

Kegelite: infrared spectroscopy and a structural hypothesis

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Abstract

The infrared spectrum of kegelite has been measured and is discussed. The infrared spectra of 1:1 molar mixtures of leadhillite with pyrophyllite or with muscovite resemble the spectrum of kegelite, the chemical formula of which is a 1:1 stoichiometric sum of those of pyrophyllite and of the leadhillite polymorphs. Kegelite is not a physical mixture of these species, but its structure may be an interlayering of phyllosilicate and leadhillite-like structures within the unit cell. The unit cell parameters of pyrophyllite-1*Tc* and the leadhillite polymorphs are compatible with each other and can be combined to give a cell with parameters close to those of the kegelite cell. A structural model for kegelite on these bases is suggested, which is also consistent with its thermal analysis. The formula of kegelite can be written as $\text{Pb}_8(\text{SO}_4)_2(\text{CO}_3)_4(\text{OH})_4(\text{AlOH})_4(\text{Si}_4\text{O}_{10})_2$.

KEYWORDS: kegelite, infrared spectrum, pyrophyllite, leadhillite.

Introduction

KEGELITE was first described, as a new mineral species from Tsumeb, Namibia, by Medenbach and Schmetzer (1976), and has been redefined recently by Dunn *et al.* (1990), who give new quantitative analyses, also thermal analysis, X-ray diffraction, infrared spectroscopic and density data, which indicate that it has the formula $\text{Pb}_8\text{Al}_4\text{Si}_8(\text{SO}_4)_2(\text{CO}_3)_4(\text{OH})_8\text{O}_{20}$ and is monoclinic, with space group *A2*, *Am* or *A2/m*, and with *Z* = 3. The extreme flexibility of the flaky and porous material made crystallographic interpretations from single-crystal X-ray measurements difficult because of pronounced streaking of the diffraction nodes, but unit cell parameters were obtained (see Table 2). A complete structure determination will have to await the discovery of more suitable material.

This paper presents the results of a study of the infrared spectrum of kegelite, which, considered with the data obtained by Dunn *et al.* (1990) and the infrared spectra of related species, suggests an interpretation of these data in terms of a structural model, the validity of which will have to await the results of more direct structural evidence.

Infrared spectroscopy

The infrared absorption spectrum of the kegelite specimen studied by Dunn *et al.* (1990)

(NMNH 147460), and consisting of crystals carefully selected under the binocular microscope for maximum purity, was measured over the 400–4000 cm^{-1} range, using a Perkin-Elmer PE 783 precision grating infrared spectrometer. In order to minimise the effects of any moisture absorbed into KBr discs, the spectrum was measured in Nujol mull between KBr plates, and also by using a 7 mm pressed-KBr disc containing 0.7 mg of kegelite, with a blank KBr disc containing the same weight of KBr as in the sample disc in the reference beam. This latter spectrum is reproduced in Fig. 1, and wavenumbers of absorption maxima and assignments are tabulated in Table 1. The assignments are based mainly on the known positions of absorption bands for isolated ions and bonds (Hertzberg, 1945; Nakamoto, 1978) and on comparison with published assignments for related species, as indicated in the notes to Table 1.

In addition to the presence of the anions reported in Dunn *et al.* (1990), the spectrum indicates that some of the hydroxide is present in a hydrogen-bonded environment, and some in a much less hydrogen-bonded environment. The carbonate ions appear not to be appreciably distorted from their normal D_{3h} symmetry, and are therefore likely to be in sites of high symmetry. The sulphate ions are somewhat distorted from their normal T_d symmetry. The

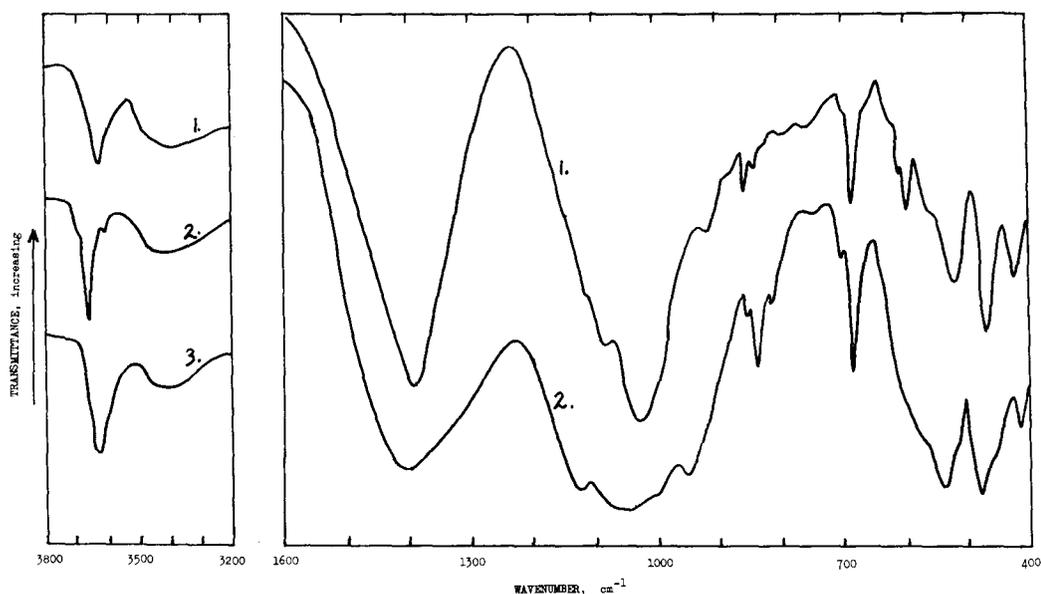


Fig. 1. Infrared spectra: 1, kegelite, off NMNH 147460; 2, a 1:1 molar physical mixture of leadhillite and pyrophyllite; 3, a 1:1 molar physical mixture of leadhillite and muscovite (O-H stretching region only).

silicate parts of the spectrum are remarkably phyllosilicate-like, as discussed below.

Structural hypothesis

The silicate parts of the infrared spectrum of kegelite are very similar to those of the phyllosilicates, in particular to those of pyrophyllite and muscovite, especially the latter. Furthermore, the O-H stretching band at 3622 cm^{-1} is similar in position and profile to that which appears in the spectra of the phyllosilicates, due to the octahedrally coordinated AlOH hydroxides.

If $(\text{AlOH})_4(\text{Si}_4\text{O}_{10})_2$, the formula for pyrophyllite, is subtracted from that of kegelite it leaves $\text{Pb}_8(\text{SO}_4)_2(\text{CO}_3)_4(\text{OH})_4$, which is that of the leadhillite polymorphs. Stoichiometrically, then, kegelite can be regarded as pyrophyllite plus a leadhillite polymorph.

The pyrophyllite structure has long been known (Gruner, 1934, and later authors), and more recently Wardle and Brindley (1972) and Lee and Guggenheim (1981) have refined the structure of a one-layer triclinic (1-*Tc*) polymorph. The layer structure is composed of two phyllosilicate sheets linked by $(\text{AlOH})_2$ units of paired AlO_6 octahedra coordinated with oxygens of the silicate sheets.

The structures of leadhillite and its polymorphs macphersonite and susannite have been determined recently, but not published except as a poster display (Highcock *et al.*, 1984). These structures are all based on different stackings of rather similar Pb_8 -containing rectangular or nearly so sub-units approximately $11.5 \times 9.1 \times 5.2\text{ \AA}$ in dimension (see Fig. 2 and Table 2). The infrared spectra of these leadhillite polymorphs are all very similar (Russell *et al.*, 1984), and also have hydrogen-bonded O-H stretching absorptions similar in profile and wavenumber to that of kegelite at 3400 cm^{-1} .

The infrared spectra of equimolar physical mixtures of leadhillite with pyrophyllite and with muscovite were measured under similar conditions to those used for the kegelite spectrum, and are reproduced in Fig. 1. The similarity between the spectrum of kegelite and those of the mixtures is remarkable, but the differences in detail show that kegelite is not a physical mixture of pyrophyllite with a leadhillite polymorph. That it is not such a physical mixture is also indicated by its exact 1:1 stoichiometry, its X-ray diffraction pattern and by its behaviour on thermal analysis, discussed below. Any interlayering of these layer lattices must therefore be within the unit cell.

Is such an interlayering possible? Kegelite,

TABLE 1. Absorption maxima in the infrared spectrum of kegelite.

Absorption maxima, cm ⁻¹ , intensities.	Assignments.
3622 mw, f sp.	O-H stretch, not appreciably hydrogen bonded. Note 1.
3400 mw, br	} O-H stretch, hydrogen bonded. Note 2.
3130 sh	
2440 vw, br	CO ₃ ²⁻ ν ₃ + ν ₁ combination band. Note 3.
1740 w, f sp	CO ₃ ²⁻ ν ₄ + ν ₁ combination band. Note 3.
1390 vs, f br	CO ₃ ²⁻ ν ₃ asymmetrical stretch. Note 4.
1150 sh (nearly buried)	} Superimposed SO ₄ ²⁻ ν ₃ asymmetrical stretch, ν ₁ symmetrical stretch, Si-O stretch and metal-O-H bending vibrations. Note 5.
1077 sh	
1039 vs	
912 wm	} CO ₃ ²⁻ ν ₂ out-of-plane bend. Note 6.
858 fw, sp	
843 w	} CO ₃ ²⁻ ν ₄ in-plane bend. Note 7.
680 m. sp	
608 sh	} SO ₄ ²⁻ ν ₄ asymmetrical bend. Note 8.
598 m, sp	
560 sh	Si-O-Si bend? Note 9.
520 ms	} Si-O bending vibrations. Note 10.
469 s	
421 ms	Pb-O stretch. Note 11.

s = strong, m = medium, w = weak, f = fairly, v = very, br = broad, sp = sharp,
sh = shoulder.

1. This closely resembles the O-H stretching band of pyrophyllite and particularly muscovite.
2. Leadhillite has a broad, H-bonded O-H absorption in just this region, centred near 3400 cm⁻¹.
3. These two weak multiphonon absorptions, common in carbonates, are combination bands as given in the table (cf. Ross and Goldsmith, 1964). Leadhillite also has these, at 2540 and 1737 cm⁻¹ (visible on the spectrum published by Russell *et al.*, 1984). The positions of these bands enables the wavenumber of the CO₃²⁻ ν₁ forbidden vibration to be calculated (~1060 cm⁻¹, see Note 5).
4. A very strong carbonate absorption, with no sign of degeneracy splitting. Thus the carbonate ion must have D_{3h} symmetry or very close to it, and is probably on a symmetrical site. In leadhillite this is at 1409 cm⁻¹ (Russell *et al.*, 1984), in caldonite at 1397 cm⁻¹ (Morgan *et al.*, 1986).
5. A complex of overlapping absorption bands. Any site symmetry distortion degeneracy splittings are hidden, but probably not large. The SO₄²⁻ ν₃ absorption in leadhillite is split, at 1042, 1087 and 1150 cm⁻¹ (Russell *et al.*, 1984), and that in caldonite is also split into three absorptions near 1000 cm⁻¹ (Morgan *et al.*, 1986). The weak (forbidden) SO₄²⁻ ν₁ symmetrical stretching vibration, revealed by anion distortion, is also buried in this complex: it is found at 964 cm⁻¹ in the spectrum of leadhillite (Russell *et al.*, 1984).
The Si-O stretching band in the muscovite layer lattice is at 1022 cm⁻¹, with shoulders at 1062, 990 and 935 cm⁻¹; in pyrophyllite, the shoulders become sharp maxima at 1121 and 950 cm⁻¹ (Farmer, 1974), doubtless due to the absence of the influence of an adjacent cation layer. In kegelite the Pb²⁺ cations could supply an effective cation layer, hence the more muscovite-like profile. In lead sulphate orthosilicates, such as mattheddleite (Livingstone *et al.*, 1987) and queitite, which also has Si₂O₇ (Povarennykh *et al.*, 1982), the S-O bands around 1050 cm⁻¹ and the Si-O absorptions around 900-950 cm⁻¹ are well-separated and distinct.
6. At 838 and 856 cm⁻¹ in leadhillite (Russell *et al.*, 1984) and at 837 and 800 (w, br) cm⁻¹ in caldonite (Morgan *et al.*, 1986).
7. Degeneracy not split, confirming D_{3h} symmetry, or close to it, for the carbonate anion. At 679 cm⁻¹ in leadhillite (Russell *et al.*, 1984).
8. The shoulder at 608 cm⁻¹ suggests distortion of the sulphate ion from T_d symmetry. At 600 cm⁻¹ in leadhillite (Russell *et al.*, 1984).
9. Si-O-Si bending vibrations are found in the 550-750 cm⁻¹ region (Strens, 1974).
10. Found in muscovite at 533 and 479 cm⁻¹ and of similar strength and profile in pyrophyllite at 541 and 483 cm⁻¹ (Farmer, 1974).
11. Found in leadhillite at 422 cm⁻¹ (Russell *et al.*, 1984) and nearby in other compounds.

pyrophyllite and the leadhillite polymorphs tend to form platy crystals of hexagonal appearance, with pearly lustres due to their pronounced cleavage. A study of their lattice parameters, considering the Pb_8 sub-units of the lead-containing species, given in Table 2, reveals an extraordinary similarity, such that a 1:1 combination of pyrophyllite-1*Tc* and leadhillite polymorph layers would be quite compatible, and could fit together to produce a cell with parameters very close to those of kegelite.

Could the kegelite structure be an alternation of pyrophyllite and leadhillite-group layers, or an insertion of pyrophyllite-1*Tc* layers into a leadhillite-group lattice? Unfortunately, these simple solutions are incompatible with the thermal analysis data. The $AlOH$ hydroxyls sandwiched between the silicate sheets of pyrophyllite and the micas lose water at 650–800 °C (Todor, 1976), whereas kegelite loses all its water by 530 °C (Dunn *et al.*, 1990). Leadhillite loses its water by 350 °C (Milodowski and Morgan, 1984). The high dehydration temperature of the phyllosilicates must be connected with the mechanism of dehydration. In the phyllosilicates, the hydroxyls are bonded to Al. In the $Al-O-H$ system the $Al-O$ bond is very strong (*cf.* corundum), but the $O-H$ bond is weak, largely because of the inductive effect of the high formal charge on the aluminium (note the well-known high acidity of the $[Al(H_2O)_6]^{3+}$ ion and the amphotericity of $Al(OH)_3$). The $Al-O-H$ system can therefore easily part with a proton, but this proton cannot easily abstract a neighbouring OH^- from its aluminium to produce the water:



As kegelite dehydrates by 530 °C the phyllosilicate system must be disrupted in such a way that its protons have access to easily removable OH^- ions, such as those in the leadhillite polymorph structures, but that the silicate sheets remain inviolate in order to explain the infrared spectrum. This opening up is likely to be symmetrical and therefore the pyrophyllite-1*Tc* double layers must be peeled apart so as to expose the $(AlOH)_2$ units, with half of these attached to each silicate sheet. This separation involves breaking $Al-O$ bonds, so suitably placed oxygens need to be attached from another source to complete the AlO_6 octahedra.

The most suitable source of such oxygen atoms are the sulphate and hydroxide ions of the leadhillite-group structures. The hydroxide-containing layers alone cannot provide sufficient

oxygens, and the carbonate groups would have to be twisted out of plane, expanding the a unit cell parameter of kegelite unacceptably and being inconsistent with the infrared data.

To derive a likely kegelite structure, the pyrophyllite half-sheets must be inserted at the correct orientation into a leadhillite-group structure along the appropriate plane (bc for leadhillite, for example) along the Pb , OH layers adjacent to the sulphate layer, separating the lead and hydroxide units so that the $(AlOH)_2$ units can coordinate with the sulphates and hydroxides (see Fig. 2). Such an insertion could not only satisfy the $Al-O$ bonding requirement but would also provide the Pb^{2+} cation layer adjacent to the silicate sheet required to explain the closer similarity of the infrared spectrum of kegelite to that of muscovite than to that of pyrophyllite.

The postulated sequence of nine layers in the unit cell of kegelite, along the a axis, would therefore be a Pb , CO_3 layer, followed by a Pb , OH layer, a silicate sheet, an $(AlOH)_2$ layer, the sulphate and hydroxide layer and finally the other $(AlOH)_2$, silicate, Pb , OH , and Pb , CO_3 layers to complete the symmetry (see Fig. 2).

On the separated pyrophyllite half-sheets only two distribution patterns of alternating $(AlOH)_2$ pairs are consistent with the unit cell requirements, with alternate hydroxide-linked pairs removed either linearly or diagonally (see Fig. 3). The resulting sheets must be inserted into the structure so that the pyrophyllite b axis becomes the kegelite c axis, which at approximately 9 Å means that rows of $(AlOH)_2$ pairs must run along the kegelite b axis, the pairs being aligned either along c (linear separation from pyrophyllite) or at 60° to c (diagonal separation from pyrophyllite) (see Fig. 3).

To complete the AlO_6 octahedra these patterns require four (pattern A of Fig. 3) or three (pattern B, with a shared O) oxygen atoms to be supplied per Pb_8 sub-unit. The sulphates alone contain sufficient oxygens for these, but the stereochemistry does not allow more than one oxygen per sulphate ion to attach to each $(AlOH)_2$ pair (two oxygens attached to one sulphur can bridge within one sheet in various patterns, but the tetrahedral configuration of sulphate does not then allow any suitable oxygen linkage patterns with the other sheet). The stereochemical requirements are considerably relaxed if the sulphates bridge with one oxygen to each sheet, and the extra oxygens are supplied by the neighbouring hydroxide groups. Pattern A of Fig. 3 would require all the hydroxides to lie within the sulphate layer to coordinate with the aluminium, but pattern B would only require the inner set of

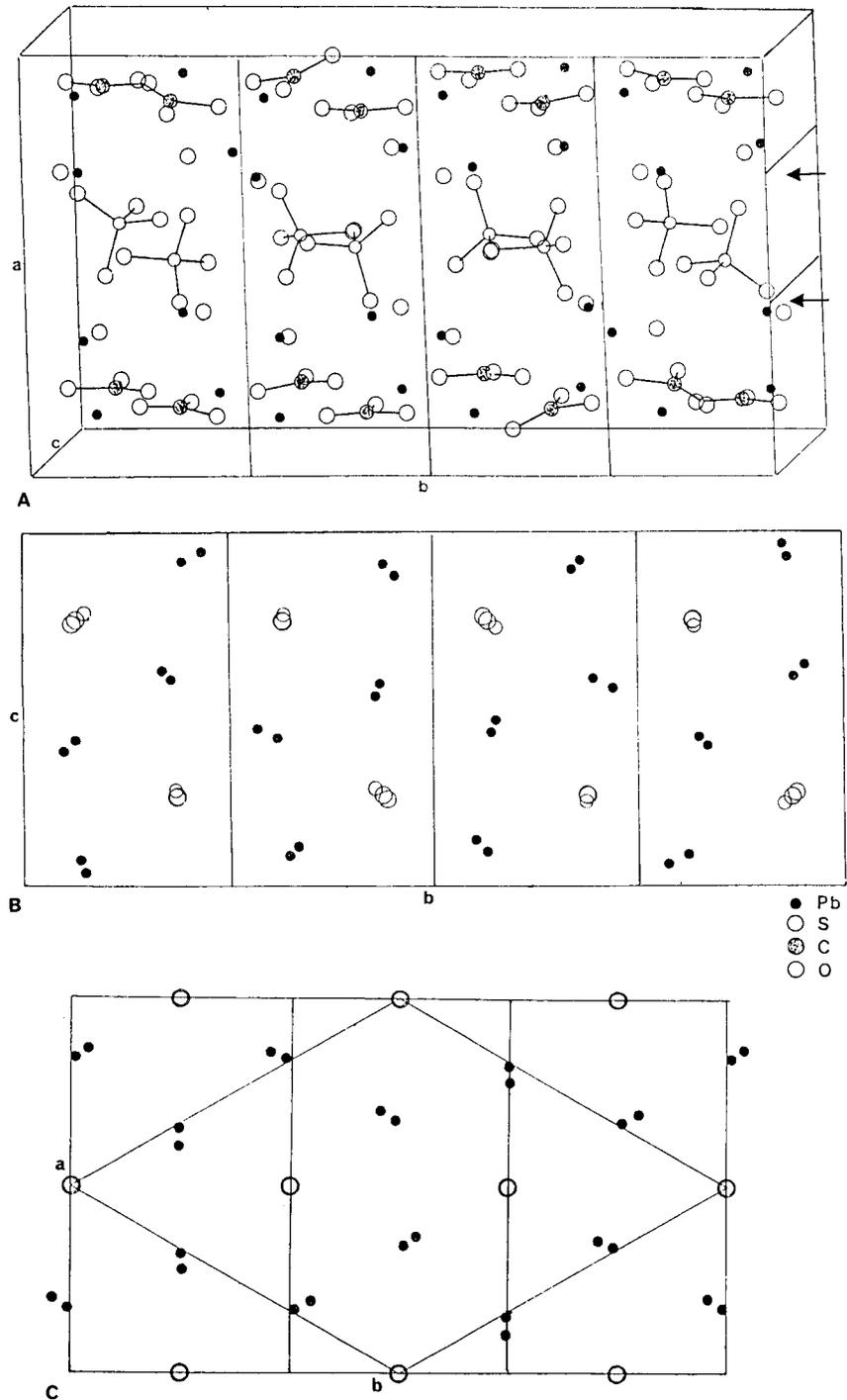


FIG. 2. Unit cells of leadhillite (A, B) and of susannite (C, *ab* plane only) (from Highcock *et al.*, 1984), divided into Pb_8 sub-units. The arrows on A indicate the likely positions of the phyllosilicate layers in kegelite. The axial lettering for kegelite is the same as that for leadhillite.

TABLE 2 Comparison of lattice parameters of suitably oriented unit cells or Pb_3 sub-units of the leadhillite polymorphs, pyrophyllite and kegelite.

Species	Comparative unit cell dimensions, Å			References
Leadhillite	a, 11.58	b/4, 5.20 (b, 20.81)	c, 9.11	Highcock <i>et al.</i> , 1984.
Susannite	c, 11.57	b/3, 5.23 (b, 15.70)	a, 9.07	Highcock <i>et al.</i> , 1984.
Macphersonite	b/2, 11.55 (b, 23.10)	c/2, 5.18 (c, 10.37)	a, 9.23	Highcock <i>et al.</i> , 1984.
Pyrophyllite-17c	c, 9.35	a, 5.16	b, 8.96	Wardle and Brindley, 1972, Lee and Guggenheim, 1981.
Kegelite	a, 21.04 (= 9.35 + 11.69)	b/3, 5.18 (b, 15.55)	c, 8.99	Dunn <i>et al.</i> , 1990.

hydroxides to coordinate with the Al, and would therefore be preferred.

The empirical formula of kegelite has all multiplicities divisible by 4 except that of sulphur, divisible by 2. The space groups $A2$ and Am proposed by Dunn *et al.* (1990) require all sites to be in multiples of 4 unless the sulphur is on a 2-fold axis or a mirror plane, when multiples of 2

are allowed. $A2/m$ requires multiples of 8 and so seems less plausible unless considerable disorder is present. In the proposed model the sulphurs occupy midline positions, bridging probably symmetrically, and so may well be on a mirror plane. If not, the structure could still be reconciled with the first two space groups if the sulphurs occupy half of four available sites. Four sulphates per

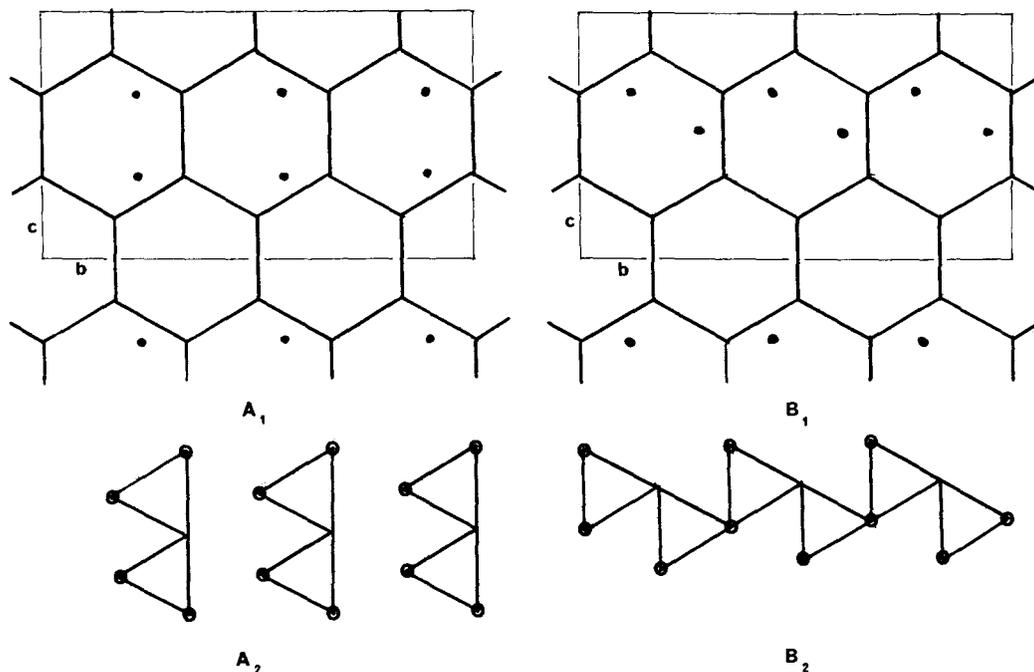


FIG. 3. A_1 , B_1 . Possible arrangements of pyrophyllite-17c half-layers within the kegelite unit cell. A, linear alignment of $(AlOH)_2$ units. B, diagonal alignment of $(AlOH)_2$ units. A_2 , B_2 . Corresponding patterns of AlO_6 octahedra in the bc plane of kegelite, showing only the faces unattached to the silicate sheet and the O atoms which have to be supplied from sulphate and hydroxide (marked o). The intersections between pairs of triangles are already occupied by OH groups of $(AlOH)_2$.

formula unit would not fit the analyses, would disturb the charge balance and would be hopelessly overcrowded because of the oxygen atoms.

With regard to the thermal analyses, the multiplicity of hydroxide environments broadens the water-loss temperature range compared with that of leadhillite. Kegelite loses SO_2 from sulphate at a somewhat lower temperature (Dunn *et al.*, 1990) than leadhillite does (Milodowski and Morgan, 1984). Aluminium sulphates lose SO_2 at a similarly lower temperature than lead sulphates do (Todor, 1976), which is consistent with the juxtaposition of Al next to sulphates in kegelite.

The infrared spectrum of the model would be expected to approximate to a superimposition of that of a leadhillite group mineral (the three polymorphs have rather similar infrared spectra) with that of a layer hydroxyaluminosilicate, as observed. Overlap of sulphate and silicate absorptions confuses assessment of sulphate coordination.

If the c parameter of pyrophyllite-1Tc is added to the appropriate parameters of the leadhillite polymorphs, a value close to the a parameter of kegelite is obtained, the other appropriate parameters of all these minerals being closely comparable with each other (Table 2). The small differences between the parameters can be accommodated by quite small shifts in the positions of the ions in the leadhillite-group structures, which also affect the symmetry within the Pb_8 sub-units and hence the repeat unit of their stacking. A stacking of alternating sub-units would forbid $Z = 3$, but would allow $Z = 6$ and a doubled b parameter for example. The b values for pyrophyllite-1Tc and kegelite are particularly close, as the fixed pyrophyllite sheet would control this value, the ionic leadhillite group structures having more flexibility in ion positions.

Conclusion

The available data suggests that the structure of kegelite is likely to be a layer-lattice one, composed of covalent phyllosilicate sheets inter-layered within the unit cell within an ionic structure resembling that of the leadhillite polymorphs (particularly susannite, as $Z = 3$), and that its formula should be written as $\text{Pb}_8(\text{SO}_4)_2(\text{CO}_3)_4(\text{OH})_4(\text{AlOH})_4(\text{Si}_4\text{O}_{10})_2$. A structural model is proposed which is consistent with the chemical and thermal analyses, physical properties, infrared spectra and unit cell parameters of kegelite. This structure may be compared with the related but simpler structure proposed for surite (Hayase *et al.*, 1978), in which

the phyllosilicate unit is not split, and sulphate and hydroxide are absent from the ionic layers.

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