2 43	1	58	52	1 -5 8	9	12	2 0 0	0	197	2 -1 1	1	388	3 -5 20	30	3	22
0 2 44	209	203	1 -5 9	10	36	2 0 0	0	206	2 -1 1	1	397	3 -5 21	31	2	19	
0 2 45	181	159	1 -6 0	1	126	2 0 0	0	215	2 -1 1	1	406	3 -5 22	32	1	257	
0 2 46	132	133	1 -6 1	3	125	2 0 0	0	224	2 -1 1	1	415	3 -5 23	33	1	282	
0 2 47	51	-47	1 -6 2	4	14	2 0 0	0	233	2 -1 1	1	424	3 -5 24	34	1	197	
0 2 48	73	77	1 -6 3	5	76	2 0 0	0	242	2 -1 1	1	433	3 -5 25	35	1	210	
0 2 49	11	18	22	1 -6 4	6	19	2 0 0	0	251	2 -1 1	1	442	3 -5 26	36	1	17
0 2 50	170	152	1 -6 5	7	22	2 0 0	0	260	2 -1 1	1	451	3 -5 27	37	1	145	
0 2 51	26	-18	1 -6 6	9	51	2 0 0	0	269	2 -1 1	1	460	3 -5 28	38	1	150	
0 2 52	202	180	1 -6 7	0	75	2 0 0	0	278	2 -1 1	1	469	3 -5 29	39	1	166	
0 2 53	24	-20	1 -6 8	1	20	2 0 0	0	287	2 -1 1	1	478	3 -5 30	40	1	189	
0 2 54	80	69	1 -6 9	2	55	2 0 0	0	296	2 -1 1	1	487	3 -5 31	41	1	210	
0 2 55	25	-21	1 -7 0	3	14	2 0 0	0	305	2 -1 1	1	496	3 -5 32	42	1	145	
0 2 56	105	109	1 -7 1	4	37	2 0 0	0	314	2 -1 1	1	505	3 -5 33	43	1	153	
0 2 57	13	14	1 -7 2	5	22	2 0 0	0	323	2 -1 1	1	514	3 -5 34	44	1	163	
0 2 58	39	37	1 -7 3	6	11	2 0 0	0	332	2 -1 1	1	523	3 -5 35	45	1	172	
0 2 59	10	9	8	1 -7 4	8	41	2 0 0	0	341	2 -1 1	1	532	3 -5 36	46	1	181
0 2 60	34	-34	1 -7 5	1	21	2 0 0	0	350	2 -1 1	1	541	3 -5 37	47	1	190	
0 2 61	43	40	1 -7 6	2	37	2 0 0	0	359	2 -1 1	1	550	3 -5 38	48	1	209	
0 2 62	35	-26	1 -7 7	3	50	2 0 0	0	368	2 -1 1	1	559	3 -5 39	49	1	228	
0 2 63	43	-13	1 -7 8	4	26	2 0 0	0	377	2 -1 1	1	568	3 -5 40	50	1	247	
0 2 64	43	40	1 -7 9	5	24	2 0 0	0	386	2 -1 1	1	577	3 -5 41	51	1	266	
0 2 65	65	79	1 -7 10	6	31	2 0 0	0	395	2 -1 1	1	586	3 -5 42	52	1	285	
0 2 66	73	60	1 -7 11	7	23	2 0 0	0	404	2 -1 1	1	595	3 -5 43	53	1	304	
0 2 67	24	-24	1 -7 12	8	50	2 0 0	0	413	2 -1 1	1	604	3 -5 44	54	1	323	
0 2 68	89	81	1 -7 13	9	34	2 0 0	0	422	2 -1 1	1	613	3 -5 45	55	1	342	
0 2 69	25	19	1 -7 14	10	59	2 0 0	0	431	2 -1 1	1	622	3 -5 46	56	1	361	
0 2 70	56	60	1 -7 15	11	12	2 0 0	0	440	2 -1 1	1	631	3 -5 47	57	1	380	
0 2 71	31	35	1 -7 16	12	48	2 0 0	0	449	2 -1 1	1	640	3 -5 48	58	1	399	
0 2 72	8	68	1 -7 17	13	48	2 0 0	0	458	2 -1 1	1	649	3 -5 49	59	1	418	
0 2 73	3	19	18	1 -7 18	14	48	2 0 0	0	467	2 -1 1	1	658	3 -5 50	60	1	437
0 2 74	6	29	28	1 -7 19	14	112	2 0 0	0	476	2 -1 1	1	667	3 -5 51	61	1	456
0 2 75	1	52	50	1 -7 20	4	176	2 0 0	0	485	2 -1 1	1	676	3 -5 52	62	1	475
0 2 76	2	12	17	1 -7 21	5	129	2 0 0	0	494	2 -1 1	1	685	3 -5 53	63	1	494
0 2 77	217	-212	1 -7 22	7	98	2 0 0	0	503	2 -1 1	1	694	3 -5 54	64	1	513	
0 2 78	281	300	1 -7 23	9	66	2 0 0	0	512	2 -1 1	1	703	3 -5 55	65	1	532	
0 2 79	129	-127	1 -7 24	11	50	2 0 0	0	521	2 -1 1	1	712	3 -5 56	66	1	551	
0 2 80	146	155	1 -7 25	12	25	2 0 0	0	530	2 -1 1	1	721	3 -5 57	67	1	570	
0 2 81	84	76	1 -7 26	13	30	2 0 0	0	539	2 -1 1	1	730	3 -5 58	68	1	589	
0 2 82	7	113	146	1 -7 27	14	98	2 0 0	0	548	2 -1 1	1	739	3 -5 59	69	1	608
0 2 83	30	26	1 -7 28	15	114	2 0 0	0	557	2 -1 1	1	748	3 -5 60	70	1	627	
0 2 84	44	50	1 -7 29	16	43	2 0 0	0	566	2 -1 1	1	757	3 -5 6				

TABLE 5 (Continued)

<i>h</i>	<i>k</i>	<i>l</i>	<i>F</i> _o	<i>F</i> _c	<i>h</i>	<i>k</i>	<i>l</i>	<i>F</i> _o	<i>F</i> _c	<i>h</i>	<i>k</i>	<i>l</i>	<i>F</i> _o	<i>F</i> _c	<i>h</i>	<i>k</i>	<i>l</i>	<i>F</i> _o	<i>F</i> _c						
4	-2	0	240	231	4	-6	3	17	-17	4	4	8	50	54	5	-2	10	17	22	5	1	2	133	118	
4	-2	1	126	-105	4	-6	4	73	73	4	-4	9	8	-11	5	-3	0	116	90	5	1	3	54	35	
4	-2	2	113	114	4	-6	6	51	57	4	-5	3	114	108	5	-3	1	20	-16	5	1	4	69	61	
4	-2	3	21	-10	4	-6	8	13	14	4	-5	4	23	-17	5	-3	2	119	100	5	1	5	46	-49	
4	-2	4	98	93	4	-7	1	55	49	4	-5	5	62	55	5	-3	3	26	-71	5	1	6	26	30	
4	-2	5	33	28	4	-7	2	11	-13	4	-5	6	31	34	5	-3	4	86	83	5	1	7	61	-70	
4	-2	6	53	46	4	-7	3	83	81	4	-5	7	54	56	5	-3	5	25	-28	5	1	8	62	70	
4	-2	7	21	-28	4	-7	5	53	56	4	-6	8	10	10	5	-3	6	90	86	5	1	9	15	8	
4	-2	8	68	69	4	-7	7	25	32	4	-6	9	54	51	5	-3	7	22	11	5	1	10	38	54	
4	-2	9	10	-6	4	-8	0	39	33	4	-6	14	31	26	5	-3	8	45	42	5	2	3	132	116	
4	-2	10	70	92	4	-8	1	17	20	4	-6	6	35	36	5	-3	9	10	3	5	2	4	46	39	
4	-3	1	236	212	4	-8	2	57	55	4	-7	1	51	50	5	-3	10	14	25	5	2	5	77	75	
4	-3	2	123	-112	4	-8	3	8	4	4	-7	3	17	14	4	-4	1	120	100	5	2	6	32	-35	
4	-3	3	112	102	4	-8	4	41	41	4	-7	4	11	13	4	-4	3	73	68	5	2	7	43	42	
4	-3	4	31	25	4	-1	2	48	64	4	-7	5	12	16	4	-4	4	67	61	5	2	8	23	18	
4	-3	5	89	83	4	-1	3	219	219	4	-7	5	178	196	4	-4	5	85	81	5	2	9	45	50	
4	-3	6	56	-57	4	-1	4	13	-5	4	-1	5	153	141	4	-4	7	55	56	5	3	3	52	43	
4	-3	7	32	38	4	-1	6	23	13	4	-1	6	23	13	4	-5	0	145	127	5	3	4	126	124	
4	-3	8	23	-26	4	-1	7	58	60	4	-1	7	58	60	4	-5	1	75	59	5	3	5	18	10	
4	-3	9	62	76	4	-1	8	50	-55	4	-2	6	63	-62	4	-5	2	94	77	5	5	6	52	53	
4	-4	10	190	158	4	-1	9	35	45	4	-2	7	76	70	4	-5	3	22	14	5	5	7	15	13	
4	-4	11	91	-76	4	-1	11	20	18	4	-2	8	12	4	4	8	87	79	5	5	8	25	22		
4	-4	12	2140	121	4	-2	3	59	-66	4	-2	4	138	135	4	-2	10	22	23	5	6	6	86	73	
4	-4	13	45	38	4	-2	5	30	-31	4	-2	5	30	-31	4	-2	11	11	-13	5	7	21	22	118	
4	-4	14	132	115	4	-2	6	67	68	4	-2	7	172	174	4	-2	13	26	25	5	4	3	123	105	
4	-4	15	53	47	4	-2	7	27	-30	4	-2	8	73	53	4	-2	1	49	42	5	4	6	16	-17	
4	-4	16	71	71	4	-2	8	59	69	4	-2	9	99	80	4	-2	2	65	65	5	4	7	27	19	
4	-4	17	39	31	4	-2	9	49	-60	4	-2	10	32	34	4	-2	6	75	62	5	4	4	26	48	
4	-4	18	9	30	32	4	-2	9	49	-60	4	-2	11	22	-25	4	-2	7	50	67	5	5	15	-19	
4	-4	19	10	37	47	4	-2	10	32	34	4	-2	11	87	65	4	-2	6	6	10	5	31	30		
4	-4	20	1	64	54	4	-3	3	133	121	4	-2	11	27	32	4	-2	7	28	32	7	7	3	28	
4	-4	21	5	23	-47	4	-3	4	30	-31	4	-2	10	25	34	4	-2	0	44	40	1	33	29	27	
4	-4	22	5	122	99	4	-3	5	53	46	4	-2	1	65	53	4	-2	2	15	11	3	28	27	27	
4	-4	23	46	38	4	-3	6	78	81	4	-2	2	13	-9	4	-2	3	15	18	5	6	5	21	24	
4	-4	24	5	108	101	4	-3	7	73	73	4	-2	3	96	79	4	-2	4	41	40	5	7	0	39	38
4	-4	25	6	6	3	4	8	19	11	4	-2	3	16	17	4	-2	5	14	-20	5	7	1	10	13	
4	-4	26	7	45	42	4	-4	9	45	50	4	-2	3	59	53	4	-2	6	29	41	5	7	0	39	38
4	-4	27	8	21	20	4	-4	3	55	-58	4	-2	4	82	66	4	-2	1	51	51	5	8	2	12	12
4	-4	28	9	19	15	4	-4	4	5	29	4	-2	5	78	74	4	-2	1	51	51	5	8	2	12	12
4	-4	29	0	29	-22	4	-4	5	6	54	4	-2	6	37	35	4	-2	2	15	11	5	8	3	21	24
4	-4	30	1	46	-35	4	-4	6	54	51	4	-2	7	61	63	4	-2	3	34	34	5	7	1	10	13

oxygen orbitals. Using the terminology of Martin and Waterman,²³ the hydrogen-bonds are in both the *syn*- and *anti*-positions, and the copper-oxygen bond in the chelate ring is *anti* (see Figure 15). When the non-

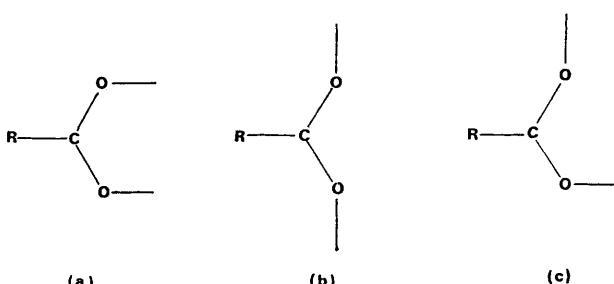


FIGURE 15 Carboxylate configurations (a) *syn,syn*, (b) *anti,anti*, (c) *syn,anti*

chelating carboxy-oxygen makes a short contact to the copper atom (glycolate and lactate), this contact is neither *syn* nor *anti*. The oxygen forms a strong hydrogen-bond in the *syn*-positions, suggesting that there is *sp²* hybridisation of the oxygen orbitals. However, the weak copper oxygen interaction is directed out of the plane of the carboxy-group and above the oxygen lone-pair rather than in the direction of the lone-pair. The interactions may be purely electrostatic and non-directional with respect to the oxygen, or perhaps utilise the oxygen π -electrons. The latter alternative seems less likely, since the interaction becomes weaker as its direction moves nearer to the perpendicular to the carboxy-group of the oxygen (lactate). Strong copper interactions in the *anti,anti*-configuration at a carboxy-

group, as in the formate tetrahydrate,²⁴ lead to super-exchange effects. However, these effects are no longer observed as the direction of interaction moves out of the plane and interaction becomes weaker.

The methoxy- and phenoxy-acetate crystal structures differ from those of the 2-hydroxycarboxylates. No hydrogen-bond can be formed between O(2) and O(3), and the molecules are now linked into three-dimensional networks by hydrogen-bonds between the water molecules and the carboxylate groups. In the methoxy-acetate, there are two hydrogen-bonds, one *syn* and one *anti*, to aquo-ligands of two neighbouring complexes. In the phenoxyacetate, hydrogen-bonding is to the chelating as well as to the non-chelating carboxy-oxygens, and bifurcated hydrogen bonds may be formed [Figure 1(b)].

EXPERIMENTAL

Preparations.—Crystals were obtained by slow growth from aqueous solutions of pH < 5 in order to repress hydrolysis. All solutions were prepared in one of two ways.

(A) Basic copper carbonate was dissolved in carboxylic acid and any excess of carbonate filtered off.

(B) Carboxylic acid-sodium carboxylate buffers were added to dilute solutions of copper perchlorate containing a small amount of perchloric acid. These solutions, which had been prepared in connection with our related thermodynamic studies,^{2,20} were all made up in a 3M-sodium perchlorate ionic medium. Crystals were analysed for copper

²³ R. L. Martin and H. Waterman, *J. Chem. Soc.*, 1959, 1359.

²⁴ R. Kiriyama, H. Ibamoto, and K. Matsuo, *Acta Cryst.*, 1954, 7, 482.

TABLE 6

Observed structure amplitudes and calculated structure factors for bis(glycollato)copper(II)

h	k	l	$ F_o $	F_c	h	k	l	$ F_o $	F_c	h	k	l	$ F_o $	F_c	h	k	l	$ F_o $	F_c		
0	0	2	20	16	1	1	1	235	220	2	3	4	17	-15	4	-2	4	136	137		
0	0	4	24	50	1	1	2	28	46	2	3	5	24	-24	4	-2	5	87	78		
0	0	6	195	185	11	1	3	308	306	2	3	6	130	110	4	-1	0	144	161		
0	0	8	115	110	1	1	4	58	39	2	3	0	125	97	4	-1	1	27	-30		
0	0	10	55	61	1	1	5	74	73	2	3	1	20	-18	4	-1	2	259	266		
0	1	2	191	173	1	1	6	34	-35	2	3	2	103	82	4	-1	3	47	40		
0	1	3	101	-84	1	1	7	79	65	2	3	3	24	-19	4	-1	4	42	59		
0	1	4	198	197	1	1	8	20	-19	2	3	4	78	57	4	-1	5	52	37		
0	1	5	10	9	1	1	9	111	108	2	3	5	26	18:	4	-1	6	104	91		
0	1	6	103	104	1	1	10	20	24	2	3	6	18	-12	4	0	0	192	206		
0	1	7	22	24	1	2	1	231	195	2	3	7	20	57	4	0	1	127	105		
0	1	8	94	90	1	2	2	86	-70	2	3	8	59	-14	4	0	2	238	205		
0	1	9	10	11	1	2	3	215	200	2	3	9	90	74	4	0	3	22	13		
0	1	10	70	77	1	2	4	26	-23	2	3	10	50	47	4	0	4	98	109		
0	2	0	71	-29	1	2	5	118	110	2	3	11	23	-22	4	0	5	73	-71		
0	2	1	229	-184	1	2	6	64	54	2	3	12	58	57	4	0	6	63	55		
0	2	2	309	248	1	2	7	30	18	2	3	13	81	71	4	1	0	160	166		
0	2	3	79	-59	1	2	8	10	-13	2	3	14	24	29	4	1	1	61	54		
0	2	4	267	265	1	2	9	109	121	2	3	15	37	32	4	1	2	137	132		
0	2	6	67	56	1	3	1	228	183	2	3	16	135	128	4	1	3	88	-89		
0	2	7	66	61	1	3	2	21	14	2	3	17	85	71	4	1	4	145	135		
0	2	8	16	18	1	3	3	3	68	50	2	3	18	63	49	4	1	5	85	76	
0	2	9	13	-10	1	3	3	4	102	-89	2	3	19	20	-15	4	2	0	144	151	
0	2	10	60	60	1	3	3	5	139	122	2	3	20	43	33	4	2	2	92	-7	
0	3	0	122	103	1	3	6	62	56	2	3	21	139	125	4	2	2	3	97		
0	3	1	90	-63	1	3	7	8	11	10	2	3	22	68	-63	4	2	2	4	81	
0	3	2	154	131	1	4	1	111	83	2	3	23	100	91	4	2	2	5	69		
0	3	3	56	47	1	4	2	2	44	32	2	3	24	43	55	4	2	2	6	129	
0	3	4	171	152	1	4	3	3	94	73	2	3	25	108	105	4	2	2	7	79	
0	3	5	38	-29	1	4	4	4	51	-39	2	3	26	34	31	4	2	2	8	88	
0	3	6	115	109	1	4	5	5	117	101	2	3	27	115	114	4	2	2	9	91	
0	3	7	13	11	1	4	6	6	11	6	2	3	28	175	177	4	2	2	10	72	
0	3	8	68	66	1	4	7	7	79	63	2	3	29	277	273	4	2	2	11	14	
0	3	9	163	138	1	4	8	9	9	13	2	3	30	104	94	4	2	2	12	-30	
0	3	10	70	52	1	5	1	5	58	41	2	3	31	144	96	4	2	2	13	40	
0	3	11	28	68	1	5	2	35	-28	2	3	32	44	38	4	2	2	14	31		
0	3	12	85	73	1	5	3	9	70	60	2	3	33	138	146	4	2	2	15	35	
0	3	13	108	95	1	5	4	4	46	40	2	3	34	22	13	4	2	2	16	27	
0	3	14	85	-73	1	5	5	5	86	74	2	3	35	130	120	4	2	2	17	34	
0	3	15	108	90	1	5	6	28	-19	2	3	36	184	-170	4	2	2	18	68		
0	3	16	80	55	1	5	7	48	45	2	3	37	61	66	4	2	2	19	62		
0	3	17	10	7	1	5	8	9	13	2	3	38	277	273	4	2	2	20	44		
0	3	18	28	23	1	5	9	58	-29	2	3	39	104	94	4	2	2	21	45		
0	3	19	85	73	1	5	10	32	-27	2	3	40	101	94	4	2	2	22	35		
0	3	20	20	18	1	5	11	74	65	2	3	41	178	168	4	2	2	23	39		
0	3	21	66	49	1	5	12	50	42	2	3	42	231	234	4	2	2	24	40		
0	3	22	57	47	1	5	13	69	55	2	3	43	125	128	4	2	2	25	44		
0	3	23	1	34	1	5	14	50	42	2	3	44	34	36	4	2	2	26	44		
0	3	24	110	88	1	5	15	90	77	2	3	45	146	146	4	2	2	27	44		
0	3	25	12	8	1	5	16	11	15	2	3	46	178	177	4	2	2	28	44		
0	3	26	100	72	1	5	17	204	197	2	3	47	221	217	4	2	2	29	44		
0	3	27	4	29	19	1	5	18	127	112	2	3	48	223	213	4	2	2	30	45	
0	3	28	5	58	36	1	5	19	32	-27	2	3	49	104	94	4	2	2	31	45	
0	3	29	2	116	-98	1	5	20	106	104	2	3	50	170	166	4	2	2	32	46	
0	3	30	3	148	118	1	5	21	37	-19	2	3	51	22	22	4	2	2	33	46	
0	3	31	4	64	44	1	5	22	17	140	2	3	52	115	114	4	2	2	34	47	
0	3	32	5	57	47	1	5	23	11	-118	2	3	53	85	-81	4	2	2	35	47	
0	3	33	6	1	34	1	5	24	48	40	2	3	54	34	26	4	2	2	36	47	
0	3	34	7	259	70	1	5	25	64	-42	2	3	55	125	128	4	2	2	37	48	
0	3	35	1	78	65	1	5	26	14	-12	2	3	56	54	54	4	2	2	38	48	
0	3	36	2	46	41	1	5	27	123	105	2	3	57	213	216	4	2	2	39	49	
0	3	37	3	23	16	1	5	28	12	-9	2	3	58	253	269	4	2	2	40	50	
0	3	38	4	110	88	1	5	29	90	77	2	3	59	162	170	4	2	2	41	51	
0	3	39	5	12	8	1	5	30	11	55	2	3	60	188	180	4	2	2	42	52	
0	3	40	6	20	18	1	5	31	258	275	2	3	61	243	238	4	2	2	43	53	
0	3	41	7	164	141	1	5	32	64	-42	2	3	62	18	15	4	2	2	44	54	
0	3	42	8	25	-22	1	5	33	25	35	2	3	63	23	21	4	2	2	45	55	
0	3	43	9	308	185	1	5	34	127	148	2	3	64	113	97	4	2	2	46	56	
0	3	44	10	249	-22	1	5	35	34	35	2	3	65	30	-29	4	2	2	47	57	
0	3	45	11	6	22	-28	1	5	36	202	225	2	3	66	162	170	4	2	2	48	58
0	3	46	12	7	103	99	1	5	37	197	199	2	3	67	187	181	4	2	2	49	59
0	3	47	13	71	69	1	5	38	280	287	2	3	68	186	180	4	2	2	50	60	
0	3	48	14	2	97	71	2	3	1	366	-377	2	3	69	241	-39	4	2	2	51	61
0	3	49	15	3	258	280	2	3	2	178	180	2	3	70	104	104	4	2	2	52	62
0	3	50	16	4	92	-83	2	3	3	140	141	2	3	71	17	-15	4	2	2	53	63
0	3	51	17	5	176	198	2	3	4	58	72	2	3	72	81	-29	4	2	2	54	64
0	3	52	18	6	60	49	2	3	5	165	205	2	3	73	40	-47	4				

TABLE 6 (Continued)

<i>h</i>	<i>k</i>	<i>l</i>	F_o	F_c	<i>h</i>	<i>k</i>	<i>l</i>	F_o	F_c	<i>h</i>	<i>k</i>	<i>l</i>	F_o	F_c	<i>h</i>	<i>k</i>	<i>l</i>	F_o	F_c
8 -3	5	10	-12		8 -1	2	61	78		8 0	4	45	54		9 -4	2	15	26	
8 -3	6	30	29		8 -1	3	38	40		8 0	5	9	5		9 -4	3	43	71	
0 -2	0	96	126		8 -1	4	14	16		8 1	0	40	49		9 -3	1	26	41	
8 -2	1	18	25		8 -1	5	15	-26		8 1	1	43	-52		9 -3	3	48	74	
8 -2	2	34	49		8 -1	6	50	62		8 1	2	48	57		9 -3	4	20	-30	
8 -2	3	19	20		8 0	0	43	56		8 1	3	9	-13		9 -3	5	21	39	
8 -2	4	36	51		8 0	1	62	-66		8 1	4	47	51		9 -2	1	50	69	
8 -2	6	58	72		8 0	2	66	85		9 -4	1	27	45		9 -2	3	13	20	
8 -1	0	86	121		8 0	3	23	19							10 -1	0	16	-20	

by complexometric titration using EDTA with murexide as indicator.²⁵

Bis(glycollato)copper(II). Method B (Found: C, 22.4; H, 2.75; Cu, 29.3. $C_4H_6CuO_6$ requires C, 22.5; H, 2.85; Cu, 29.7%).

Aquobis-(DL-lactato)copper(II) hemihydrate. Method A, the blue surfaces of these crystals became coated with a white sheen on standing, owing to efflorescence. Analyses of various batches indicated a variable loss of water. An anhydrous sample was obtained by dehydration in a vacuum-desiccator over phosphorus pentoxide (Found: C, 29.9; H, 4.15; Cu, 26.3. $C_6H_{10}CuO_6$ requires C, 29.8; H, 4.15; Cu, 26.3%).

Diaquobis-(2-hydroxy-2-methylpropionato)copper(II). Method A (Found: C, 31.3; H, 6.2; Cu, 20.6. $C_8H_{18}CuO_8$ requires C, 31.4; H, 5.95; Cu, 20.8%).

Diaquobis(methoxyacetato)copper(II). Method B (Found: C, 25.8; H, 5.1; Cu, 23.0. $C_6H_{14}CuO_8$ requires C, 25.9; H, 5.05; Cu, 22.9%).

Diaquobis(phenoxyacetato)copper(II). Method B (Found: C, 47.9; H, 4.5; Cu, 15.6. $C_{16}H_{18}CuO_8$ requires C, 47.8; H, 4.5; Cu, 15.8%). In the absence of sodium perchlorate, clusters of green needle-shaped anhydrous bis(phenoxyacetato)copper(II) crystals were obtained. These are the subject of a continuing investigation.

X-Ray Photography.—Crystals of all five complexes were monoclinic needles. The unit cell dimensions were obtained from calibrated zero-layer Weissenberg photographs taken about the needle axis and a second principal axis of the crystal so that one photograph was taken about the unique axis of the crystal. The X-ray intensities were estimated visually from sets of multiple-film Weissenberg photographs about the needle axis of the crystals. The results from individual layers were placed on a common scale either by using the exposure time to X-rays from a stabilised source and standardised development conditions [diaquobis(phenoxyacetato)copper(II), diaquobis-(2-hydroxy-2-methylpropionato)copper(II), and aquobis-(lactato)copper(II) hemihydrate] or by using intensities collected about a second axis [bis(glycollato)copper(II) and diaquobis(methoxyacetato)copper(II)] with layer scale factor calculated by the method of Hamilton, Rollett, and Sparks.²⁶ All intensities were corrected for Lorentz and polarisation factors but not for extinction or absorption.

Calculations.—Structure factors, cycles of least-squares refinement, Fourier syntheses, and interatomic distances were computed on an English Electric KDF9 computer using J. S. Rollett's 'Novtape' programme²⁷ as modified and extended by J. Hodder. Atomic scattering factors for carbon, oxygen, hydrogen, and copper were taken from International Tables.²⁸ An anomalous-dispersion correction was applied to the scattering curve for copper.

²⁵ G. Schwarzenbach, 'Complexometric Titrations,' Methuen, London, 1957.

²⁶ W. C. Hamilton, J. S. Rollett, and R. A. Sparks, *Acta Cryst.*, 1965, **18**, 129.

Details of individual structure analyses follow. In each case, the authors mainly responsible for the experimental work are given in parentheses.

Diaquobis(phenoxyacetato)copper(II) (R. A. Armstrong).— $C_{16}H_{18}CuO_8$, $M = 401.85$, Monoclinic prismatic, $a = 16.26 \pm 0.04$, $b = 7.28 \pm 0.03$, $c = 23.99 \pm 0.06$ Å, $\gamma = 110.4 \pm 0.3^\circ$. $U = 2653$ Å³, $D_m = 1.531 \pm 0.01$ (by flotation), $Z = 6$, $D_c = 1.510$. Space group $P2_1/n$ (C_{2h}^5 , No. 14), general positions $\pm(x, y, z; \frac{1}{2} + x, \frac{1}{2} + y, \frac{1}{2} - z)$. Cu- K_α radiation, $\mu = 15.35$ cm⁻¹, crystal size $0.12 \times 0.15 \times 1.0$ mm. Optically biaxial.

To have six bis(phenoxyacetato)copper(II) molecules in a unit cell in space group $P2_1/n$ it is necessary that they should be of at least two crystallographically dissimilar types. If two molecules lie with their copper atoms at the inversion centres at 0,0,0 and $\frac{1}{2},\frac{1}{2},\frac{1}{2}$, then the other four may have copper atoms in general positions and the molecules need not be centrosymmetric. The oscillation photographs about the a -axis shows layers 0, 3, 6, and 9 much more intense than the rest, suggesting molecules in similar orientations separated from each other by $a/3$.

A three-dimensional Patterson function sharpened to 'point atoms at rest' was computed from 2383 independent reflections. The most prominent vectors could be explained if two copper atoms were at the special positions 0,0,0 and $\frac{1}{2},\frac{1}{2},\frac{1}{2}$, and a further four at the general positions 0.33, 0.08, 0, and symmetry-related positions. A three-dimensional electron-density distribution was computed from the phases given by the proposed copper positions. Prominent independent maxima in this distribution correspond to three phenoxyacetate residues and three water molecules. The trial parameters were refined by two least-squares cycles assuming isotropic thermal motion for all atoms and unit weights for all observations. The weighting scheme

$$w = \{1 + [(|F_o| - a)/b]^2\}^{-1} \quad (2)$$

was then adopted with $a = b = 100$ [on the scale of Table 4]. After two further refinement cycles $R = 0.131$. At this stage it was possible to locate hydrogen atoms from a three-dimensional difference synthesis calculated so that the two-dimensional sections were parallel to the plane of the phenoxyacetate molecules of the complexes. The hydrogen atoms were placed within the maxima on the difference synthesis at accepted bonded distances from the atoms to which they were attached. Three very strong reflections showing very considerable extinction effects were removed, and after two further least-squares cycles convergence was reached with $R = 0.113$. In all least-squares refinement cycles, a block diagonal approximation to the normal matrix was reduced to three blocks 39×39 , 36×36 , and 75×75 , such that interactions of the temperature and overall scale parameters, and the space parameters, and of

²⁷ J. S. Rollett, unpublished work.

²⁸ International Tables for X-ray Crystallography, vol. 3 Kynoch Press, Birmingham, 1962, p. 220.

TABLE 7

Observed structure amplitudes and calculated structure factors for aquobis-(DL-lactato)copper(II) hemihydrate

h	k	l	$ F_o $	F_o	α	h	k	l	$ F_o $	F_o	α	h	k	l	$ F_o $	F_o	α	h	k	l	$ F_o $	F_o	α	
0	2	4	145	213	171	1	6	0	45	114	-180	2	5	3	125	146	-179	3	7	1	220	219	4	
0	3	3	194	246	175	1	6	2	444	420	-173	2	6	2	149	148	-159	7	3	81	88	-72	0	
0	4	0	75	49	180	1	6	4	169	201	-180	4	48	67	149	149	-149	0	64	41	41	-87	0	
0	4	2	275	194	129	1	7	1	153	147	-135	7	1	393	375	-7	8	2	24	26	-87	-10		
0	4	4	94	102	-50	1	7	3	42	50	158	7	3	158	168	12	8	4	47	37	-10	-10		
0	5	1	297	310	24	1	8	0	490	357	0	8	0	296	256	0	9	1	303	305	177	177		
0	5	3	298	302	-21	1	8	2	473	134	1	8	2	266	272	4	9	3	117	127	142	142		
0	6	0	137	137	0	1	8	4	130	207	-5	8	4	91	92	-7	10	2	59	59	47	47		
0	6	2	300	232	-115	1	9	1	297	275	28	9	1	186	203	173	10	4	17	27	155	155		
0	6	4	77	90	145	1	9	3	150	150	35	10	0	296	267	180	11	1	215	214	1	1		
0	7	1	239	245	175	1	10	0	314	268	180	10	0	296	262	179	11	3	119	119	-12	-12		
0	7	3	322	320	157	1	10	2	121	192	167	10	4	59	77	-162	12	0	79	67	0	0		
0	8	0	343	270	180	1	10	4	97	105	168	11	1	43	35	170	12	2	33	37	-23	-23		
0	8	2	105	72	124	1	11	1	383	330	-152	11	3	86	84	-177	12	4	48	54	-11	-11		
0	9	4	71	75	-128	1	11	3	156	159	-153	12	0	237	192	0	13	1	33	40	-136	-136		
0	9	9	1	273	295	-17	1	12	0	68	58	130	12	2	259	225	-1	14	0	77	83	-180	-180	
0	9	3	300	283	-6	1	12	2	171	121	-163	12	4	57	68	30	14	2	92	76	-174	-174		
0	10	0	102	78	0	1	12	4	47	54	-163	13	1	86	85	-39	14	4	79	68	160	160		
0	10	2	87	95	32	1	13	1	203	195	11	13	3	54	55	-11	15	3	35	38	25	25		
0	10	4	67	82	22	1	13	3	67	53	-31	14	0	29	38	180	16	0	109	92	0	0		
0	11	1	292	214	157	1	14	0	107	100	0	14	2	155	129	177	16	2	86	84	12	12		
0	11	3	172	170	-170	1	14	2	108	113	-7	14	4	24	29	-160	17	3	16	15	138	138		
0	12	2	141	132	-151	1	14	5	50	59	-3	15	1	97	118	144	18	0	139	104	-180	-180		
0	12	4	47	52	176	1	15	1	154	146	-175	16	3	66	62	-180	19	1	53	51	-173	-173		
0	13	1	79	70	-56	1	15	3	62	73	114	16	2	23	28	180	14	-25	-1	30	44	104	104	
0	13	3	195	165	0	1	16	2	49	53	127	16	2	55	63	-5	4	-24	2	15	32	159	159	
0	14	2	230	218	7	1	16	4	24	37	157	17	1	88	100	-31	4	-24	0	27	29	-180	-180	
0	14	4	60	74	-5	1	17	1	94	102	9	18	2	31	28	165	4	-23	-1	31	47	-80	-80	
0	15	3	52	50	-40	1	17	3	40	51	-47	19	1	60	59	154	4	-22	-2	35	49	11	11	
0	15	0	195	183	180	1	18	0	87	70	0	20	2	44	50	-157	4	-22	0	90	110	0	0	
0	15	2	217	196	177	1	18	2	54	47	-28	21	-3	61	69	-21	4	-21	-3	15	21	33	33	
0	15	4	44	50	-176	1	18	4	57	50	-5	20	-1	90	95	-13	4	-20	-4	78	89	163	163	
0	17	3	32	39	168	1	19	1	28	30	-133	20	-2	53	61	173	4	-19	-3	96	121	-173	-173	
0	18	2	125	129	-6	1	19	3	16	14	175	21	-3	63	63	-178	4	-18	-2	103	108	20	20	
0	18	4	21	33	34	1	20	0	35	48	180	23	-1	35	35	-21	4	-17	-3	151	183	6	6	
0	19	1	94	91	-8	1	20	2	48	40	-173	24	-2	21	32	7	4	-20	-2	73	97	-161	-161	
0	19	3	71	65	7	1	22	0	28	34	0	21	-3	61	69	-21	4	-19	-1	73	80	167	167	
0	24	2	52	55	174	2	24	0	46	56	0	20	-2	53	61	-180	4	-18	-2	103	108	20	20	
0	21	1	78	78	164	2	24	-1	45	53	-179	19	-3	68	102	169	4	-14	-4	59	63	163	163	
0	21	3	43	52	-176	2	24	-2	78	74	179	18	-2	44	64	-168	4	-17	-1	203	210	-8	-8	
0	22	0	48	64	-180	2	22	0	27	49	180	19	-3	68	102	169	4	-16	-2	61	61	58	-148	-148
1	-23	1	26	31	-42	2	21	-3	56	28	175	19	-1	138	127	155	4	-17	-3	151	183	6	6	
1	-22	2	79	79	-166	2	21	-1	50	57	-23	18	-2	87	90	167	4	-16	-2	61	61	58	-148	-148
1	-22	0	62	73	-180	2	20	-4	31	42	-2	19	-2	87	90	167	4	-16	-2	186	200	-180	-180	
1	-21	3	27	33	-81	2	20	-2	110	122	-16	17	-3	81	98	20	4	-19	-1	73	80	167	167	
1	-21	1	42	52	158	2	20	0	72	68	0	17	-1	97	111	-25	4	-18	-4	48	72	-34	-34	
1	-20	4	29	43	-7	2	19	-3	24	32	-102	16	-4	123	152	-4	4	-15	-3	119	123	-164	-164	
1	-20	2	90	105	-1	2	19	-6	56	57	90	15	-3	226	213	6	4	-14	-4	59	63	163	163	
1	-19	0	62	54	0	2	18	-4	45	59	-179	15	-1	138	127	155	4	-14	-2	33	21	104	104	
1	-19	3	32	34	4	2	18	-2	105	102	157	18	-2	87	90	-168	4	-14	-3	353	353	-180	-180	
1	-19	1	113	125	-10	2	18	0	71	52	-180	15	-1	68	67	-168	4	-13	-3	162	159	-7	-7	
1	-18	4	14	24	151	2	17	-3	64	73	164	14	-4	115	132	173	4	-13	-1	288	264	12	12	
1	-18	2	81	70	161	2	17	-1	98	108	-134	14	-2	187	170	-145	4	-12	-4	91	87	-12	-12	
1	-17	3	95	110	161	2	16	-4	42	58	-2	11	-1	161	142	-10	4	-9	-3	129	131	-75	-75	
1	-17	1	217	221	172	2	16	-5	32	45	7	10	-2	120	148	170	4	-8	-4	199	206	11	11	
1	-16	1	25	32	-138	2	16	0	86	90	0	9	-3	54	53	-45	4	-8	-2	332	368	-13	-13	
1	-16	2	33	34	-153	2	15	-3	119	127	13	12	-4	91	90	-9	4	-8	0	432	400	0	0	
1	-16	0	70	58	-180	2	15	-1	119	99	42	12	-2	101	111	70	4	-7	-3	143	110	67	67	
1	-15	2	70	74	-12	2	14	-4	93	105	164	14	-2	233	240	0	4	-5	-3	237	123	-171	-171	
1	-15	1	141	141	-13	2	14	-4	151	146	-150	11	-2	211	201	87	4	-5	-1	182	157	137	137	
1	-14	4	78	82	15	2	13	-5	322	307	-165	11	-2	202	185	-65	4	-4	-1	186	100	23	23	
1	-14	0	147	147	147	2	13	-5	201	205	-165	10	-3	282	282	8	4	-6	-2	340	341	178	178	
1	-14	3	94	97	-180	2	12	-6	76	66	-108	7	-1	500	465	13	4	-6	-2	260	234	-55	-55	
1	-14	1	309	252	-25	2	10	-4	80	83	48	8	-2	202										

TABLE 7 (Continued)

<i>h</i>	<i>k</i>	<i>l</i>	<i>F</i> _o	<i>F</i> _c	α	<i>h</i>	<i>k</i>	<i>l</i>	<i>F</i> _o	<i>F</i> _c	α	<i>h</i>	<i>k</i>	<i>l</i>	<i>F</i> _o	<i>F</i> _c	α	<i>h</i>	<i>k</i>	<i>l</i>	<i>F</i> _o	<i>F</i> _c	α		
4	7	3	25	39	-80	5	8	0	66	61	-180	7	-24	-2	24	20	31	8	-16	-4	41	65	-17		
4	8	0	232	211	-180	5	8	2	74	56	-176	7	-24	0	28	54	0	8	-16	-2	113	131	7		
4	8	2	320	310	178	5	9	4	71	73	132	7	-23	-3	13	29	164	8	-16	0	23	32	0		
4	8	4	103	107	175	5	9	1	65	72	50	7	-22	-2	35	50	175	8	-15	-3	85	87	28		
4	9	1	149	156	129	5	9	3	103	103	0	7	-22	0	64	96	-180	8	-14	-1	65	80	-25		
4	9	3	83	70	-150	5	10	0	176	142	0	7	-21	-3	17	35	-30	8	-14	-2	138	163	-160		
4	10	0	115	105	0	5	10	2	119	110	20	7	-21	-1	30	26	-51	8	-14	-3	52	48	-136		
4	10	2	198	182	-2	5	10	4	89	93	-13	7	-20	-4	62	80	19	8	-13	-1	95	96	-171		
4	10	4	91	91	-1	5	11	3	53	45	146	7	-20	-2	84	99	-25	8	-12	-2	201	224	15		
4	11	1	139	131	-36	5	11	0	217	177	-180	7	-20	0	116	159	0	8	-12	0	125	163	0		
4	11	3	125	120	38	5	12	2	136	100	-161	7	-19	-3	40	49	-7	8	-11	-3	42	37	-30		
4	12	0	47	23	0	5	12	4	106	96	-176	7	-19	-1	56	65	-1	8	-11	-1	163	180	8		
4	12	2	23	21	83	5	13	1	48	41	-153	7	-18	-4	83	107	-173	8	-10	-4	63	72	170		
4	12	4	28	35	-173	5	13	3	33	27	-127	7	-18	-2	112	139	144	8	-10	-2	133	136	-174		
4	13	1	94	85	157	5	14	0	111	92	0	7	-18	0	209	223	-180	8	-10	-4	63	72	170		
4	13	3	107	109	-155	5	14	2	19	10	2	7	-17	-3	103	116	169	8	-10	-2	123	124	-180		
4	14	0	21	29	180	5	15	1	64	47	34	7	-16	-4	50	64	19	8	-9	-3	139	127	142		
4	14	2	63	70	-158	5	16	0	82	66	-180	7	-16	-2	57	74	-59	8	-8	-4	35	35	-30		
4	14	4	20	20	63	5	16	2	67	94	-180	7	-16	0	148	159	0	8	-7	-3	72	74	-174		
4	15	1	72	58	-2	5	16	-25	-1	28	28	157	7	-15	-3	140	149	-2	8	-7	-8	0	55	-180	
4	15	3	98	77	0	5	16	-24	0	45	56	19	7	-15	-1	183	186	-15	8	-6	-4	53	46	22	
4	16	0	25	12	0	5	16	-24	0	48	67	0	7	-14	-2	108	109	4	8	-6	-2	235	224	1	
4	16	2	73	56	20	5	16	-23	-3	41	60	24	7	-14	-2	106	109	4	8	-5	-1	107	86	-147	
4	17	1	39	41	-165	5	16	-23	1	38	62	-1	7	-14	-2	108	109	4	8	-4	-2	232	240	177	
4	18	0	21	17	180	5	16	-22	-2	39	60	-160	7	-14	0	23	20	0	8	-4	-2	166	199	-180	
5	-25	-1	47	54	155	5	16	-21	-3	28	58	-175	7	-13	-3	133	154	-163	8	-5	-1	165	156	133	
5	-24	-2	15	27	171	5	16	-21	1	75	94	-169	7	-13	-1	203	201	-179	8	-4	-2	120	110	-27	
5	-23	-3	58	71	20	5	16	-20	-4	15	53	16	7	-12	-4	99	86	160	8	-4	-2	232	240	177	
5	-23	-1	59	75	-6	5	16	-20	2	85	96	3	7	-12	-2	98	105	170	8	-4	-1	138	130	-70	
5	-22	-2	36	48	17	5	16	-19	-3	71	103	-26	7	-12	0	30	39	-180	8	-3	-1	11	11	0	
5	-22	0	61	86	0	5	16	-19	-2	131	137	10	7	-11	-3	83	85	44	8	-2	0	73	76	0	
5	-21	-3	42	74	-143	5	16	-18	-2	38	6	0	7	-10	-4	147	158	30	8	0	0	108	89	180	
5	-21	-1	82	92	-167	5	16	-18	0	106	134	0	7	-10	-2	113	117	-7	8	0	0	27	21	117	
5	-21	-2	66	67	-153	5	16	-17	-1	122	114	-168	7	-10	0	102	91	0	8	0	0	115	106	-178	
5	-21	0	122	150	-180	5	16	-16	-4	73	92	-180	7	-9	-3	63	72	-165	8	1	1	47	42	157	
5	-19	-3	30	46	43	5	16	-16	-2	160	179	176	7	-9	-1	134	148	-145	8	1	3	83	69	-169	
5	-19	-1	56	64	31	5	16	-16	0	261	273	-180	7	-8	-4	100	104	-176	8	2	0	223	218	0	
5	-18	-2	77	71	23	5	16	-15	-3	122	114	-28	7	-8	-2	263	275	-178	8	2	2	2	2	-25	
5	-18	0	134	151	0	5	16	-14	-4	92	98	19	7	-7	-3	40	41	-5	8	3	1	102	85	15	
5	-17	-3	19	14	37	5	16	-14	-2	152	160	0	7	-6	-4	86	70	31	8	3	3	112	98	-10	
5	-17	-1	31	46	-8	5	16	-13	5	56	59	158	7	-6	-2	245	253	6	8	4	2	41	31	114	
5	-16	-4	52	68	-165	5	16	-13	-1	121	120	-146	7	-6	0	210	231	0	8	5	1	129	107	-167	
5	-16	-2	92	90	167	5	16	-12	-4	128	131	-172	7	-5	-3	31	31	-23	8	4	4	41	26	166	
5	-16	0	150	173	-180	5	16	-12	-2	188	191	-179	7	-4	-4	40	31	-151	8	5	3	118	85	170	
5	-15	-3	105	97	-154	5	16	-12	0	73	78	-180	7	-4	-2	169	163	-177	8	5	3	121	98	5	
5	-15	-1	177	177	-158	5	16	-11	-3	56	60	-163	7	-4	0	103	106	-180	8	7	3	63	51	-4	
5	-14	-4	93	112	36	5	16	-11	-1	107	98	73	7	-3	-3	27	34	-169	8	8	2	34	27	-115	
5	-14	-2	141	148	-28	5	16	-10	-4	151	147	-13	7	-3	-1	146	147	-144	8	9	1	117	71	178	
5	-14	0	166	212	0	5	16	-10	-2	232	249	-1	7	-2	-2	151	166	-19	8	9	2	34	27	-115	
5	-13	-3	115	113	7	5	16	-10	0	128	132	0	7	-2	-2	110	132	0	8	-23	-3	25	44	-164	
5	-13	-1	157	170	12	5	16	-9	-3	147	141	24	7	-1	-1	179	180	15	8	-22	-2	38	90	0	
5	-12	-4	127	133	-154	5	16	-8	-4	146	148	18	7	0	0	57	58	-144	8	-24	-2	34	45	132	
5	-12	-2	213	218	-142	5	16	-8	-2	148	139	177	7	1	1	227	200	-177	8	-23	-3	33	48	143	
5	-11	-3	201	179	158	5	16	-7	-3	201	194	-162	7	1	3	233	206	-179	8	-22	-2	24	36	-30	
5	-10	-4	75	73	21	5	16	-6	-4	25	25	8	7	2	0	61	59	-180	8	-21	-3	40	59	18	
5	-10	-2	131	151	-67	5	16	-6	-2	151	161	30	7	2	2	6	65	-151	8	-20	-2	21	28	-128	
5	-9	-3	534	328	-19	5	16	-6	0	67	76	0	7	2	4	61	65	-142	8	-19	-3	63	61	30	
5	-8	-1	373	384	-11	5	16	-5	-3	157	149	5	7	3	1	117	110	11	8	-18	-4	46	54	-157	
5	-8	-2	337	327	7	5	16	-4	-2	46	38	-161	7	3	4	156	156	-6	8	-17	-2	70	80	-6	
5	-8	0	111	126	0	5	16	-4	-2	119	106	-153	7	4	4	2	76	75	54	8	-16	0	79	91	0
5	-7	-3	215	210	178	5	16	-4	0	91	93	-180	7	7	1	3	31	34	-103	8	-15	-3	67	74	-173
5	-7	-1	305	300	-178	5	16	-3	-3	71	50	113	7	7	3	18	12	28	8	-15</td					

TABLE 7 (Continued)

<i>h</i>	<i>k</i>	<i>l</i>	F_o	F_c	α	<i>h</i>	<i>k</i>	<i>l</i>	F_o	F_c	α	<i>h</i>	<i>k</i>	<i>l</i>	F_o	F_c	α	<i>h</i>	<i>k</i>	<i>l</i>	F_o	F_c	α
9	1	1	94	85	-33	10	-15	-3	67	72	143	10	5	1	54	31	8	11	-7	-1	129	96	-154
9	1	3	86	68	21	10	-15	-1	158	170	-176	11	-23	-1	47	50	18	11	-5	-1	110	92	34
9	2	0	260	205	0	10	-14	-4	32	24	-21	11	-22	-2	14	21	-10	11	-3	-1	70	56	-131
9	2	2	129	106	-0	10	-14	-2	40	47	177	11	-22	0	15	28	0	11	0	0	82	46	0
9	2	4	102	64	3	10	-13	-3	68	64	-4	11	-21	-1	72	69	-165	11	0	2	53	43	-2
9	3	1	46	54	93	10	-12	-2	67	56	-6	11	-20	-2	38	45	180	11	2	0	92	46	180
9	4	0	200	149	-180	10	-11	-3	89	79	-153	11	-20	0	31	44	-180	12	-22	0	33	49	-180
9	4	2	117	97	176	10	-10	-1	72	48	-173	11	-19	-3	60	72	18	12	-21	-1	46	65	-170
9	5	1	26	24	-110	10	-10	-2	56	59	168	11	-19	-1	62	82	-6	12	-20	0	40	48	0
9	5	3	14	8	-158	10	-10	0	34	39	-180	11	-18	-2	60	56	-12	12	-19	-1	21	36	-34
9	6	0	86	44	0	10	-9	-3	88	82	28	11	-18	0	44	54	0	12	-18	0	26	30	-180
9	6	2	47	39	2	10	-8	-1	130	111	-50	11	-17	-3	58	51	-166	12	-17	-1	89	70	138
9	7	1	48	33	-18	10	-8	-2	59	65	-2	11	-16	-2	41	50	156	12	-15	-1	89	76	-31
10	-23	-1	30	54	31	10	-8	0	100	110	0	11	-16	0	56	55	180	12	-14	-2	30	22	10
10	-22	-2	59	66	19	10	-7	-1	72	61	140	11	-14	-2	56	68	-12	12	-14	0	56	62	0
10	-21	-3	14	18	107	10	-6	0	125	133	-180	11	-14	0	66	90	0	12	-13	-1	14	38	180
10	-20	-2	84	103	-163	10	-3	-1	31	45	-47	11	-13	-3	75	50	20	12	-12	0	88	87	-180
10	-20	0	25	25	180	10	-1	-1	98	82	165	11	-13	-1	116	39	46	12	-11	-1	68	70	21
10	-18	-4	50	63	-14	10	0	0	81	63	0	11	-12	-2	83	81	-176	12	-9	-1	25	33	-128
10	-18	-2	140	162	5	10	0	2	34	35	28	11	-12	0	86	105	-180	12	-7	-1	104	80	170
10	-18	0	107	124	0	10	1	1	142	111	2	11	-11	-3	103	87	-175	12	-5	-1	129	84	-3
10	-17	-3	28	34	-69	10	1	3	70	49	-4	11	-11	-1	30	29	-174	12	-3	-1	97	62	170
10	-17	-1	75	97	21	10	2	2	30	24	-102	11	-10	-2	90	88	15	13	-15	-1	75	40	12
10	-16	-4	69	67	165	10	3	1	145	105	-180	11	-10	0	97	89	0	13	-13	-1	115	67	-155
10	-16	-2	76	88	-179	10	4	0	60	35	-180	11	-9	-1	67	57	-0	13	-11	-1	100	58	37
10	-16	0	89	87	180	10	4	2	46	32	153	11	-8	0	44	43	180	13	-9	-1	59	37	-153
																							33

the space parameters of the centrosymmetric molecule and the space parameters of the molecule in the general position, were ignored. Table 4 gives the observed structure amplitudes and the structures factors calculated from the atomic parameters in Table 3(a).

Diaquobis(methoxyacetato)copper(II) (J. G. Forrest).— $C_6H_{14}CuO_8$, $M = 277\cdot6$, Monoclinic prismatic, $a = 6\cdot92 \pm 0\cdot02$, $b = 7\cdot26 \pm 0\cdot02$, $c = 10\cdot10 \pm 0\cdot03 \text{ \AA}$, $\gamma = 96\cdot7 \pm 0\cdot3^\circ$, $U = 505\cdot2 \text{ \AA}^3$, $D_m = 1\cdot820 \pm 0\cdot005$ (by flotation), $Z = 2$, $D_c = 1\cdot819$. Space group $P2_1/n$ (C_{2h}^5 , No. 14), general position $\pm (x, y, z; \frac{1}{2} + x, \frac{1}{2} + y, \frac{1}{2} - z)$. Cu- K_α radiation, $\mu = 33\cdot8 \text{ cm}^{-1}$. Crystal size $0\cdot4 \times 0\cdot15 \times 0\cdot08 \text{ mm}$. Optically biaxial.

A trial structure was deduced from a Patterson function sharpened to 'point atoms at rest' computed from 722 independent reflections. The trial was refined by F_o synthesis and six cycles of full matrix least-squares refinement, assuming isotropic temperature factors and giving each observation unit weight. The refinement converged with $R = 0\cdot115$. Some adjustment at the layer scale factors was made after the third refinement cycle. At no stage was it possible to locate hydrogen atoms from a difference synthesis. Table 5 gives observed structure amplitude and structure factors calculated from the atomic parameters in Table 3(b).

Bis(glycollato)copper(II) (J. G. Forrest).— $C_4H_6CuO_6$, $M = 213\cdot6$, Monoclinic prismatic, $a = 7\cdot93 \pm 0\cdot02$, $b = 5\cdot08 \pm 0\cdot02$, $c = 8\cdot68 \pm 0\cdot02 \text{ \AA}$, $\gamma = 111\cdot1 \pm 0\cdot3^\circ$, $U = 326\cdot3 \text{ \AA}^3$, $D_m = 2\cdot164 \pm 0\cdot01$ (by flotation), $Z = 2$, $D_c = 2\cdot165$. Space group $P2_1/a$ (C_{2h}^5 , No. 14), general positions $\pm (x, y, z; \frac{1}{2} - x, \frac{1}{2}, \frac{1}{2} + z)$. Cu- K_α radiation, $\mu = 46\cdot0 \text{ cm}^{-1}$. Crystal size $0\cdot4 \times 0\cdot2 \times 0\cdot15 \text{ mm}$. Optically biaxial.

A trial structure was deduced from a three-dimensional Patterson function sharpened to 'point atoms at rest' computed from 602 independent reflections. The trial structure was improved firstly by F_o synthesis and then by four cycles of full matrix least-squares refinement assuming isotropic temperature factors and giving each observation unit weight. The refinement converged with $R = 0\cdot136$. Attempts to locate the hydrogen atoms by difference syntheses failed. Table 6 gives the observed structure amplitudes and structure factors calculated from the atomic parameters in Table 3(c).

Aquabis-(DL-lactato)copper(II) Hemihydrate (J. R. Carruthers and R. A. Armstrong).— $C_6H_{10}CuO_6 \cdot 1\frac{1}{2}H_2O$,

$M = 268\cdot7$, Monoclinic, $a = 10\cdot53 \pm 0\cdot03$, $b = 20\cdot55 \pm 0\cdot05$, $c = 5\cdot66 \pm 0\cdot02$, $\gamma = 119\cdot3^\circ \pm 0\cdot30$, $U = 1\cdot067\cdot9 \text{ \AA}^3$, $D_m = 1\cdot71 \pm 0\cdot03$ (by flotation), $Z = 4$, $D_c = 1\cdot671$. Space group $A2(C_2^3$, No. 5); a piezoelectric test shows no inversion centre. Cu- K_α radiation, $\mu = 32\cdot3 \text{ cm}^{-1}$.

From the three-dimensional Patterson function sharpened to 'point atoms at rest' computed from 981 independent reflections, it was clear that the copper atom was in a general position in space group $A2$. The F_o synthesis computed from the phases given by the tentative copper atom position and the three-dimensional minimum function constructed by superposing the copper-copper vectors had similar sets of maxima. Surprisingly, after careful study of both maps, it was only possible to assign maxima to four atoms [O(1), O(13), O(3), and O(11)] of the copper co-ordination polyhedron and no others. However, the F_o synthesis phased on the five proposed atom positions showed the location of both lactate groups and a water molecule. This trial structure was improved by six cycles of full matrix least-squares refinement, assuming isotropic thermal motion, unit weights for the observations, and applying 0·5 shifts. After the third cycle the layer scale factors were adjusted manually, and the residual was 0·15. A difference map computed at this residual had a maximum of about four electrons at the site O(5). The agreement between observed and calculated densities assuming a monohydrate was not good nor were the analytical results. It was therefore decided to introduce one 'half water molecule' at this position, and to refine its positional and temperature parameters and occupation number along with all other atomic parameters for the structure at $R = 0\cdot15$ by the full matrix least-squares method with the weighting scheme (2) with $a = b = 1000$. The refinement converged after three cycles at $R = 0\cdot121$. At no stage in the refinement was it possible to locate hydrogen atoms from difference syntheses with any certainty. Table 7 lists observed structure amplitudes and calculated structure factors based on the final atomic parameters in Table 3(d).

Diaquobis-(2-hydroxy-2-methylpropionato)copper(II) (R. A. Armstrong).—When a pale blue needle crystal of this complex was first subjected to X -ray photography it showed a diffraction pattern expected for a triclinic twinned crystal with a needle axis spacing of $5\cdot80 \text{ \AA}$. As X -ray photography progressed the diffraction pattern changed during an exposure time of several days. One of several new

TABLE 8

Observed structure amplitudes and calculated stucture factors for diaquobis-(2-hydroxy-2-methylpropionato)copper(II)

<i>h k l</i>	<i>F</i> _o <i>F</i> _c	<i>h k l</i>	<i>F</i> _o <i>F</i> _c	<i>h k l</i>	<i>F</i> _o <i>F</i> _c	<i>h k l</i>	<i>F</i> _o <i>F</i> _c
0 0 13	25 41	7 1 -3	52 36	4 2 9	62 58	5 3 3	-1 46 45
0 0 12	60 61	7 1 -4	35 15	4 2 8	28 35	5 5 5	0 33 42
0 0 11	106 102	7 1 -6	45 42	4 2 7	54 51	5 5 5	2 51 63
0 0 10	79 74	7 1 -7	53 50	4 2 5	70 65	5 5 5	3 50 62
0 0 9	45 60	7 1 -8	46 47	4 2 4	60 59	5 5 5	3 34 49
0 0 8	80 90	7 1 -9	42 43	4 2 3	163 167	5 5 5	7 37 53
0 0 7	93 92	9 1 5	41 40	4 2 2	121 120	5 5 5	8 29 45
0 0 6	103 122	9 1 4	28 30	4 2 1	49 30	5 5 5	7 3 11 14
0 0 5	116 114	9 1 3	18 19	4 2 0	46 27	5 5 5	7 3 8 20
0 0 4	130 143	9 1 2	24 21	4 2 -1	65 52	5 5 5	7 3 5 30
0 0 3	225 218	9 1 1	18 15	4 2 -2	149 143	5 5 5	7 3 4 39
0 0 2	14 13	9 1 0	45 33	4 2 -3	127 122	5 5 5	7 3 3 24
0 0 1	27 38	9 1 -1	24 21	4 2 -4	96 97	5 5 5	7 3 2 33
0 0 0	13	9 1 -2	14 14	4 2 -5	126 139	5 5 5	7 3 1 26
0 0 0	12	9 1 -3	28 15	4 2 -6	55 58	5 5 5	7 3 1 29
0 0 0	11	9 1 -4	14 5	4 2 -7	39 53	5 5 5	7 3 1 31
0 0 0	10	9 1 -5	27 20	4 2 -8	130 135	5 5 5	7 3 1 31
0 0 0	9	9 1 -6	29 28	4 2 -9	83 77	5 5 5	7 3 1 31
0 0 0	8	9 1 -7	25 25	4 2 -10	16 14	5 5 5	7 3 1 31
0 0 0	7	9 1 -8	25 25	4 2 -11	27 37	5 5 5	7 3 1 31
0 0 0	6	9 1 -9	25 25	4 2 -12	26 34	5 5 5	7 3 1 31
0 0 0	5	9 1 -10	25 25	4 2 -13	27 30	5 5 5	7 3 1 31
0 0 0	4	9 1 -11	25 25	4 2 -14	26 28	5 5 5	7 3 1 31
0 0 0	3	9 1 -12	25 25	4 2 -15	21 21	5 5 5	7 3 1 31
0 0 0	2	9 1 -13	25 25	4 2 -16	14 14	5 5 5	7 3 1 31
0 0 0	1	9 1 -14	25 25	4 2 -17	32 31	5 5 5	7 3 1 31
0 0 0	0	9 1 -15	25 25	4 2 -18	62 67	5 5 5	7 3 1 31
0 0 0	102	98 106	1 1 1 4	16 18	6 6 6	6 6 6	0 0 0
0 0 0	3	18 22	1 1 1 2	167 149	6 6 6	6 6 6	0 0 0
0 0 0	2	114 -86	1 1 1 3	123 109	6 6 6	6 6 6	0 0 0
0 0 0	1	117 142	1 1 1 4	177 154	6 6 6	6 6 6	0 0 0
0 0 0	12	33 45	1 1 1 5	201 188	6 6 6	6 6 6	0 0 0
0 0 0	11	67 64	1 1 1 6	80 70	6 6 6	6 6 6	0 0 0
0 0 0	10	70 74	1 1 1 7	7 16	6 6 6	6 6 6	0 0 0
0 0 0	9	32 44	1 1 1 8	50 48	6 6 6	6 6 6	0 0 0
0 0 0	8	39 45	1 1 1 9	35 41	6 6 6	6 6 6	0 0 0
0 0 0	7	56 66	1 1 1 10	74 67	6 6 6	6 6 6	0 0 0
0 0 0	6	45 44	1 1 1 11	83 70	6 6 6	6 6 6	0 0 0
0 0 0	5	159 165	1 1 1 12	39 39	6 6 6	6 6 6	0 0 0
0 0 0	4	37 41	1 1 1 13	226 191	6 6 6	6 6 6	0 0 0
0 0 0	3	229 292	1 1 1 14	253 254	6 6 6	6 6 6	0 0 0
0 0 0	2	20 27	1 1 1 15	113 98	6 6 6	6 6 6	0 0 0
0 0 0	1	41 46	1 1 1 16	182 174	6 6 6	6 6 6	0 0 0
0 0 0	0	10 40	1 1 1 17	126 114	6 6 6	6 6 6	0 0 0
0 0 0	8	23 19	1 1 1 18	67 76	6 6 6	6 6 6	0 0 0
0 0 0	7	37 37	1 1 1 19	7 35 37	6 6 6	6 6 6	0 0 0
0 0 0	6	26 15	1 1 1 20	10 68 59	6 6 6	6 6 6	0 0 0
0 0 0	5	56 47	1 1 1 21	11 60 52	6 6 6	6 6 6	0 0 0
0 0 0	4	12 14	1 1 1 22	3 85 64	6 6 6	6 6 6	0 0 0
0 0 0	3	22 9	1 1 1 23	123 120	6 6 6	6 6 6	0 0 0
0 0 0	2	65 -32	1 1 1 24	5 82 77	6 6 6	6 6 6	0 0 0
0 0 0	1	66 71	1 1 1 25	6 71 52	6 6 6	6 6 6	0 0 0
0 0 0	0	131 64	1 1 1 26	7 20 23	6 6 6	6 6 6	0 0 0
0 0 0	10	18 21	1 1 1 27	9 33 32	6 6 6	6 6 6	0 0 0
0 0 0	9	11 12	1 1 1 28	1 55 52	6 6 6	6 6 6	0 0 0
0 0 0	8	17 22	1 1 1 29	45 45	6 6 6	6 6 6	0 0 0
0 0 0	7	18 20	1 1 1 30	42 58	6 6 6	6 6 6	0 0 0
0 0 0	6	32 28	1 1 1 31	10 19 29	6 6 6	6 6 6	0 0 0
0 0 0	5	35 24	1 1 1 32	11 23 34	6 6 6	6 6 6	0 0 0
0 0 0	4	37 31	1 1 1 33	10 77 73	6 6 6	6 6 6	0 0 0
0 0 0	3	45 41	1 1 1 34	9 53 58	6 6 6	6 6 6	0 0 0
0 0 0	2	50 53	1 1 1 35	8 51 56	6 6 6	6 6 6	0 0 0
0 0 0	1	127 144	1 1 1 36	7 103 98	6 6 6	6 6 6	0 0 0
0 0 0	0	226 -18	1 1 1 37	6 113 110	6 6 6	6 6 6	0 0 0
0 0 0	9	45 59	1 1 1 38	5 131 125	6 6 6	6 6 6	0 0 0
0 0 0	8	4 16 22	1 1 1 39	4 116 112	6 6 6	6 6 6	0 0 0
0 0 0	7	28 33	1 1 1 40	3 154 158	6 6 6	6 6 6	0 0 0
0 0 0	6	41 49	1 1 1 41	2 234 262	6 6 6	6 6 6	0 0 0
0 0 0	5	26 13	1 1 1 42	1 43 47	6 6 6	6 6 6	0 0 0
0 0 0	4	152 128	1 1 1 43	0 74 -54	6 6 6	6 6 6	0 0 0
0 0 0	3	160 126	1 1 1 44	-12 128 162	6 6 6	6 6 6	0 0 0
0 0 0	2	130 110	1 1 1 45	-3 129 128	6 6 6	6 6 6	0 0 0
0 0 0	1	133 105	1 1 1 46	-14 233 248	6 6 6	6 6 6	0 0 0
0 0 0	0	10 71	1 1 1 47	-15 141 139	6 6 6	6 6 6	0 0 0
0 0 0	11	44 46	1 1 1 48	-16 60 82	6 6 6	6 6 6	0 0 0
0 0 0	10	54 58	1 1 1 49	-7 125 127	6 6 6	6 6 6	0 0 0
0 0 0	9	13 52	1 1 1 50	-8 140 138	6 6 6	6 6 6	0 0 0
0 0 0	8	1 -70 58	1 1 1 51	-9 119 117	6 6 6	6 6 6	0 0 0
0 0 0	7	2 -3 155 119	1 1 1 52	-10 19 43	6 6 6	6 6 6	0 0 0
0 0 0	6	3 -3 54 44	1 1 1 53	-11 23 25	6 6 6	6 6 6	0 0 0
0 0 0	5	4 -27 30	1 1 1 54	-12 62 64	6 6 6	6 6 6	0 0 0
0 0 0	4	5 -8 80 64	1 1 1 55	-13 45 50	6 6 6	6 6 6	0 0 0
0 0 0	3	6 -6 88 65	1 1 1 56	-14 50 56	6 6 6	6 6 6	0 0 0
0 0 0	2	7 -126 116	1 1 1 57	-15 11 17	6 6 6	6 6 6	0 0 0
0 0 0	1	8 -137 113	1 1 1 58	-16 59 52	6 6 6	6 6 6	0 0 0
0 0 0	0	9 -79 67	1 1 1 59	-17 32 34	6 6 6	6 6 6	0 0 0
0 0 0	11	10 65 56	1 1 1 60	-18 78 59	6 6 6	6 6 6	0 0 0
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0 0 0	9	12 37 40	1 1 1 62	-20 5 74	6 6 6	6 6 6	0 0 0
0 0 0	8	15 47 49	1 1 1 63	-21 61 44	6 6 6	6 6 6	0 0 0
0 0 0	7	10 101 80	1 1 1 64	-22 4 47	6 6 6	6 6 6	0 0 0
0 0 0	6	1 51 36	1 1 1 65	-23 111 123	6 6 6	6 6 6	0 0 0
0 0 0	5	-2 84 76	1 1 1 66	-24 124 129	6 6 6	6 6 6	0 0 0
0 0 0	4	-3 21 23	1 1 1 67	-25 1 57 -37	6 6 6	6 6 6	0 0 0
0 0 0	3	-4 37 14	1 1 1 68	-26 0 119 -91	6 6 6	6 6 6	0 0 0
0 0 0	2	-5 52 40	1 1 1 69	-27 -2 87 82	6 6 6	6 6 6	0 0 0
0 0 0	1	-6 106 82	1 1 1 70	-28 -3 76 59	6 6 6	6 6 6	0 0 0
0 0 0	0	-7 100 83	1 1 1 71	-29 -4 111 106	6 6 6	6 6 6	0 0 0
0 0 0	9	-8 43 39	1 1 1 72	-30 -5 104 96	6 6 6	6 6 6	0 0 0
0 0 0	8	-11 21 30	1 1 1 73	-31 -6 32 28	6 6 6	6 6 6	0 0 0
0 0 0	7	-12 21 28	1 1 1 74	-32 -7 67 57	6 6 6	6 6 6	0 0 0
0 0 0	6	-10 47 46	1 1 1 75	-33 -8 129 120	6 6 6	6 6 6	0 0 0
0 0 0	5	-11 21 26	1 1 1 76	-34 -9 107 82	6 6 6	6 6 6	0 0 0
0 0 0	4	-12 100 83	1 1 1 77	-35 -10 12 58 46	6 6 6	6 6 6	0 0 0
0 0 0	3	-13 22 44	1 1 1 78	-36 -11 10 37 43	6 6 6	6 6 6	0 0 0
0 0 0	2	-14 11 23	1 1 1 79	-37 -12 10 37 42	6 6 6	6 6 6	0 0 0
0 0 0	1	-15 41 40	1 1 1 80	-38 -13 10 37 41	6 6 6	6 6 6	0 0 0
0 0 0	0	-16 7 60	1 1 1 81	-39 -14 10 37 40	6 6 6	6 6 6	0 0 0
0 0 0	9	-17 106 82	1 1 1 82	-40 -15 10 37 39	6 6 6	6 6 6	0 0 0
0 0 0	8	-18 11 23	1 1 1 83	-41 -16 10 37 38	6 6 6	6 6 6	0 0 0
0 0 0	7	-19 47 46	1 1 1 84	-42 -17 10 37 37	6 6 6	6 6 6	0 0 0
0 0 0	6	-20 10 37	1 1 1 85	-43 -18 10 37 36	6 6 6	6 6 6	0 0 0
0 0 0	5	-21 12 37	1 1 1 86	-44 -19 10 37 35	6 6 6	6 6 6	0 0 0
0 0 0	4	-22 11 37	1 1 1 87	-45 -20 10 37 34	6 6 6	6 6 6	0 0 0
0 0 0	3	-23 12 40	1 1 1 88	-46 -21 10 37 33	6 6 6	6 6 6	0 0 0
0 0 0	2	-24 11 40	1 1 1 89	-47 -22 10 37 32	6 6 6	6 6 6	0 0 0
0 0 0	1	-25 12 40	1 1 1 90	-48 -23 10 37 31	6 6 6	6 6 6	0 0 0
0 0 0	0	-26 11 40	1 1 1 91	-49 -24 10 37 30	6 6 6	6 6 6	0 0 0
0 0 0	9	-27 12 40	1 1 1 92	-50 -25 10 37 29	6 6 6	6 6 6	0 0 0
0 0 0	8	-28 11 40	1 1 1 93	-51 -26 10 37 28	6 6 6	6 6 6	0 0 0
0 0 0	7	-29 12 40	1 1 1 94	-52 -27 10 37 27	6 6 6	6 6 6	0 0 0
0 0 0	6	-30 11 40	1 1 1 95	-53 -28 10 37 26	6 6 6	6 6 6	0 0 0
0 0 0	5	-31 12 40	1 1 1 96	-54 -29 10 37 25	6 6 6		

monoclinic forms of the crystal resulted; one with unit cell dimensions $a = 10.8 \pm 0.05$, $b = 5.80 \pm 0.03$, $c = 39.98 \pm 0.1$ Å, $\alpha = 110 \pm 1^\circ$ and space group $P2_1/c$; and another with the crystal data quoted below and whose structure has been determined. Investigation of the first form is in progress. Although crystals have been prepared and stored in a variety of ways, only the triclinic twin has been found at the beginning of X -ray photography of a particular crystal.

Diaquobis-(2-hydroxy-2-methylpropionato)copper(II), $C_8H_{18}CuO_8$, $M = 305.8$. Monoclinic prismatic (after exposure to $Cu-K_\alpha$ radiation), $a = 10.25 \pm 0.05$, $b = 5.80 \pm 0.03$, $c = 11.24 \pm 0.05$, $\alpha = 105.9^\circ \pm 0.5$, $U = 642.5$ Å 3 , $D_m = 1.647 \pm 0.01$, $Z = 2$, $D_c = 1.581$. Space group $C2/m$ (C_{2h}^3 , No. 12), general positions $(0,0,0; \frac{1}{2},\frac{1}{2},0) \pm (x,y,z; \bar{x},y,\bar{z})$. $Cu-K_\alpha$ radiation, $\mu = 17.3$ cm. $^{-1}$. Crystal size $0.15 \times 0.18 \times 0.32$ mm.

Space group $C2$ or $C2/m$ are possible, from absent reflections. With two molecules to the unit cell the copper atoms lie on the two-fold axis, and if space group $C2/m$ is used, then the chelate rings and the copper atom must lie in the mirror plane. A three-dimensional unsharpened Patterson function calculated from 406 independent reflections could be interpreted assuming that the space group was $C2/m$ with the copper atoms at $0,0,0$ and $\frac{1}{2},\frac{1}{2},0$. After a preliminary F_o synthesis, the trial was refined to convergence assuming isotropic thermal motion, and R was 0.214. Examination of the difference synthesis and the X -ray photographs suggested that the copper atom was executing anisotropic vibrations with the maximum displacement along a perpendicular to the plane of the chelate ring. There was no compelling evidence in the difference synthesis to suggest using anisotropic temperature factors for the light atoms.

Refinement continued using an anisotropic temperature factor for the copper atom. After five more refinement cycles, R dropped to 0.16, but a difference synthesis still indicated residual electron-density near the copper atom along a . After a re-examination of the difference synthesis at $R = 0.214$, the copper atoms were replaced by an arrangement whereby one half copper atom was placed at $0,0,0$ and $\frac{1}{2},\frac{1}{2},\frac{1}{2}$ and two one quarter copper atoms were placed along a at each side of the half copper atom, related by the mirror planes and about $\frac{1}{2}$ Å distant. Then, starting with the atomic parameters for $R = 0.214$, the space and isotropic temperature factors of the light atoms and the occupation numbers only for the copper atoms were refined by least-squares using the full normal matrix. The refinement converged after five cycles with $R = 0.129$. The difference synthesis was much cleaner in the region of the copper than after the anisotropic refinement. The thermal parameters of the light atoms were very similar after both refinements, but the bonded distances at $R = 0.129$ were markedly better than at $R = 0.160$, and the e.s.d.'s for the refined parameters calculated from the normal matrices were uniformly smaller at $R = 0.129$ than at $R = 0.160$. It was concluded that the last refinement came nearest to the true model. Table 8 gives observed structure amplitudes and structure factors calculated from the parameters in Table 3(e), which gave $R = 0.129$.

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