

The crystal structure of faustite and its copper analogue turquoise

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ABSTRACT

The crystal structure of faustite, $\text{ZnAl}_6(\text{PO}_4)_4(\text{OH})_8 \cdot 4\text{H}_2\text{O}$, was determined using single-crystal data (Mo- $K\alpha$ X-radiation, CCD area detector, 1624 unique reflections, $R1 = 4.91\%$, $wR2 = 9.23\%$), and compared with results of a reinvestigation of the structure of its copper analogue turquoise, $\text{CuAl}_6(\text{PO}_4)_4(\text{OH})_8 \cdot 4\text{H}_2\text{O}$ (2737 unique reflections, $R1 = 2.81\%$, $wR2 = 6.90\%$). Both are isostructural and crystallize in space group $P\bar{1}$, with $a = 7.419(2)$ [turquoise: 7.410(1)], $b = 7.629(3)$ [7.633(1)], $c = 9.905(3)$ [9.904(1)] Å, $\alpha = 69.17(2)$ [68.42(1)], $\beta = 69.88(2)$ [69.65(1)], $\gamma = 64.98(2)$ [65.05(1)]°, $V = 462.2(3)$ [460.27(10)] Å³, and $Z = 1$. The structure consists of distorted MO_6 polyhedra ($M = \text{Zn}, \text{Cu}$), AlO_6 octahedra and PO_4 tetrahedra. By edge- and corner-sharing of these polyhedra a fairly dense three-dimensional framework is formed which is further strengthened by a system of hydrogen bonds. The metal atoms in the unique MO_6 ($M = \text{Zn}$ or Cu) polyhedron show a distorted [2+2+2]-coordination, the distortion being more pronounced in turquoise. About 10% of the M site is vacant in both minerals. In turquoise, a previously undetected structural site with a very low occupancy of (possibly) Cu is present at the position $(\frac{1}{2}, 0, \frac{1}{2})$.

KEYWORDS: faustite, turquoise, crystal structure, crystal chemistry, phosphates.

Introduction

FAUSTITE, $\text{ZnAl}_6(\text{PO}_4)_4(\text{OH})_8 \cdot 4\text{H}_2\text{O}$ (Erd *et al.*, 1953), is the rare Zn analogue of turquoise, $\text{CuAl}_6(\text{PO}_4)_4(\text{OH})_8 \cdot 4\text{H}_2\text{O}$, with which it forms a solid solution series. Both belong to the turquoise group which also includes the members chalcociderite, $\text{CuFe}_6^{3+}(\text{PO}_4)_4(\text{OH})_8 \cdot 4\text{H}_2\text{O}$, aheylite, $(\text{Fe}^{2+}, \text{Zn})\text{Al}_6(\text{PO}_4)_4(\text{OH})_8 \cdot 4\text{H}_2\text{O}$ and planerite, $\square\text{Al}_6(\text{PO}_4)_2(\text{PO}_3\text{OH})_2(\text{OH})_8 \cdot 4\text{H}_2\text{O}$ ($\square =$ vacancy) (see recent review of Foord and Taggart, 1998 for further details). The crystal structure of faustite has not been verified or refined to date, although it was assumed to be isostructural with turquoise, which had been determined to be triclinic, with space group $P\bar{1}$ (Cid-Dresdner, 1965; $R = 7\%$).

Recently, the first occurrence of single, well-developed crystals of faustite has been reported

from the Iron Monarch iron ore deposit in South Australia (Pring *et al.*, 2000). The pale brownish rose to colourless, platy or prismatic faustite crystals (up to 0.1 mm) are accompanied by hausmannite and kleemannite, $\text{ZnAl}_2(\text{PO}_4)_2(\text{OH})_2 \cdot 3\text{H}_2\text{O}$. Chemical analyses have demonstrated (Pring *et al.*, 2000) that the faustite contains little or no copper and only very small quantities of iron (probably Fe^{3+} substituting for Al; see Foord and Taggart, 1998). The present article reports the refinement of the crystal structure of one of these faustite crystals. For comparison purposes, we also reinvestigated the crystal structure of turquoise because the earlier structure determination (Cid-Dresdner, 1965) reported unusually large ranges for the Al–O and P–O bond lengths.

Experimental

A very small chip cut from a colourless single crystal of faustite was selected for the structure analysis. The crystal chip with the approximate

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dimensions $0.10 \times 0.05 \times 0.03$ mm was mounted on a Nonius KappaCCD diffractometer equipped with a 300 μm diameter capillary-optics collimator for improved resolution. Preliminary investigations gave a triclinic unit cell very similar to that reported for turquoise (Cid-Dresdner, 1965). For the intensity data collection at room temperature, we employed Mo- $K\alpha$ X-radiation ($\lambda = 0.71073$ Å) and a detector distance of 28 mm. Using φ and ω as rotation axes (rotation width 2°), 382 data frames were acquired, each collected over 365 s (see Table 1 for further details). The data were processed using the Nonius program suite DENZO-SMN and corrected for Lorentz, polarization and background effects. Because of the small absorption coefficient, $\mu = 2.12$ mm $^{-1}$, and the small size of the crystal chip used, no absorption correction was applied.

A pale bluish crystal of turquoise with the approximate dimensions $0.12 \times 0.08 \times 0.05$ mm was chosen for the reinvestigation of the turquoise structure. The crystal was taken from a specimen showing well-developed, pseudorhomboidal crystals, also from Iron Monarch, South Australia. Semiquantitative energy-dispersive chemical analyses of three different turquoise crystals showed them to contain only traces of Zn, K or Ca as impurity elements. The selected turquoise crystal was measured under very similar conditions (see above and Table 1). A triclinic unit cell was found which was very similar to that reported by Cid-Dresdner (1965). Again, no absorption correction was applied. Normalized structure factor statistics indicated the centrosymmetric space group $P\bar{1}$ for both faustite and turquoise crystals, in agreement with the observed crystal habits.

TABLE 1. Crystal data, data collection information, and refinement details for faustite and turquoise.

	Faustite	Turquoise
Formula (simplified)	ZnAl ₆ (PO ₄) ₄ (OH) ₈ ·4H ₂ O	CuAl ₆ (PO ₄) ₄ (OH) ₈ ·4H ₂ O
Formula weight	815.26	813.43
Space group	$P\bar{1}$	$P\bar{1}$
<i>a</i> , <i>b</i> , <i>c</i> (Å)	7.419(2), 7.629(3), 9.905(3)	7.410(1), 7.633(1), 9.904(1)
α , β , γ (°)	69.17(2), 69.88(2), 64.98(2)	68.42(1), 69.65(1), 65.05(1)
<i>V</i> (Å ³), <i>Z</i>	462.2(3), 1	460.27(10), 1
<i>F</i> (000)	408.0	407.0
μ (mm ⁻¹)	2.12	1.98
ρ_{calc} (g·cm ⁻³)	2.929	2.935
Data collection and refinement		
Total no. of frames	382	304
Collect. time per frame (s)	365	220
Collection mode	sphere	sphere
θ_{max} (°)	30.0 (used: 25.03)	30.49
<i>h</i> , <i>k</i> , <i>l</i> ranges	-8 → 8, -9 → 9, -11 → 11	-10 → 10, -10 → 10, -13 → 13
Total refls. measured	4389	5193
Unique reflections	1624	2737
<i>R</i> 1(<i>F</i>), <i>wR</i> 2(<i>F</i> ²)	4.91%, 9.23%*	2.81%, 6.90%†
'Observed' refls.	1443 [<i>F</i> _o > 4σ(<i>F</i> _o)]	2437 [<i>F</i> _o > 4σ(<i>F</i> _o)]
Extinct. factor	-	0.0173(16)
No. of refined parameters	211	213
'Goodness of Fit' (GoF)	1.212	1.106
(Δ/σ) _{max}	0.0001	0.0001
$\Delta\rho_{\text{min}}$, $\Delta\rho_{\text{max}}$ (e/Å ³)	-0.49, 0.52	-0.60, 0.75

Note: Unit-cell parameters were refined from all recorded reflections. Scattering factors for neutral atoms were employed in the refinement.

* $w = 1/[\sigma^2(F_o^2) + (0.02P)^2 + 1.9P]$; $P = ([\text{max of } (0 \text{ or } F_o^2)] + 2F_c^2)/3$

† $w = 1/[\sigma^2(F_o^2) + (0.02P)^2 + 0.5P]$; $P = ([\text{max of } (0 \text{ or } F_o^2)] + 2F_c^2)/3$

CRYSTAL STRUCTURE OF FAUSTITE

The positions of the metal and oxygen atoms in faustite were located using direct methods (SHELXS-97, Sheldrick, 1997a) and subsequent Fourier and difference Fourier syntheses. The deduced structure model is basically that of turquoise (Cid-Dresdner, 1965). Further difference Fourier syntheses revealed all 8 H atoms present in the unit cell. The positions of the H atoms given here are those which are in fact bonded to the respective O atoms, and not the symmetry-equivalent positions as listed in Cid-Dresdner (1965). Full-matrix least-squares refinement on F^2 (SHELXL-97, Sheldrick, 1997b), involving anisotropic displacement parameters for all non-H atoms, gave a discrepancy factor $R1$ of ~6%. At this stage, residuals of the electron density were in the range of $\pm 1 \text{ e}/\text{\AA}^3$. The refinement of turquoise proceeded in the same way, with $R1 \approx 3.5\%$ in the later stages of the refinement and electron density residuals $< \pm 1.2 \text{ e}/\text{\AA}^3$.

The approximate Zn:Cu ratio in the faustite crystal chip measured was then determined after the data collection by semiquantitative energy-dispersive X-ray analysis, giving a value of ~10:1. Additional point analyses of one small cluster of faustite crystals showed that the Zn:Cu ratio can be fairly variable (probably even within a single crystal), ranging from ~12:1 to ~3:1. Very minor Fe was detected in both the measured crystal and the crystal cluster.

It is known that both turquoise and faustite form a complete solid solution series with planerite, $\square\text{Al}_6(\text{PO}_4)_2(\text{PO}_3\text{OH})_2(\text{OH})_8 \cdot 4\text{H}_2\text{O}$ (i.e. a member of the turquoise group with dominant vacancies on the transition metal site), and that cuprian planerite is in fact more common than turquoise (Foord and Taggart, 1988). Accordingly, we refined the occupancy of the M ($M = \text{Zn}, \text{Cu}$) site in faustite and turquoise, neglecting the very minor Cu-for-Zn substitution in faustite. The occupancy factors obtained were 0.927(5) and 0.874(2) for faustite and turquoise, respectively, indicating that the M site contains 7 and 13% vacancies. If the M site was constrained to be fully occupied, the R factors increased considerably. During the last refinement stage for turquoise, the difference Fourier map showed a distinct peak at the position $(\frac{1}{2}, 0, \frac{1}{2})$. Considering the site coordination and bond lengths, a minor amount of Cu was preliminarily assigned to this position and U_{eq} was fixed at 0.02. Refinement of the occupancy factor gave a very low value, 0.0193(12). This site is discussed further below.

The final refinement cycle for faustite led to $R1 = 4.91\%$, $wR2 = 9.23\%$. Final R factors for turquoise were $R1 = 2.81\%$, $wR2 = 6.90\%$. Residual electron densities in both crystals were $< \pm 1 \text{ e}/\text{\AA}^3$. The final positional and displacement parameters are presented in Tables 2 and 3, and selected bond lengths and angles in Table 4. Probable hydrogen bonds are listed in Table 5. Results of empirical bond-valence calculations (Brown and Altermatt, 1985; Brese and O'Keeffe, 1991; Brown, 1996) are given in Table 6. Lists of observed and calculated structure factors (Table 7) have been deposited with the Editor.

Discussion

Structure connectivity and cation coordinations

The structure contains distorted MO_6 polyhedra ($M = \text{Zn}, \text{Cu}$), AlO_6 octahedra and PO_4 tetrahedra. The dominant structure element is an $[\text{Al}_2\text{MAl}_2(\text{O}, \text{OH}, \text{H}_2\text{O})_{18}]$ polyhedral cluster formed by a central MO_6 polyhedron sharing four edges with two edge-sharing $\text{Al}(1)\text{O}_6$ - $\text{Al}(2)\text{O}_6$ dimers (Fig. 1). The AlO_6 octahedra in the cluster are corner-linked to the $\text{Al}(3)\text{O}_6$ octahedron which shares Oh(9) with $\text{Al}(1)\text{O}_6$ and MO_6 , and Oh(12) with $\text{Al}(2)\text{O}_6$ and MO_6 . Further corner-linkage via the two non-equivalent

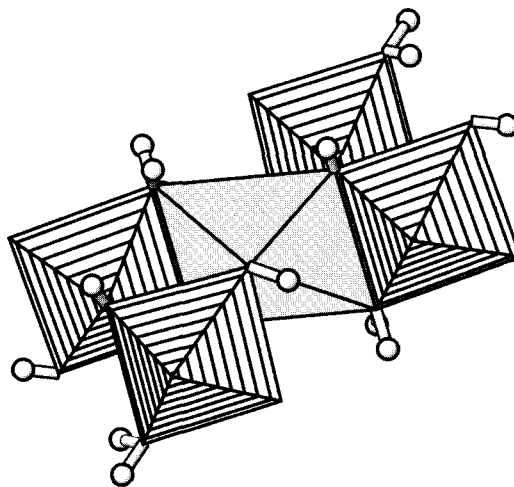


FIG. 1. The dominant element in the faustite (and turquoise) structure: an $[\text{Al}_2\text{MAl}_2(\text{O}, \text{OH}, \text{H}_2\text{O})_{18}]$ polyhedral cluster formed by a central (shaded) MO_6 polyhedron ($M = \text{Zn}, \text{Cu}$) which shares four of its edges with two edge-sharing $\text{Al}(1)\text{O}_6$ - $\text{Al}(2)\text{O}_6$ dimers. All drawings were done using ATOMS (Shape Software, 1999).

TABLE 2. Fractional atomic coordinates and isotropic displacement parameters for faustite and turquoise (every second row).

Atom	<i>x</i>	<i>y</i>	<i>z</i>	U_{eq}/U	Occupancy
Zn	0	0	0	0.0301(4)	0.927(5)
Cu	0	0	0	0.00875(13)	0.874(2)
Al1	0.2835(2)	0.1804(2)	0.74960(16)	0.0127(4)	
	0.28321(9)	0.17672(8)	0.75262(6)	0.00739(14)	
Al2	0.7478(2)	0.1861(2)	0.27319(16)	0.0123(4)	
	0.75315(9)	0.18560(8)	0.27287(6)	0.00683(14)	
Al3	0.2431(2)	0.5022(2)	0.24472(16)	0.0083(3)	
	0.24489(8)	0.50221(7)	0.24459(6)	0.00571(14)	
P(1)	0.35248(19)	0.38778(19)	0.94179(14)	0.0121(3)	
	0.35046(7)	0.38634(7)	0.94324(5)	0.00608(12)	
P(2)	0.84463(19)	0.38499(19)	0.45620(14)	0.0119(3)	
	0.84272(7)	0.38707(7)	0.45631(5)	0.00540(12)	
O(1)	0.0697(5)	0.3606(5)	0.3841(4)	0.0147(8)	
	0.0680(2)	0.3629(2)	0.38468(15)	0.0089(3)	
O(2)	0.8067(5)	0.3433(5)	0.6245(4)	0.0142(8)	
	0.8045(2)	0.34485(19)	0.62581(14)	0.0078(3)	
O(3)	0.2799(5)	0.3545(5)	0.1113(4)	0.0151(8)	
	0.2767(2)	0.3558(2)	0.11311(14)	0.0084(3)	
Ow(4)	0.0667(6)	0.0611(6)	0.1897(4)	0.0208(9)	
	0.0632(2)	0.0661(3)	0.19825(18)	0.0130(3)	
Oh(5)	0.2379(6)	0.0722(6)	0.6281(4)	0.0145(8)	
	0.2377(2)	0.0732(2)	0.62792(17)	0.0081(3)	
Oh(6)	0.7325(6)	0.0794(6)	0.1266(5)	0.0140(9)	
	0.7355(2)	0.0828(2)	0.12361(16)	0.0075(3)	
O(7)	0.2957(5)	0.4056(5)	0.6052(4)	0.0148(8)	
	0.2962(2)	0.40354(19)	0.60706(14)	0.0089(3)	
O(8)	0.3261(5)	0.2269(5)	0.9022(4)	0.0149(8)	
	0.3253(2)	0.2230(2)	0.90554(15)	0.0101(3)	
Oh(9)	0.9882(5)	0.2880(5)	0.8390(4)	0.0146(8)	
	0.9863(2)	0.2807(2)	0.84408(17)	0.0088(3)	
Ow(10)	0.5758(5)	0.0513(6)	0.6830(4)	0.0167(8)	
	0.5741(2)	0.0494(2)	0.68631(17)	0.0113(3)	
O(11)	0.7823(5)	0.4059(5)	0.1313(4)	0.0152(8)	
	0.7848(2)	0.4090(2)	0.13125(15)	0.0093(3)	
Oh(12)	0.4632(6)	0.2949(5)	0.3269(4)	0.0146(8)	
	0.4649(2)	0.2940(2)	0.32633(17)	0.0080(3)	
O(13)	0.7924(5)	0.2274(5)	0.4283(3)	0.0128(8)	
	0.7883(2)	0.22916(19)	0.43134(14)	0.0080(3)	
O(14)	0.5786(5)	0.3666(5)	0.8964(4)	0.0145(8)	
	0.5771(2)	0.3662(2)	0.89794(14)	0.0082(3)	
H(1)	0.133(13)	-0.055(14)	0.238(10)	0.08(3)	
	0.128(5)	-0.028(5)	0.224(4)	0.032(9)	
H(2)	0.148(17)	0.111(18)	0.184(13)	0.13(5)	
	0.116(5)	0.132(6)	0.176(4)	0.040(10)	
H(3)	0.653(11)	0.111(11)	0.608(8)	0.05(2)	
	0.628(5)	0.117(5)	0.609(4)	0.040(9)	
H(4)	0.607(12)	-0.067(12)	0.689(9)	0.06(3)	
	0.608(5)	-0.057(5)	0.679(3)	0.035(9)	
H(5)	0.185(10)	0.128(9)	0.580(7)	0.02(2)	
	0.204(5)	0.129(5)	0.573(3)	0.023(9)	
H(6)	0.671(9)	0.135(9)	0.091(7)	0.01(2)	
	0.691(4)	0.143(4)	0.074(3)	0.018(8)	
H(7)	0.984(8)	0.339(8)	0.895(6)	0.017(16)	
	0.987(5)	0.318(5)	0.889(4)	0.029(10)	
H(8)	0.444(9)	0.315(9)	0.390(7)	0.02(2)	
	0.448(4)	0.287(4)	0.397(3)	0.021(8)	
X (turqu.)*	0.5	0	0.5	0.02	~0.03

Note: Oxygens of hydroxyl groups are designated as Oh, those of water molecules as Ow.

* *U* was fixed (for possible identity of *X* see text). Given occupancy refers to *X* = Cu

CRYSTAL STRUCTURE OF FAUSTITE

TABLE 3. Anisotropic displacement parameters for faustite and turquoise (every second row).

Atom	U_{11}	U_{22}	U_{33}	U_{23}	U_{13}	U_{12}
Zn	0.0271(7)	0.0217(6)	0.0263(7)	-0.0087(5)	0.0128(5)	-0.0069(5)
Cu	0.0074(2)	0.00798(19)	0.0066(2)	-0.00188(14)	0.00203(13)	-0.00159(13)
Al1	0.0145(8)	0.0107(7)	0.0134(8)	-0.0031(6)	-0.0038(6)	-0.0043(6)
	0.0102(3)	0.0059(3)	0.0073(3)	-0.0011(2)	-0.0030(2)	-0.0036(2)
Al2	0.0151(8)	0.0108(7)	0.0116(8)	-0.0024(6)	-0.0039(6)	-0.0044(6)
	0.0081(3)	0.0061(3)	0.0071(3)	-0.0016(2)	-0.0024(2)	-0.0029(2)
Al3	0.0079(7)	0.0072(7)	0.0097(7)	-0.0031(6)	-0.0020(6)	-0.0016(6)
	0.0061(3)	0.0059(3)	0.0053(3)	-0.0020(2)	-0.0006(2)	-0.0023(2)
P(1)	0.0130(7)	0.0117(7)	0.0125(7)	-0.0040(5)	-0.0032(5)	-0.0040(5)
	0.0066(2)	0.0066(2)	0.0058(2)	-0.00183(16)	-0.00136(16)	-0.00288(17)
P(2)	0.0127(7)	0.0112(7)	0.0124(7)	-0.0034(5)	-0.0037(5)	-0.0037(5)
	0.0059(2)	0.0057(2)	0.0053(2)	-0.00167(16)	-0.00111(16)	-0.00262(17)
O(1)	0.0143(18)	0.0126(17)	0.0153(18)	-0.0024(15)	-0.0032(15)	-0.0039(15)
	0.0071(6)	0.0103(6)	0.0086(6)	-0.0022(5)	-0.0002(5)	-0.0037(5)
O(2)	0.0179(19)	0.0132(18)	0.0127(18)	-0.0036(15)	-0.0035(15)	-0.0061(15)
	0.0094(6)	0.0084(6)	0.0057(6)	-0.0033(5)	-0.0008(5)	-0.0029(5)
O(3)	0.0182(18)	0.0146(18)	0.0138(18)	-0.0050(15)	-0.0008(15)	-0.0079(15)
	0.0113(7)	0.0095(6)	0.0050(6)	-0.0030(5)	-0.0002(5)	-0.0047(5)
Ow(4)	0.020(2)	0.020(2)	0.022(2)	-0.0044(18)	-0.0066(17)	-0.0053(19)
	0.0130(7)	0.0120(8)	0.0146(8)	0.0003(6)	-0.0057(6)	-0.0060(7)
Oh(5)	0.018(2)	0.0122(19)	0.014(2)	-0.0030(17)	-0.0082(17)	-0.0031(16)
	0.0119(7)	0.0068(6)	0.0069(7)	0.0000(5)	-0.0054(5)	-0.0033(5)
Oh(6)	0.015(2)	0.013(2)	0.014(2)	-0.0023(17)	-0.0056(18)	-0.0032(17)
	0.0087(6)	0.0069(6)	0.0069(7)	-0.0005(5)	-0.0041(5)	-0.0020(5)
O(7)	0.0161(19)	0.0100(17)	0.0182(19)	-0.0037(15)	-0.0041(15)	-0.0042(15)
	0.0094(7)	0.0064(6)	0.0089(6)	-0.0009(5)	-0.0022(5)	-0.0019(5)
O(8)	0.0193(19)	0.0131(18)	0.0159(18)	-0.0021(15)	-0.0076(15)	-0.0071(15)
	0.0142(7)	0.0083(6)	0.0109(7)	-0.0027(5)	-0.0054(5)	-0.0044(5)
Oh(9)	0.017(2)	0.0152(19)	0.014(2)	-0.0084(16)	-0.0025(15)	-0.0047(15)
	0.0085(7)	0.0104(7)	0.0092(7)	-0.0059(5)	-0.0001(5)	-0.0035(5)
Ow(10)	0.018(2)	0.013(2)	0.018(2)	-0.0016(17)	-0.0023(16)	-0.0074(17)
	0.0096(7)	0.0092(7)	0.0129(7)	-0.0017(5)	-0.0006(5)	-0.0039(6)
O(11)	0.0143(18)	0.0123(18)	0.0171(19)	-0.0016(15)	-0.0051(15)	-0.0033(15)
	0.0091(6)	0.0075(6)	0.0099(7)	-0.0008(5)	-0.0024(5)	-0.0024(5)
Oh(12)	0.015(2)	0.015(2)	0.013(2)	-0.0062(17)	-0.0024(16)	-0.0034(15)
	0.0078(6)	0.0092(6)	0.0056(7)	-0.0025(5)	-0.0014(5)	-0.0013(5)
O(13)	0.0178(18)	0.0124(17)	0.0095(18)	-0.0021(14)	-0.0032(14)	-0.0067(15)
	0.0123(7)	0.0078(6)	0.0079(6)	-0.0025(5)	-0.0038(5)	-0.0051(5)
O(14)	0.0112(18)	0.0149(18)	0.0174(19)	-0.0039(15)	-0.0036(14)	-0.0043(14)
	0.0065(6)	0.0091(6)	0.0078(6)	-0.0005(5)	-0.0016(5)	-0.0030(5)

PO₄ tetrahedra forms a three-dimensional framework. Hydrogen bonds provide additional links between the polyhedral components (see also below). For further details and alternative descriptions of the connectivity, the reader is referred to Cid-Dresdner (1965) and the description of the structure of chalcociderite, CuFe₆³⁺(PO₄)₄(OH)₈·4H₂O, the Fe analogue of turquoise (Giuseppetti *et al.*, 1989).

In both faustite and turquoise the transition metal atoms show a [2+2+2]-coordination, which is, however, less pronounced in faustite (2x 1.921, 2x 2.197 and 2x 2.335 Å) than in turquoise (2x 1.895, 2x 2.121 and 2x 2.427 Å). In chalcociderite, CuFe₆³⁺(PO₄)₄(OH)₈·4H₂O, the Cu atom shows a comparable [2+2+2]-coordination (2x 1.914, 2x 2.079 and 2x 2.484 Å; Giuseppetti *et al.*, 1989). The bond distances in turquoise are in accordance

TABLE 4. Selected bond distances (Å) for the coordination polyhedra in faustite and turquoise ($M = \text{Zn}$ and Cu , respectively).

	Faustite	Turquoise		Faustite	Turquoise
$M-\text{Oh}(6) \times 2$	1.921(4)	1.8947(14)	$\text{Al}(1)-\text{O}(7)$	1.817(4)	1.8192(14)
$-\text{Oh}(9) \times 2$	2.196(4)	2.1209(15)	$-\text{O}(8)$	1.821(4)	1.8240(14)
$-\text{Ow}(4) \times 2$	2.334(4)	2.4272(16)	$-\text{Oh}(5)$	1.862(4)	1.8682(15)
$\langle M-\text{O} \rangle$	2.15	2.148	$-\text{Ow}(10)$	1.944(4)	1.9323(16)
			$-\text{Oh}(6)$	1.953(4)	1.9467(16)
			$-\text{Oh}(9)$	1.980(4)	1.9887(16)
			$\langle \text{Al}(1)-\text{O} \rangle$	1.896	1.897
$\text{P}(1)-\text{O}(8)$	1.516(3)	.5209(14)	$\text{Al}(2)-\text{O}(11)$	1.814(4)	1.8152(14)
$-\text{O}(11)$	1.536(4)	1.5305(14)	$-\text{O}(13)$	1.831(3)	1.8383(14)
$-\text{O}(14)$	1.532(4)	1.5341(14)	$-\text{Oh}(5)$	1.837(4)	1.8412(16)
$-\text{O}(3)$	1.538(4)	1.5378(13)	$-\text{Oh}(12)$	1.868(4)	1.8888(15)
$\langle \text{P}(1)-\text{O} \rangle$	1.531	1.531	$-\text{Oh}(6)$	1.954(4)	1.9710(15)
			$-\text{Ow}(4)$	2.121(4)	2.0492(17)
$\text{P}(2)-\text{O}(1)$	1.531(4)	1.5280(14)	$\langle \text{Al}(2)-\text{O} \rangle$	1.904	1.901
$-\text{O}(7)$	1.531(3)	1.5291(14)			
$-\text{O}(2)$	1.528(3)	1.5343(14)	$\text{Al}(3)-\text{O}(14)$	1.887(4)	1.8788(14)
$-\text{O}(13)$	1.541(3)	1.5407(13)	$-\text{O}(2)$	1.895(3)	1.8841(14)
$\langle \text{P}(2)-\text{O} \rangle$	1.533	1.533	$-\text{O}(1)$	1.903(4)	1.8999(14)
			$-\text{O}(3)$	1.908(3)	1.9040(14)
			$-\text{Oh}(12)$	1.909(4)	1.9063(15)
			$-\text{Oh}(9)$	1.977(4)	2.0108(16)
			$\langle \text{Al}(3)-\text{O} \rangle$	1.913	1.914

TABLE 5. Probable hydrogen bonds (Å) in faustite and turquoise.

	Faustite	Turquoise
$\text{Ow}(4)-\text{H}(1)\cdots\text{O}(2)$	2.913(6)	2.906(2)
$-\text{H}(1)\cdots\text{O}(10)$	2.995(6)	2.952(2)
$\text{Ow}(4)-\text{H}(2)\cdots\text{O}(3)$	3.005(6)	2.961(2)
$-\text{H}(2)\cdots\text{O}(8)$	3.023(5)	3.020(2)
$\text{Ow}(10)-\text{H}(3)\cdots\text{O}(13)$	2.716(5)	2.702(2)
$\text{Ow}(10)-\text{H}(4)\cdots\text{O}(12)$	2.814(6)	2.806(2)
$-\text{H}(4)\cdots\text{O}(3)$	3.015(5)	2.983(2)
$\text{Oh}(5)-\text{H}(5)\cdots\text{O}(1)$	2.865(6)	2.855(2)
$-\text{H}(5)\cdots\text{O}(2)$	2.990(5)	3.001(2)
$\text{Oh}(6)-\text{H}(6)\cdots\text{O}(14)$	2.724(6)	2.685(2)
$\text{Oh}(9)-\text{H}(7)\cdots\text{O}(3)$	2.707(5)	2.720(2)
$-\text{H}(7)\cdots\text{O}(11)$	3.029(5)	3.020(2)
$\text{Oh}(12)-\text{H}(8)\cdots\text{O}(7)$	2.875(5)	2.896(2)

Note: Values given are O—O distances. Refined (unrestrained) donor—hydrogen distances range between 0.59 and 0.88 Å (faustite), and 0.69 and 0.83 Å (turquoise)

with the bimodal distribution of Cu—O distances (maxima at 1.97 and 2.44 Å) observed for Jahn-Teller-distorted $\text{Cu}^{2+}-\text{O}$ polyhedra in copper oxysalt minerals (Eby and Hawthorne, 1993). The atomic environment of Cu in turquoise is in accordance with published electron paramagnetic resonance and optical absorption spectra (Sharma *et al.*, 1988). Bond-valence sums for Zn and Cu are very close to ideal values (Table 6).

The partially occupied site in turquoise at $(\frac{1}{2}, 0, \frac{1}{2})$, on which a very minor amount of Cu was assumed to be located, has bond distances to nearest oxygens compatible with a $[2+2+2]$ -coordination of Cu, 2x 1.89, 2x 2.25, 2x 2.29 Å. The environment of this site is also very similar to that of the M site (Fig. 2). However, trace impurity cations such as Na or Ca, or water molecules could occupy the site, if one considers the possibility of positional disorder of these cations. No residual electron density was observed for faustite. Occupancy of the additional site in turquoise may be correlated with the lower partial occupancy of the principal Cu site at $(0, 0, 0)$. According to Foord and Taggart (1988),

CRYSTAL STRUCTURE OF FAUSTITE

TABLE 6. Results of empirical bond-valence calculations for faustite and turquoise; for the latter only sum values (in brackets) are given.

	<i>M</i>	Al(1)	Al(2)	Al(3)	P(1)	P(2)	Sum
O(1)	—	—	—	0.506	—	1.218	1.72 (1.74)
O(2)	—	—	—	0.517	—	1.228	1.75 (1.74)
O(3)	—	—	—	0.499	1.195	—	1.69 (1.70)
Ow(4)	0.182 ↓x2	—	0.281	—	—	—	0.65 (0.47)
Oh(5)	—	0.565	0.605	—	—	—	1.17 (1.15)
Oh(6)	0.556 ↓x2	0.442	0.441	—	—	—	1.44 (1.43)
O(7)	—	0.638	—	—	—	1.218	1.86 (1.86)
O(8)	—	0.632	—	—	1.269	—	1.90 (1.88)
Oh(9)	0.265 ↓x2	0.411	—	0.414	—	—	1.09 (1.08)
Ow(10)	—	0.453	—	—	—	—	0.45 (0.47)
O(11)	—	—	0.644	—	1.202	—	1.85 (1.69)
Oh(12)	—	—	0.556	0.498	—	—	1.05 (1.03)
O(13)	—	—	0.615	—	—	1.186	1.80 (1.79)
O(14)	—	—	—	0.528	1.215	—	1.74 (1.75)
Sum	2.01 (1.99)	3.14 (3.14)	3.14 (3.13)	2.96 (2.97)	4.88 (4.88)	4.85 (4.85)	

Note: the bond-valence parameters used are from Brese and O'Keeffe (1991). The sum values for Zn and Cu were calculated without taking into account the minor presence of vacancies on these sites (see text and Table 2)

an increase in the number of vacancies occurs along the solid solution series planerite ($\square\text{Al}_6(\text{PO}_4)_2(\text{PO}_3\text{OH})_2(\text{OH})_8 \cdot 4\text{H}_2\text{O}$) – turquoise ($\text{CuAl}_6(\text{PO}_4)_4(\text{OH})_8 \cdot 4\text{H}_2\text{O}$) causes an increase of the protonation of the PO_4 tetrahedra of the structure. Our results suggest that the necessary charge balance might also be achieved by introducing either additional Cu^{2+} or other cations in the partially occupied site at $(\frac{1}{2}, 0, \frac{1}{2})$.

The AlO_6 octahedra and PO_4 tetrahedra in both faustite and turquoise are somewhat distorted. The relatively large range of Al–O and P–O bond lengths reported for turquoise by Cid-Dresdner (1965) ($\langle\text{P–O}\rangle$ 1.54 Å, range 1.52–1.56 Å; $\langle\text{Al–O}\rangle$ 1.92 Å, range 1.81–2.16 Å) was only partially confirmed; the ranges found in the present study are slightly smaller (Table 4), and the PO_4 tetrahedra are clearly more regular ($\langle\text{P–O}\rangle$ 1.532 Å, range 1.521–1.541 Å). Very similar ranges are present in faustite ($\langle\text{P–O}\rangle$ 1.532 Å, range 1.516–1.541 Å; $\langle\text{Al–O}\rangle$ 1.904 Å, range 1.814–2.121 Å). The PO_4 tetrahedra in chalcocyanite are comparable ($\langle\text{P–O}\rangle$ 1.536 Å, range 1.526–1.547 Å; Giuseppetti *et al.*, 1989).

The average Al(3)–O bond distance in faustite, 1.913 Å, is slightly larger than those of Al(1)–O and Al(2)–O (1.896 and 1.904 Å, respectively), but the distances are still identical within standard uncertainties. The Al(3) O_6 octahedron is clearly more regular than the Al(1) O_6 and Al(2) O_6 octahedra, and its shortest Al(3)–O bond, 1.887(4) Å, is considerably longer than those in the other two octahedra, 1.817(4) and 1.814(4) Å. Although this suggested the presence of very minor Fe on the Al(3) site, an occupancy factor refinement and the calculated bond-valence sums (Table 6) provided no convincing evidence for this. Nonetheless, in the final Fourier difference maps, residual electron density was only found close to Al(3) but not Al(1) or Al(2). The Al(3) O_6 octahedron in turquoise has a similar geometry (Table 4), although no possible substituent for Al(3) has been detected by the chemical analyses. A different Fe^{3+} –Al substitution model has been suggested for ferrian turquoise ('rashleighite'; Cid-Dresdner and Villarroel, 1972), where Fe^{3+} seemed to statistically occupy all three Al sites. On the other hand, during a

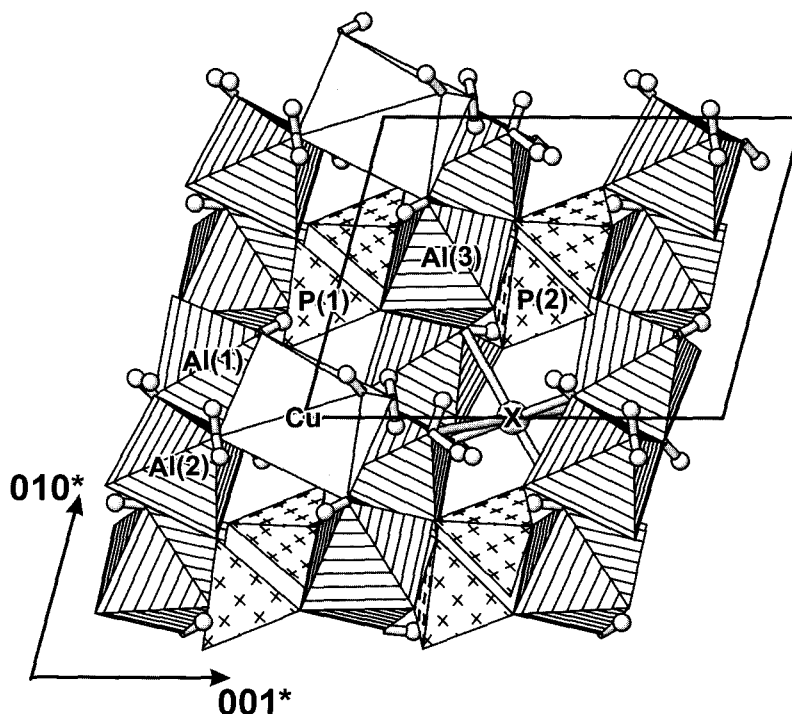


FIG. 2. View of the turquoise structure along the a -axis (the unit cell is outlined). PO_4 tetrahedra are marked with crosses, AlO_6 octahedra with parallel lines, and the distorted CuO_6 polyhedron is shaded. The large grey sphere represents the partially occupied site ('X') at the position $(\frac{1}{2}, 0, \frac{1}{2})$ (see text). The small grey spheres are H atoms.

structure analysis of an Al-bearing chalcociderite, $\text{Cu}(\text{Fe}_{3.46}^{3+}\text{Al}_{0.54})(\text{PO}_4)_4(\text{OH})_8 \cdot 4\text{H}_2\text{O}$, Giuseppetti *et al.* (1989) found evidence that Fe prefers the most regular $(\text{Fe,Al})\text{O}_6$ octahedron, i.e. that octahedron which corresponds to $\text{Al}(3)\text{O}_6$ in faustite [note that Fe(1), Fe(2A) and Fe(2B) of Giuseppetti *et al.* (1989) correspond to Al(3), Al(1) and Al(2) in the present work].

Hydrogen bonding

Probable hydrogen bonds, listed in Table 5, show reasonable bond distances and angles which are similar for both faustite and turquoise. The hydrogen bonds are medium strong to very weak, with O...O distances between 2.70 and 3.03 Å. The observed O-H...O angles accord with the geometry of hydrogen bonds with variable strengths, although the unconstrained refinement of the H atom positions leads to increased uncertainties for these angles. Some of the bonds may be bifurcated (see Table 5) and the

hydrogen bonding system appears to be flexible to some extent. The calculated bond-valence sums for the O, Oh and Ow atoms indicate a 'mixed' ligand character for some of them (Table 6). The isotropic displacement parameters of the H atoms do, however, not suggest any pronounced disorder of the H positions. The partially occupied site in turquoise at $(\frac{1}{2}, 0, \frac{1}{2})$ may be the acceptor of H bonds donated by the H(4) and H(8) atoms, if one assumes that the site is occupied by water molecules.

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CRYSTAL STRUCTURE OF FAUSTITE

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